



HAL
open science

Adsorptive Removal of Rhodamine B Dye Using Carbon Graphite/CNT Composites as Adsorbents: Kinetics, Isotherms and Thermodynamic Study

Sabrina Zghal, Ilyes Jedidi, Marc Cretin, Sophie Cerneaux, Makki Abdelmouleh

► To cite this version:

Sabrina Zghal, Ilyes Jedidi, Marc Cretin, Sophie Cerneaux, Makki Abdelmouleh. Adsorptive Removal of Rhodamine B Dye Using Carbon Graphite/CNT Composites as Adsorbents: Kinetics, Isotherms and Thermodynamic Study. *Materials*, 2023, 16 (3), pp.1015. 10.3390/ma16031015 . hal-04054095

HAL Id: hal-04054095

<https://hal.umontpellier.fr/hal-04054095v1>

Submitted on 18 Oct 2023

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.



Distributed under a Creative Commons Attribution 4.0 International License

Article

Adsorptive Removal of Rhodamine B Dye Using Carbon Graphite/CNT Composites as Adsorbents: Kinetics, Isotherms and Thermodynamic Study

Sabrine Zghal¹, Ilyes Jedidi^{1,2,*} , Marc Cretin³ , Sophie Cerneaux³  and Makki Abdelmouleh^{1,*} 

¹ Laboratory of Materials Science and Environment (LMSE), Faculty of Science of Sfax, University of Sfax, Rte. Soukra Km 4, Sfax 3000, Tunisia

² Department of Engineering, College of Engineering and Technology, University of Technology and Applied Science, Al Jamiaa Street, Suhar 311, Oman

³ Institut Européen des Membranes, IEM—UMR 5635, ENSCM, CNRS, Université de Montpellier, 34095 Montpellier, France

* Correspondence: ilyesjedidi.soh@cas.edu.om (I.J.); makki.abdmouleh@fss.rnu.tn (M.A.)

Abstract: The study of the adsorption efficiency of new carbon/CNT composites was undertaken to remove a cationic dye, Rhodamine B (RhB), from dye-contaminated wastewater. Indeed, we investigated the effect of different experimental parameters such as time, initial concentration of dye and temperature on the adsorption of RhB by the carbon composites (KS44-0 and KS44-20). The results showed that the adsorption uptake increased with the initial concentration and solution temperature while maintaining a relatively constant pH. The presence of the carbon nanotubes provided more active sites for dye removal and improved the adsorption behavior of Rhodamine B dye. The analysis of the experimental data was conducted using model equations, such as Langmuir, Freundlich and Temkin isotherms. As regards the Freundlich isotherm model, it was the best fit for the equilibrium data obtained from the experiments. The applicability of the pseudo-second-order equation could be explained assuming that the overall adsorption rate is limited by the rate of adsorbate transport that occurs on the pore surfaces of adsorbents. Furthermore, the intraparticle diffusion and Bangham models were used to investigate the diffusion mechanism of RhB adsorption onto carbon composites. They showed that multiple adsorption stages occurred simultaneously via pore surface diffusion. Concerning the thermodynamic parameters (ΔG° , ΔH° , and ΔS°), they were calculated and explained in the mean of the chemical structure of the adsorbate. Negative standard Gibbs free energy change values ($\Delta G^\circ_{\text{ads}}$) at all temperatures suggested that the adsorption process was spontaneous, and the positive values of the standard enthalpy change of adsorption ($\Delta H^\circ_{\text{ads}}$) revealed the reaction to be endothermic. The values of standard enthalpy ($\Delta H^\circ_{\text{ads}}$) and activation energy (E_a) indicated that the adsorption process corresponds to physical sorption. The mechanisms for the removal of Rhodamine B dye from wastewater using carbon composite were predicted. RhB is a planar molecule that is readily adsorbed, in which adsorbed molecules are bound by hydrophobic or other weak interactions due to the π - π interactions between the dyes' aromatic backbones and the hexagonal skeleton of graphite and carbon nanotubes. Thus, the graphite carbon/carbon nanotube composite is believed to play a major role in organic pollutant reduction.

Keywords: carbon nanotubes; Rhodamine B dye; adsorption; kinetic; isotherm; thermodynamic



Citation: Zghal, S.; Jedidi, I.; Cretin, M.; Cerneaux, S.; Abdelmouleh, M. Adsorptive Removal of Rhodamine B Dye Using Carbon Graphite/CNT Composites as Adsorbents: Kinetics, Isotherms and Thermodynamic Study. *Materials* **2023**, *16*, 1015. <https://doi.org/10.3390/ma16031015>

Academic Editors: Maning Liu, Haichang Zhang, Daohai Zhang and Zhifeng Deng

Received: 14 December 2022

Revised: 7 January 2023

Accepted: 18 January 2023

Published: 22 January 2023



Copyright: © 2023 by the authors. Licensee MDPI, Basel, Switzerland. This article is an open access article distributed under the terms and conditions of the Creative Commons Attribution (CC BY) license (<https://creativecommons.org/licenses/by/4.0/>).

1. Introduction

Synthetic dyes represent a relatively large group of organic chemical compounds, virtually found in all areas of our daily life. They make up the largest family of dyestuffs, accounting for more than half of global production. They are widely used, particularly in the dyeing factories of textile fibers, leather articles, stationary products, plastics, etc. [1]. Wastewater from the textile sector represents a real challenge in terms of dye removal since

it frequently contains diverse colors and heavy metals. Thanks to their unique physico-chemical characteristics, reactive dyes are the favored choice in the textile industry [2]. Most dye molecules are highly colored organic compounds due to the presence of chromophore and its fixed property to the acid or basic auxochromic groups (such as OH, SO₃H, NH₂, NR₂, etc.). These dyed molecules, which are widely used to impart color various substances, can be dissolved in an adequate organic solvent or incorporated into a solid matrix by physical and chemical interactions (for instance, the incorporation of groups that reduce solubility (e.g., amide groups), or the lake formation of insoluble salts of carboxylic or sulfonic acids groups) [3]. Given that these dyes have chemical structures that contain amino and carboxyl groups and are therefore highly poisonous, mutagenic, and carcinogenic, excessive quantities of these dyes in wastewater constitute a serious hazard to human health and the aquatic ecology [4]. The hazard presented by the increasing amount of dye used on a daily basis, especially on our ecosystem, has prompted a serious quest for more efficient, low-cost adsorbents.

According to earlier studies, Rhodamine B may be removed from aqueous solutions using activated carbon made from almond shells [5], and that from Palm shell was successfully tested as an adsorbent [6]. Furthermore, raw Irvingia gabonenses (dika nut) (DN) and its acid-treated form (ADN) increased the absorption of Rhodamine B (RhB) dye from aqueous solutions [7]. Activated carbon obtained from the pericarp of rubber fruit was more efficient in optimizing the adsorption process in comparison with commercial activated carbon [8]. The effects of several experimental factors such as contact time, dye concentration, activated carbon quantity, and temperature were investigated. The equilibrium results suggested that the Langmuir model was better suited to represent Congo red adsorption and had high reusability potential.

Despite the fact that there are several challenges with graphene-based composite and carbon nanotube synthesis, such as separation/dispersion, controlled manipulation, dispersion and defect-free manufacturing, graphene and carbon nanotubes are one of the most extensively utilized adsorbents. This is due to their considerable adsorption capacity towards the removal of toxic inorganic and organic contaminants from water and wastewater to reduce their environmental impact [9,10].

Our most recent publication, which uses the ferrocene chemical, describes a one-step approach for creating porous carbon graphite/carbon nanotube composites. The latter acted as both a carbon source and catalyst precursor agent for the in situ growth of carbon nanotubes (CNTs), and both phenolic resin and graphite powder were used as carbon precursors for the carbonaceous material synthesis [11]. We found that carbon nanotubes grown inside a porous carbon structure significantly improved the material's adsorption performance toward the synthetic Acid Orange 7 azo-dye (AO7) when tested as an adsorbent for the treatment of wastewater containing persistent organic pollutants. This was in contrast to the same material prepared without CNTs. [12].

In the present research work, we mainly concentrated on the application of these graphite carbon/carbon nanotube composites for the adsorption-based removal of organic dye from aqueous systems. To investigate the feasibility of our approach, we selected Rhodamine B (RhB) as a model pollutant for a long-lasting dye molecule to explore the potential use of our porous carbon composites as cost-effective carbon adsorbents through the adsorption method. Therefore, this strategy is expected to enhance the adsorption capacity of the resulting porous carbon by avoiding carbon nanotubes' growth within the porosity of the composites materials while maintaining the efficiency of RhB dye pollutants' adsorption. Overall, this study has made progress in proving the viability of synthesized carbon composites in environmental remediation and wastewater organic dye treatment. The effect of various operational parameters (contact time, initial dye concentration, temperature and adsorbent nature) on the dye removal efficiency onto carbon composites was examined. The experimental data were also interpreted using a variety of adsorption kinetics, processes, isotherm models and thermodynamics studies.

2. Materials and Methods

2.1. Materials

Two different carbon composites were elaborated based on our previous research [11]: one without ferrocene powder named KS44-0, and another containing a weight ratio of 20 wt.% of ferrocene which was added as a catalyst for the carbon nanotubes' (CNTs) growth, named KS44-20'.

Rhodamine B (Basic violet 10, Tetraethylrhodamine, [9-(o-Carboxyphenyl)-6-(diethylamino)-3H-xanthen-3-ylidene] diethylammonium chloride), ($C_{28}H_{31}ClN_2O_3$, 80%) was obtained from BIO BASIC. Benzenesulfonic acid sodium salt (4-(2-Hydroxy-1-naphthylazo)) ($C_6H_5SO_3Na$), anhydrous sodium sulphate (Na_2SO_4 , 99%), ferrous sulphate heptahydrate ($FeSO_4 \cdot 7H_2O$, >99%) and sulfuric acid (H_2SO_4 , 98%) were purchased from Sigma-Aldrich.

2.2. Adsorption Experiments

The adsorption experiments were performed using Rhodamine B (RhB) as adsorbate and carbon composite materials as adsorbents. An amount of 0.6 g of carbon composites was dispersed through a soaking method into 30 mL RhB aqueous solutions of different initial concentrations which varied from 2.5 mg/L to 20 mg/L. All the experiments were carried out by adjusting the pH at acidic pH (pH = 3) using H_2SO_4 (1 M), over a temperature range of 20 °C (± 2 °C) to 60 °C (± 2 °C). The absorbance of RhB dye was measured using a spectrophotometer, Jenway 6300 (Barioworld Scientific Ltd., Dunmow, UK), at a wavelength of 555 nm, and the adsorption percentages of RhB were calculated as follows [13]:

$$RhB_{ads} (\%) = \frac{A_i - A_e}{A_i} \times 100 \quad (1)$$

where $RhB_{ads} (\%)$ is the adsorption percentage uptake, A_i is the initial absorbance of RhB and A_e the equilibrium absorbance of RhB.

The amounts of RhB adsorbed onto the unit amount of carbon composites ($mg \cdot g^{-1}$) q_e were calculated according to Equation (2) [13]:

$$q_e = \frac{(C_0 - C_e) V}{W} \quad (2)$$

with C_0 and C_e being the initial and equilibrium concentrations of RhB ($mg \cdot L^{-1}$), respectively. V represents the volume of the RhB solution (L), and W is the weight of the carbon composites (g).

To determine the concentration of Rhodamine B, a calibration curve was established at the maximal absorbance wavelength of 555 nm. A graphical representation of the calibration curve and the chemical structure of a Rhodamine B dye molecule are shown in Figure S1.

3. Results

3.1. Kinetics Study

3.1.1. Effect of Initial Dye Concentration

Figure 1 depicts the adsorption kinetics of RhB onto carbon graphite (KS44-0) and carbon graphite/CNT (KS44-20) composite materials as a function of time (t), expressed as a percentage, for different RhB initial concentrations (C_0) in the range of 2.5–20 mg/L. The results revealed that the adsorption capacity uptake of RhB increased onto carbon composites with the increase in the initial dye concentration. As shown in Figure 1, the dye adsorption process of RhB onto KS44-0 and KS44-20 reached an equilibrium after 120 min of contact time for all the initial concentrations tested. At the lower initial concentration of 2.5 mg/L, the dye adsorption efficiency values of RhB were 98% and 55% for KS44-20 and KS44-0, respectively. However, at a 20 mg/L initial RhB concentration, the adsorption quantity on the two supports was divided by 2. In fact, the increase in the initial RhB concentration favored the adsorbent–adsorbate interactions and led to the saturation of

active sites accessible at the inner and outer surfaces of the porous carbon composites. This is because the driving force for mass transfer (i.e., the gradient of concentration) increased with the increase in the initial RhB dye concentration [14]. Actually, the existence of carbon nanotubes that expanded dramatically within the substance is the reason behind the higher RhB adsorption capacity and effectiveness seen with KS44-20 (Figure 2d). Carbon nanotubes with various structures that appeared all over the sample, on the inner faces of the pores and onto the surface, drastically changed the surface properties, thus leading to significant changes in the specific surface area and porosity of the KS44-20 composite compared to the KS44-0 composite (Figure 2b,d) [11,12].

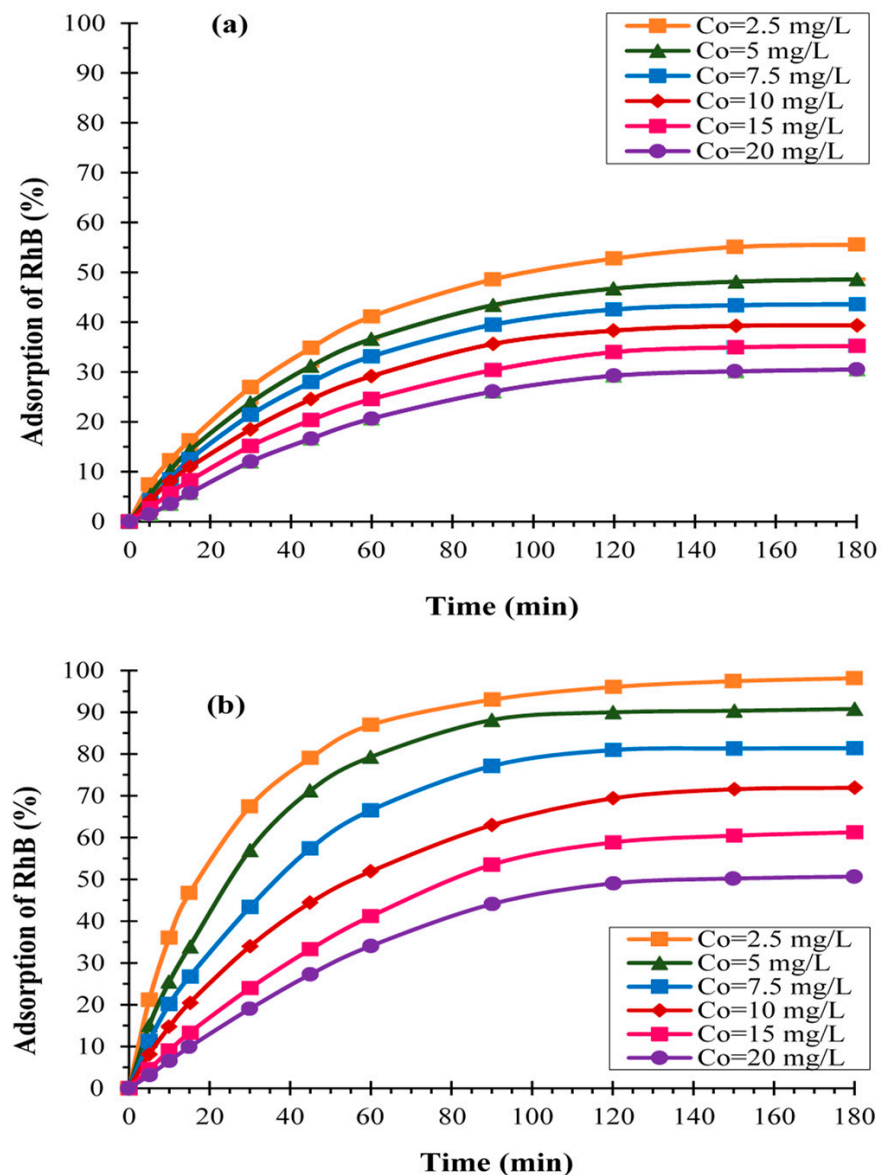


Figure 1. Evolution of the adsorption of RhB dye (%) at different initial concentrations on (a) KS44-0 and (b) KS44-20 ($T = 20\text{ }^{\circ}\text{C}$, $\text{pH} = 3$ and time = 3 h).

Additionally, morphological and structural analyses of carbon composite materials, as well as information pertaining to their surfaces, showed that KS44-20 had a more significant porous structure than that of KS44-0 (Figure 2a,c). The pore shape difference between KS44-0 and KS44-20 was well-illustrated by scanning electron microscopy, BET the surface areas and mercury porosimetry studies (Table 1) [11,12].

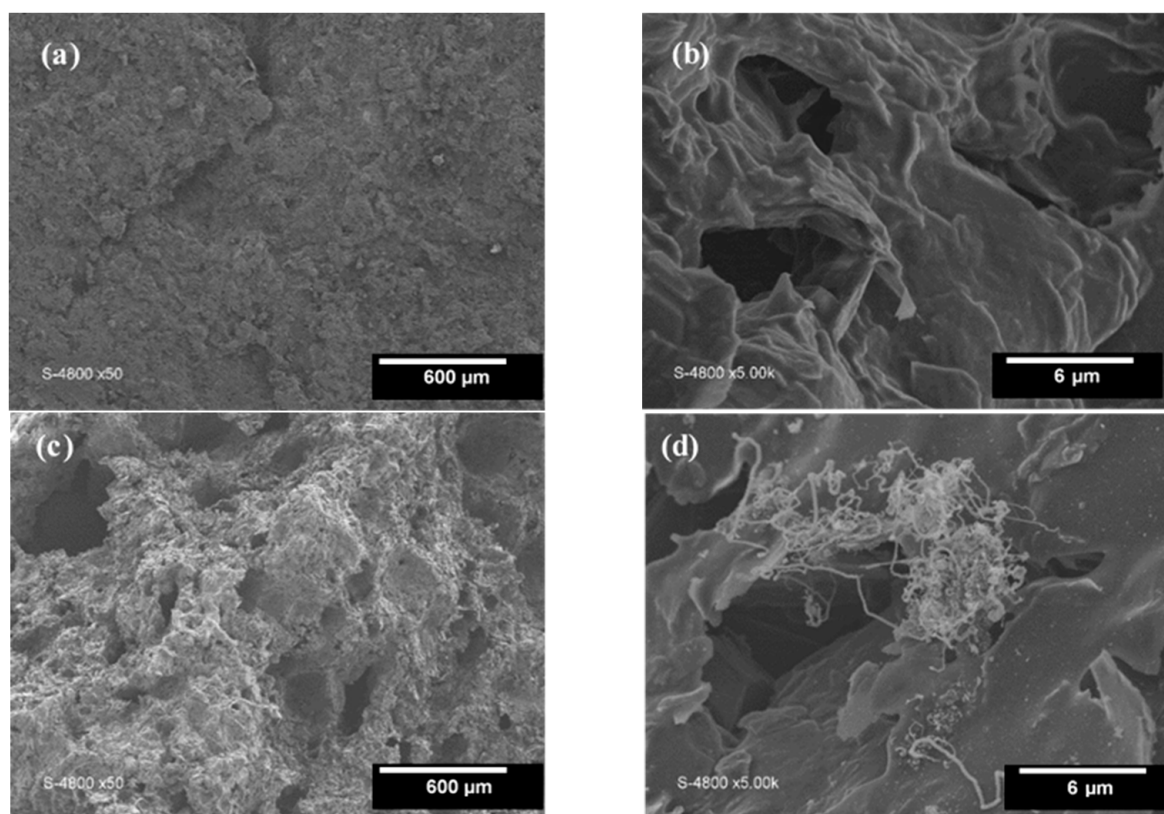


Figure 2. SEM micrographs at different scales of the material: (a,b) carbon graphite (KS44-0) and (c,d) carbon graphite/CNT composite (KS44-20).

Table 1. Textural parameters of KS44-0 and KS44-20 detected with BET method and mercury intrusion porosimetry technique [12].

Composites	Porosity (%)	Average Pore Diameter (4 V/A) (nm)	BET Surface Area (m ² /g)
KS44-0	45.65	102.20	85.67
KS44-20	52.08	150.20	131.47

3.1.2. Effect of Temperature

The experiments were performed with the carbon adsorbents KS44-20 and KS44-0 using a 10 mg/L initial dye concentration at different temperatures of 20, 30, 40, 50 and 60 °C (Figure 3). The percentage removal of RhB increased with the increase in temperature up to 60 °C. In fact, the adsorption increased from 71 to 85% and from 39 to 51% for KS44-20 and KS44-0, respectively. The kinetic energy of the adsorbent molecules was therefore affected by the RhB solution's temperature, which could result in a greater adsorption capacity on the surfaces and in the pores of both carbon composites (KS44-0 and KS44-20). Due to intermolecular forces of attraction between a solid and the material adsorbed, porous adsorbents favor the probability of adsorbate diffusion [15]. As a result, the adsorption process seems to be controlled by adsorbate diffusion and mass transfer onto different porous adsorbents, which is also favored as temperature rises. The adsorption process is likewise endothermic for RhB dye, as shown by the preceding data [16].

3.1.3. Adsorption Kinetics

The experimental adsorption data were analyzed using the pseudo-first-order [17–19] and pseudo-second-order [17–19] Elovich models [20,21] to explore the kinetics of the RhB

adsorption process over carbon composites (KS44-0 and KS44-20). Different linear form calculations were used (Equations (3)–(6)) and are presented in Figure 4.

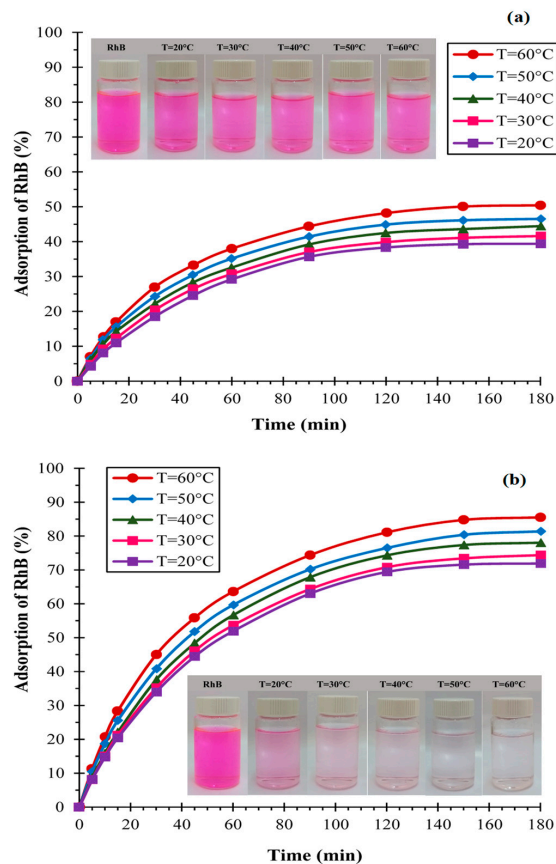


Figure 3. Evolution of the adsorption of RhB dye at different temperatures on (a) KS44-0 and (b) KS44-20.

The pseudo-first-order kinetic rate expression of Lagergren is in the form of:

$$\frac{1}{q_t} = \frac{k_1}{q_1 \times t} + \frac{1}{q_1} \tag{3}$$

where q_t and q_e ($\text{mg}\cdot\text{g}^{-1}$) are the adsorbed amounts of RhB at time t and at equilibrium, respectively, and k_1 (min^{-1}) is the adsorption rate constant. The pseudo-first-order model was examined by the extrapolation of the plot $\ln(q_e - q_t)$ vs. (time), according to Equation (4):

$$\ln(q_e - q_t) = \ln(q_e) - \frac{k_1}{2.303} \times t \tag{4}$$

The pseudo-second-order kinetics was examined by the extrapolation of the plot (t/q_t) vs. (time), according to Equation (5):

$$\frac{t}{q_t} = \frac{1}{(k_2 \times q_e^2)} + \frac{t}{q_e} \tag{5}$$

where q_e ($\text{mg}\cdot\text{g}^{-1}$) is the maximum adsorption capacity and k_2 ($\text{g}\cdot\text{mg}^{-1}\cdot\text{min}^{-1}$) is the rate constant of adsorption.

The Elovich model can be described in a linear form as defined by Equation (6):

$$q_t = \frac{1}{\beta} \text{Ln}(\alpha\beta) + \frac{1}{\beta} \text{Ln}(t) \tag{6}$$

where q_t is the amount adsorbed at time t ($\text{mg}\cdot\text{g}^{-1}$), α represents the initial adsorption rate ($\text{mg}\cdot\text{g}^{-1}\cdot\text{min}^{-1}$) and β is related to the desorption constant for each experiment ($\text{g}\cdot\text{mg}^{-1}$).

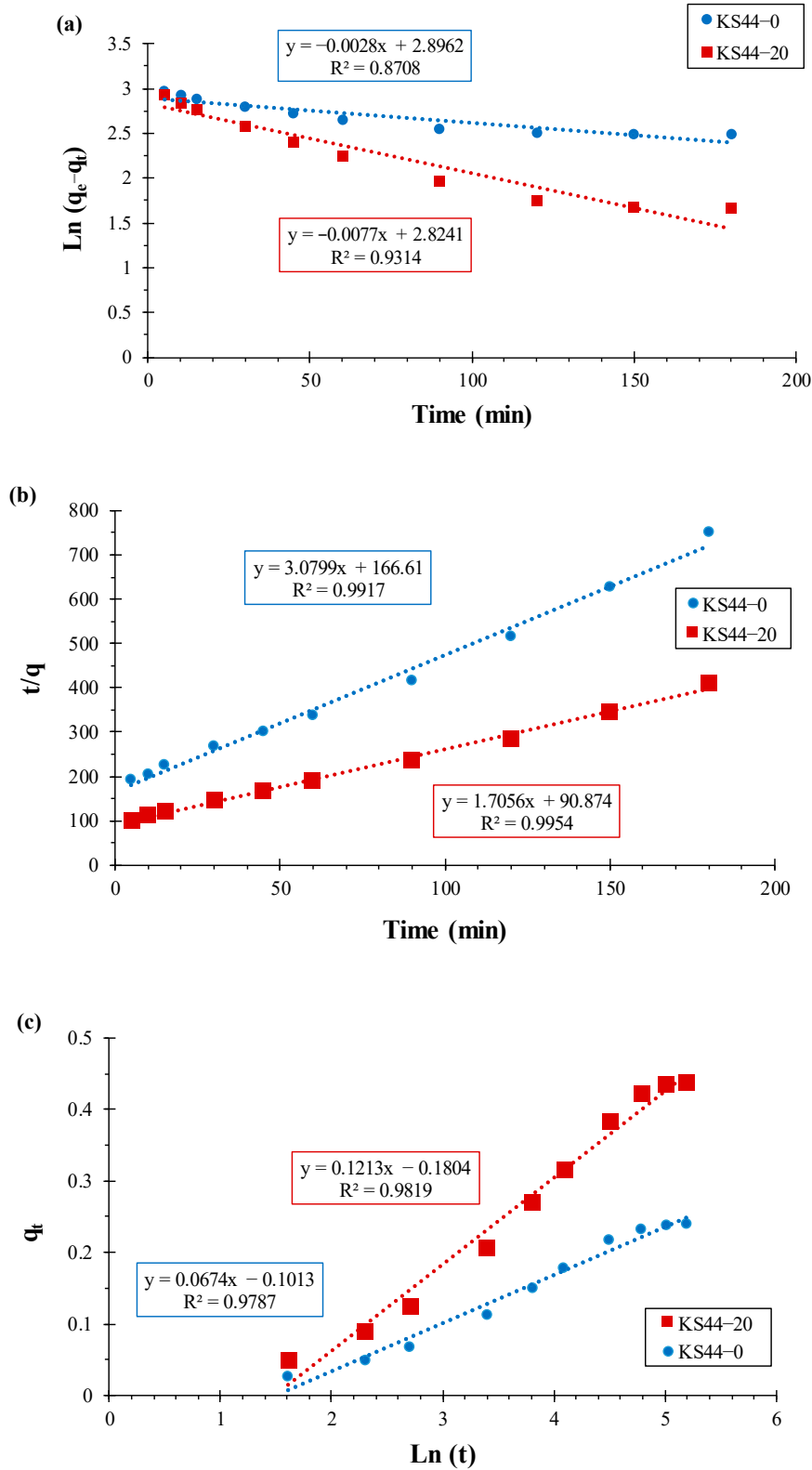


Figure 4. Kinetic parameters of the pseudo-first-order (a), pseudo-second-order (b) and the Elovich model (c) for RhB adsorption on KS44-0 and KS44-20.

The extracted adsorption parameters of the various kinetic models from Figure 4 are tabulated in Table 2, from which it is clear that the correlation coefficient (R^2), indicating the proximity of the results, shows the mechanism under study. When the value is close to 1, it is evident that the reaction is more likely to follow it. The pseudo-second-order model revealed a better fit to the adsorption data, in which the calculated values of q_e (0.249 and 0.452 $\text{mg}\cdot\text{g}^{-1}$) were the closest to the experimental values of $q_{e(\text{exp})}$ (0.239 and 0.437 $\text{mg}\cdot\text{g}^{-1}$) for the KS44-0 and KS44-20 composites, respectively. This demonstrated that the adsorption kinetics of RhB dye onto carbon composites is better described by a pseudo-second-order model.

Table 2. Kinetic parameters for RhB dye adsorption on KS44-0 and KS44-20: ($C_0 = 10 \text{ mg}\cdot\text{L}^{-1}$, $\text{pH} = 3$ and $T = 20 \text{ }^\circ\text{C}$).

Composite	Pseudo-First-Order			Pseudo-Second-Order			Elovich			$q_{e(\text{exp})}$ ($\text{mg}\cdot\text{g}^{-1}$)
	k_1 (min^{-1})	$q_e(\text{cal})$ ($\text{mg}\cdot\text{g}^{-1}$)	R^2	k_2 ($\text{g}\cdot\text{mg}^{-1}\cdot\text{min}^{-1}$)	$q_e(\text{cal})$ ($\text{mg}\cdot\text{g}^{-1}$)	R^2	β ($\text{g}\cdot\text{mg}^{-1}$)	α ($\text{mg}\cdot\text{g}^{-1}\cdot\text{min}^{-1}$)	R^2	
KS44-0	0.002	18.105	0.870	1.095	0.249	0.991	0.067	0.101	0.978	0.239
KS44-20	0.007	16.845	0.931	0.536	0.452	0.995	0.121	0.180	0.981	0.437

3.1.4. Adsorption Mechanism

The intra-particle diffusion [7,20] and Bangham models were applied to identify the adsorption mechanism. According to this concept, the intra-particle diffusion model developed by Weber and Morris has been widely applied to examine the rate-limiting step during adsorption [22], and is given by the following equation:

$$q_t = k_{id} \times t^{\frac{1}{2}} + A \quad (7)$$

where q_t ($\text{mg}\cdot\text{g}^{-1}$) is the amount of RhB adsorbed at time t , k_{id} ($\text{mg}\cdot\text{g}^{-1}\cdot\text{min}^{-1/2}$) is the rate constant for intra-particle diffusion and A is a constant.

Kinetic data were further used to determine the controlling step occurring in the adsorption system using Bangham model that is based on the assumption that pore diffusion takes place during adsorption [23,24], in the following form:

$$\log\left(\frac{C_0}{C_0 - q_t \times m}\right) = \log\left(\frac{K_B}{2.303 \times V}\right) + \alpha \text{Ln}(t) \quad (8)$$

where C_0 ($\text{mg}\cdot\text{L}^{-1}$) is the initial concentration of dye solution, V (mL) is the volume of solution, q_t ($\text{mg}\cdot\text{g}^{-1}$) is the amount of dye adsorbed at time t , m ($\text{g}\cdot\text{L}^{-1}$) is the weight of adsorbent per liter of solution and ($\alpha < 1$) and K_B are constants.

The kinetic parameters for the adsorption of RhB onto carbon composites were determined using the slope-intercept form (Figure 5) and are summarized in Table 3. The graphical representation of q_t versus $t^{1/2}$ that describes model dependencies obtained for RhB dye adsorption (Figure 5) did not pass through the origin. Next, there were two separate linear regions independently evaluated in both cases. From this figure, it is evident that intra-particle diffusion is a significant step in the process of the adsorption of RhB onto KS44-0 and KS44-20, especially in the first 60 min of the experiment. Figure 5 shows a linear plot of kinetic data, not passing through the origin, which indicates that the diffusion of adsorbate into the pores of adsorbents is not the only rate-controlling step. Nonetheless, the Bangham plot's linearity indicates that more than one process may be at work when it comes to the uptake of RhB onto carbon composite materials (Figure 6). As indicated in Table 3, the correlation coefficients (R^2) of the intra-particle diffusion and the Bangham model converge to 1, which indicates the applicability of these two adsorption kinetic models. Thus, the intra-particle diffusion model, which is highly supported by the Bangham model, is the second-best model to produce a decent match to the experiment, followed by a pseudo-second-order. This suggests that the data fit and mainly control

the adsorption process of RhB onto carbon composites with three kinetic models for all experimental conditions.

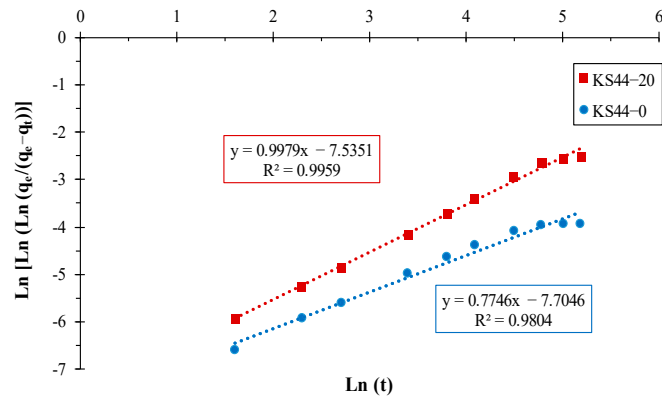


Figure 5. Intra-diffusion kinetic model representation for RhB adsorption on KS44-0 and KS44-20.

Table 3. Kinetic parameters obtained from intra-particle diffusion and Bangham models for RhB adsorption on KS44-0 and KS44-20.

Composites	Intra-Particle Diffusion Model						Bangham Model		
	Region I			Region II			K_B (min ^{-α})	A	R ²
	k_{dif} (min ⁻¹)	C (mg.L ⁻¹)	R ²	k_{dif} (min ⁻¹)	C (mg.L ⁻¹)	R ²			
KS44-0	0.027	0.037	0.998	0.010	0.105	0.846	0.774	-7.704	0.980
KS44-20	0.049	0.065	0.999	0.021	0.166	0.885	0.997	-7.535	0.995

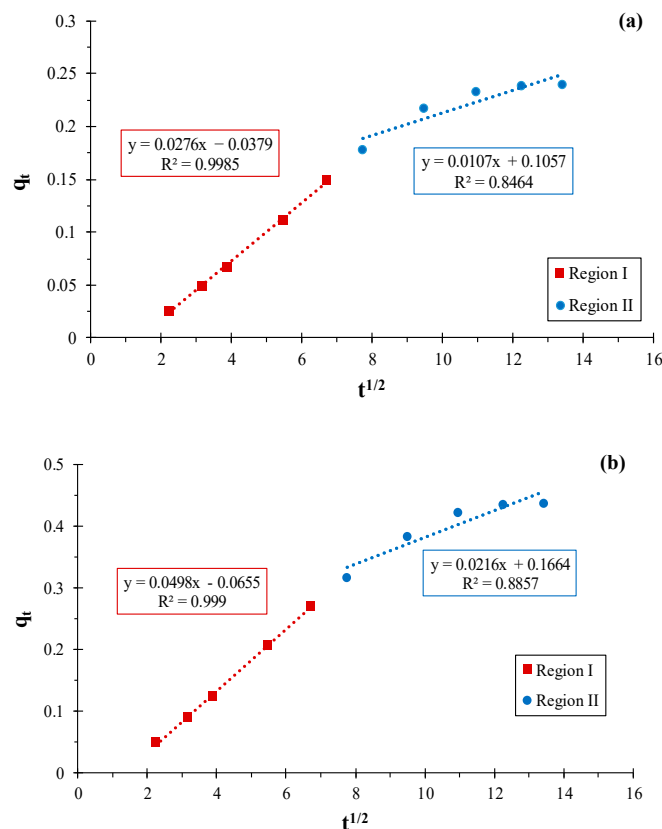


Figure 6. Bangham kinetic model representation for RhB adsorption on (a) KS44-0 and (b) KS44-20.

3.2. Adsorption Isotherm Study

The adsorption isotherms of RhB dye on KS44-0 and KS44-20 composites at room temperature were studied (Figure 7). Indeed, the adsorption isotherm describes the interaction of dyes with adsorbents, which are determinant factors for the performance of adsorption processes. According to the concentration of adsorbate remaining in solution C_e (in $\text{mg}\cdot\text{L}^{-1}$), it expresses the amount of adsorbate on the adsorbent q_e (in $\text{mg}\cdot\text{g}^{-1}$ of adsorbent). The adsorption isotherms indicate that the amount of adsorbed RhB increased with the increase in initial concentration C_e . The results demonstrate that the adsorption occurred quickly in the early stages of the experiment before processing slowed down to reach a plateau, a sign that porous composite surfaces are saturated. In the low C_e range, the adsorption of RhB adsorbent on KS44-20 was closer to the theoretical line, corresponding to 100% adsorption, which suggests a higher binding affinity of RhB dye onto KS44-20.

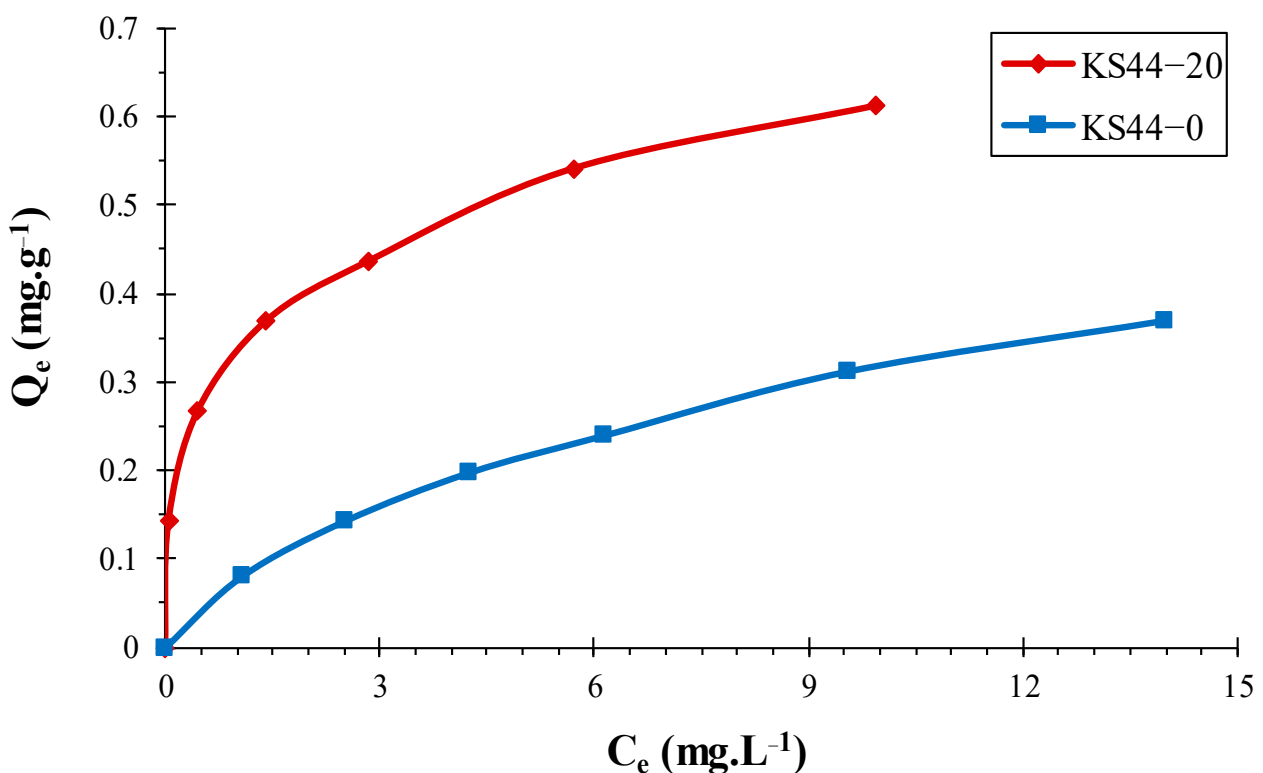


Figure 7. Adsorption isotherms of RhB on carbon composites KS44-0 and KS44-20. ($T = 20\text{ }^\circ\text{C}$, $\text{pH} = 3$, and time = 3 h).

The equilibrium nature of adsorption has been described by a number of isotherm equations. Adsorption data were fitted to the Langmuir [6,9], Freundlich [20,21] and Temkin [21,25] equations, and the isotherm parameters were calculated and reported accordingly (Figure 8). The linear transformation of the Langmuir adsorption isotherm, assuming that the adsorbate adsorption is limited to one molecular layer and the surface of the adsorbent is homogeneous, is given in Equation (9):

$$\frac{C_e}{q_e} = \frac{C_e}{q_{max}} + \frac{1}{(q_{max} \times K_L)} \quad (9)$$

where K_L ($\text{L}\cdot\text{mg}^{-1}$) is the Langmuir adsorption constant related to the energy of adsorption, q_e ($\text{mg}\cdot\text{g}^{-1}$) is the equilibrium adsorbate amount on the adsorbent, C_e ($\text{mg}\cdot\text{L}^{-1}$) is the equilibrium dye concentration and q_{max} ($\text{mg}\cdot\text{g}^{-1}$) is the maximum adsorption retained by the adsorbent at monolayer coverage [26–28].

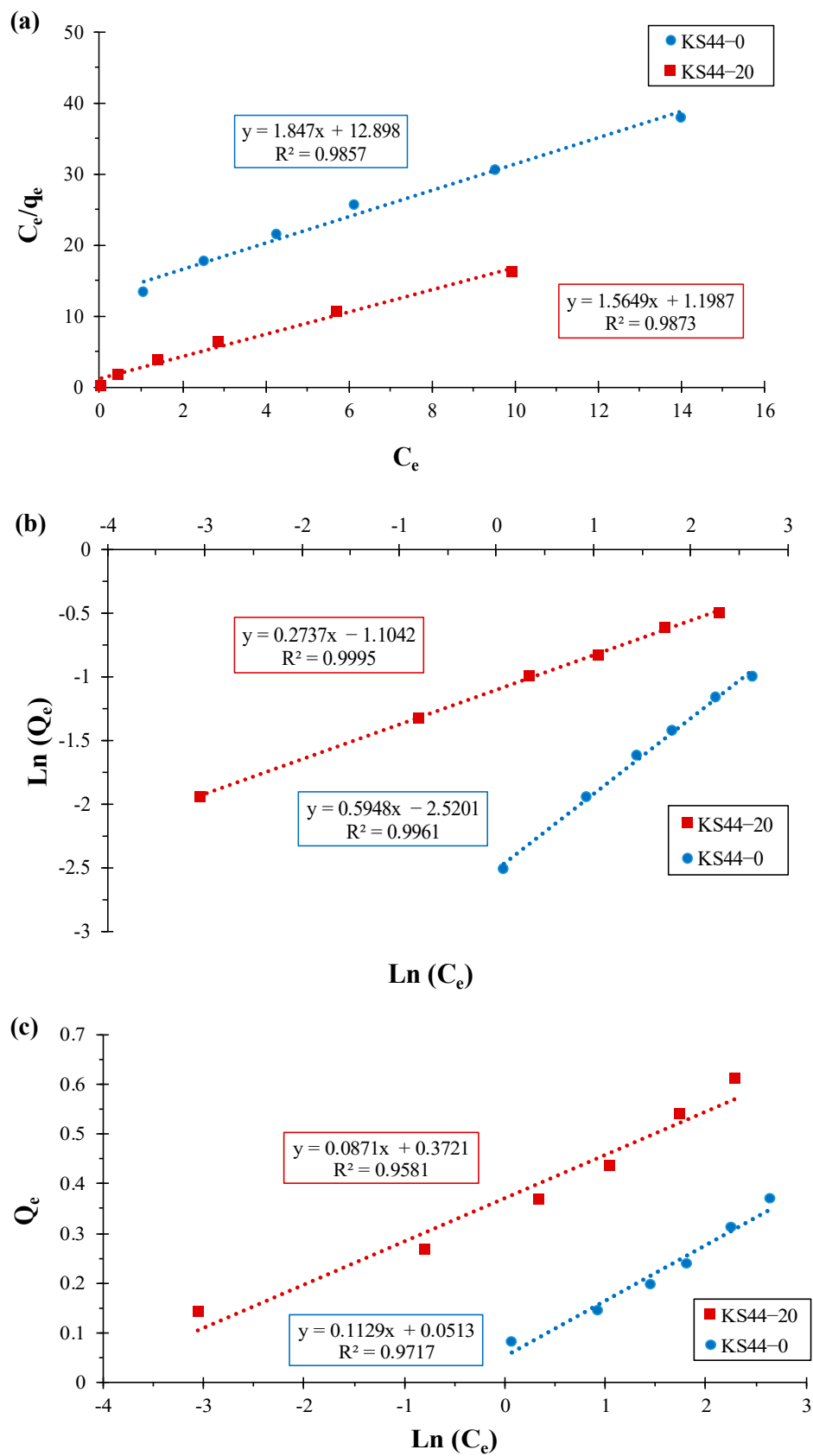


Figure 8. Transformed adsorption isotherm for (a) the Langmuir model, (b) the Freundlich model and (c) the Temkin model for RhB adsorption on KS44-0 and KS44-20 composites.

The linear transformation of the Freundlich adsorption isotherm, which reflects the description of multilayer adsorption with interactions between adsorbed molecules and heterogeneous surfaces, is given as Equation (10):

$$\ln(q_e) = \left(\frac{1}{n}\right) \times \ln(C_e) + \ln(K_F) \quad (10)$$

where K_F is the Freundlich adsorption capacity and $(1/n)$ is the Freundlich adsorption intensity.

The Temkin adsorption isotherm's linear transformation, which does not only examine the effects of some indirect adsorbate/adsorbent interactions on the adsorption process but also assumes that the heat of adsorption of all molecules would decrease linearly with the coverage of the adsorbed molecule layer, is given as Equation (11):

$$q_e = B_T \ln(K_T) + B_T \ln(C_e) \quad (11)$$

where B_T ($\text{J}\cdot\text{mol}^{-1}$) is the Temkin adsorption constant and K_T ($\text{L}\cdot\text{mg}^{-1}$) is the Temkin isotherm constant.

Table 4 provides an overview of the parameters and correlation coefficients of the Langmuir, Freundlich, and Temkin isotherm models used to study RhB adsorption on KS44-0 and KS44-20. The Freundlich model yielded the highest correlation coefficients of the data fitting ($R^2 > 0.99$). Therefore, the Freundlich model, which describes multilayer adsorption (more than one layer of molecules), in which adsorbed molecules are bound by hydrophobic or other weak interactions, provides a good representation of the adsorption isotherm of RhB dye from carbon composites [29–31].

Table 4. Langmuir, Freundlich and Temkin isotherm model parameters for RhB adsorption on KS44-0 and KS44-20 composites.

Composites	Langmuir Isotherm			Freundlich Isotherm			Temkin Isotherm		
	q_{\max} ($\text{mg}\cdot\text{g}^{-1}$)	K_L ($\text{L}\cdot\text{mg}^{-1}$)	R^2	K_F ($\text{mg}\cdot\text{g}^{-1}$)	n	R^2	B_T ($\text{J}\cdot\text{mol}^{-1}$)	K_T ($\text{L}\cdot\text{mg}^{-1}$)	R^2
KS44-0	0.541	6.983	0.985	0.080	1.681	0.996	0.112	1.575	0.971
KS44-20	0.639	0.765	0.987	0.331	3.653	0.999	0.087	71.671	0.958

3.3. Adsorption Thermodynamics

Thermodynamic parameters such as standard Gibbs free energy change ($\Delta G^\circ_{\text{ads}}$), standard enthalpy change ($\Delta H^\circ_{\text{ads}}$) and standard entropy change ($\Delta S^\circ_{\text{ads}}$) were determined using the following equations [32–34]:

$$K_D = \frac{q_e}{C_e} \quad (12)$$

$$\Delta G^\circ = -R T \times \ln(K_d) \quad (13)$$

$$\ln(K_d) = \frac{\Delta S^\circ}{R} - \frac{\Delta H^\circ}{R T} \quad (14)$$

$$\Delta G^\circ_{\text{ads}} = \Delta H^\circ_{\text{ads}} - T \times \Delta S^\circ_{\text{ads}} \quad (15)$$

where R ($8.314 \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$) is the universal gas constant, T (K) is the absolute temperature scale, K_D ($\text{L}\cdot\text{g}^{-1}$) is the equilibrium dissociation constant and C_e ($\text{mg}\cdot\text{L}^{-1}$) and q_e ($\text{mg}\cdot\text{g}^{-1}$) are the equilibrium concentration of dye in solution and the amount of dye adsorbed onto the adsorbent at equilibrium, respectively.

The values of $\Delta H^\circ_{\text{ads}}$ and $\Delta S^\circ_{\text{ads}}$ were obtained from the linear Van't Hoff plots of $\ln(K_D)$ versus $1/T$ [35–37] (depicted in Figure 9a), and the thermodynamic parameters obtained for the adsorption process are summarized in Table 5.

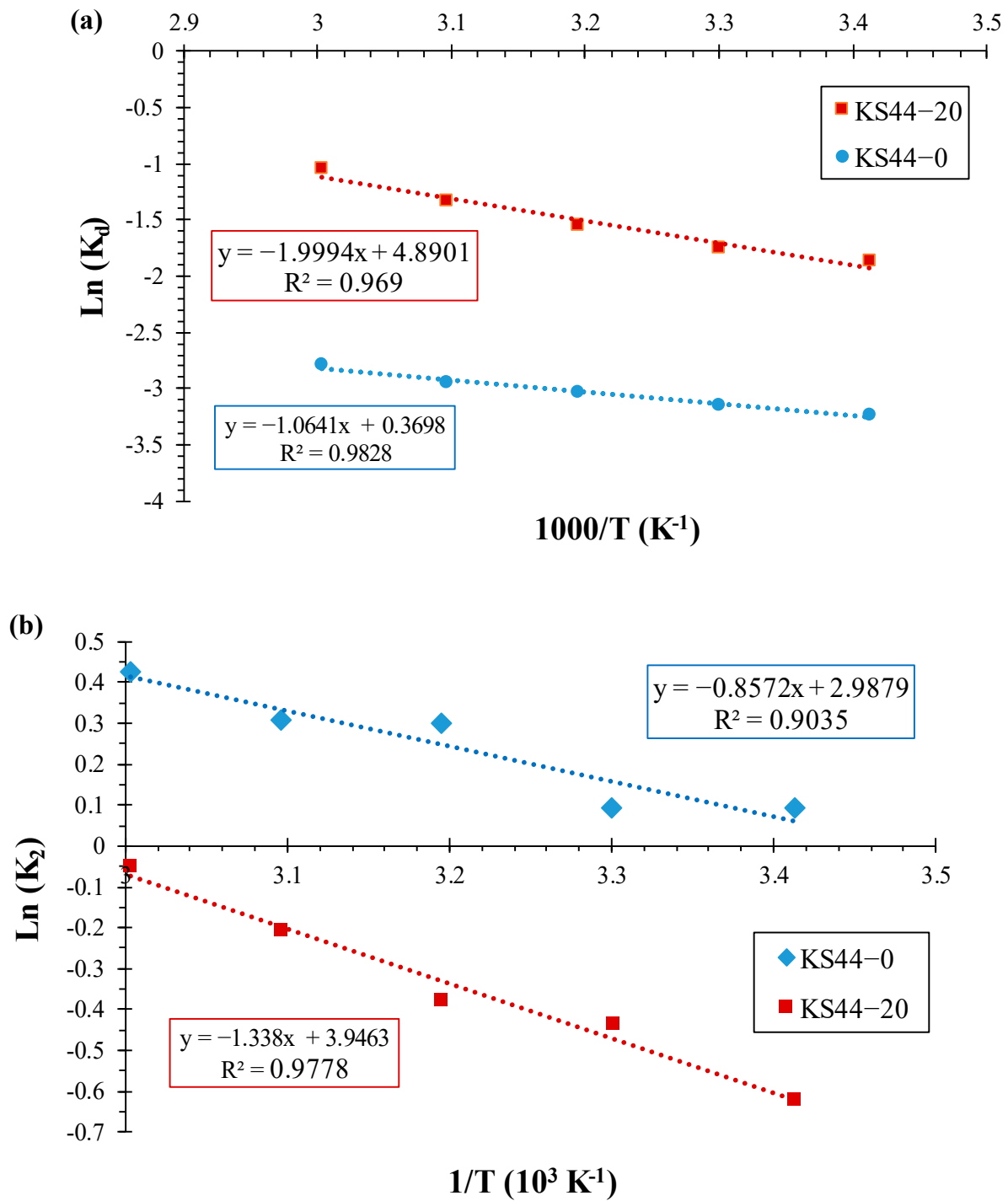


Figure 9. (a) The Van't Hoff plots and (b) the Arrhenius plot for the determination of activation energy for adsorption of RhB on KS44-0 and KS44-20 at different temperatures.

Table 5. Thermodynamics parameters for the adsorption of RhB onto KS44-0 and KS44-20.

Composites	ΔH° ($\text{kJ}\cdot\text{mol}^{-1}$)	ΔS° ($\text{J}\cdot\text{mol}^{-1}$)	ΔG° ($\text{kJ}\cdot\text{mol}^{-1}$)					R^2
			T = 20 °C	T = 30 °C	T = 40 °C	T = 50 °C	T = 60 °C	
KS44-0	9.079	4.010	-1.165	-1.206	-1.246	-1.286	-1.326	0.982
KS44-20	16.747	41.583	-1.158	-1.198	-1.238	-1.278	-1.318	0.969

The influence of temperature on the retention rate of the RhB dye on the adsorbents (KS44-0 and KS44-20) is illustrated in Figure 9a. According to the adsorption theory, the adsorption rate and adsorption capacity increase with the increase in temperature (from 20 °C to 60 °C). In accordance with the first law of thermodynamics, this phenomenon does not only suggest that the reaction of RhB dye adsorption on composites is endothermic ($\Delta H^\circ_{\text{ads}} > 0$), but also explains why reactions occur more rapidly at higher temperatures [38,39].

From Table 5, the positive values of standard enthalpy change ($\Delta H^\circ_{\text{ads}}$) reveal the endothermic nature of the adsorption process that absorbs thermal energy from its surroundings, namely heat transfer into the system.

This is in accordance with the increase in adsorption capacity upon increasing temperature. The heat evolved during physical adsorption is of the same order of magnitude as that of condensation, (i.e., 2.1–20.9 kJ·mol⁻¹), while the heat of chemisorption generally falls into the range of 80–200 kJ·mol⁻¹ [40–43]. Hence, the values of standard enthalpy change ($\Delta H^\circ_{\text{ads}} = 8.846$ and 16.623 kJ·mol⁻¹) were less than 20.9 kJ·mol⁻¹ for KS44-0 and KS44-20, respectively. This indicates that RhB dye adsorption by using carbon composites is attributed to a physical adsorption process rather than a chemical adsorption process. As a result of this process, the second law of thermodynamics states that the positive values of standard entropy change ($\Delta S^\circ_{\text{ads}}$) reflect an increase in disorder at the adsorbate/adsorbent interface in the adsorption process and reveal the binding affinity of the adsorbate to adsorbent materials toward RhB dye molecules. For a full understanding of this alloying effect, the increase in the standard entropy change ($\Delta S^\circ_{\text{ads}}$) can provide a larger number of adsorption sites. It can also process more fixation binding sites, allowing the adsorption of RhB dye, which increases the RhB adsorption capacity as the surface and internal structure of the porous materials become more fully saturated by the adsorbate. The thermodynamic evidence of the viability and spontaneity of the RhB dye adsorption process on KS44-0 and KS44-20 composite materials is provided by the negative values of standard Gibbs free energy change ($\Delta G^\circ_{\text{ads}}$) at different temperatures (from 20 °C to 60 °C).

The expression of Arrhenius' law reveals the activation energy (E_a) for the dye adsorption by using the pseudo-second-order rate constant (Figure 10), as given by Equation (16) [44,45]:

$$\ln(K_2) = \ln(A_0) - \frac{E_a}{RT} \quad (16)$$

where K_2 is the rate constant for the reaction at a given temperature (L·g⁻¹), A_0 is the temperature-independent factor (g·mg⁻¹·min⁻¹), E_a is the activation energy (kJ·mol⁻¹), R is the universal gas constant (8.314 kJ·mol⁻¹) and T is the solution temperature (K).

The activation energy was determined from the slope of the straight line of plotting $\ln K_2$ versus $1/T$ (Figure 9b). If the value of activation energy requirement is small ($E_a = 0$ –40 kJ·mol⁻¹), and since the forces involved are weak, the adsorption process flows by chemical ion exchange. If the activation energy $E_a < 40$ kJ·mol⁻¹, then the adsorption process is physical in nature, whereas if the E_a value is within the range of 40–800 kJ·mol⁻¹, then the adsorption process is chemisorption in nature and involves forces much stronger (chemical bonds) than physical adsorption [46–50]. Therefore, the values of activation energy (E_a) (i.e., minimum energy required for a reaction to occur), calculated using the Arrhenius equation ($E_a = 7.126$ and 11.124 kJ·mol⁻¹), are less than 40 kJ·mol⁻¹ for KS44-0 and KS44-20, respectively (Table 6). This implies that the adsorption process is controlled by a physisorption interaction mechanism, and thus, activation energy requirements are small (below 40 kJ·mol⁻¹). These data were consistent with other results from several investigations, which suggested that the adsorption has a multilayer character and involves weak intermolecular forces [51–54].

It is worthwhile mentioning that RhB dye solution has two primary distinctive absorption bands (Figure 11). The ultraviolet (UV) light spans a range of wavelengths between 200 and 400 nm, which represents the aromatic rings of RhB ($\pi \rightarrow \pi^*$ transition of C=C, C=O and C=N groups). The visible radiation bands at 555 nm ($n \rightarrow \pi^*$ transition of C=N and C=O groups) are attributable to the color of dye solution, which was used for the

monitoring of dye decolorization [55]. The high adsorption performance of the carbon graphite/carbon nanotube composite material towards RhB dye might be explained by the feasible adsorption mechanism involved in the adsorption process (Figure 11). This startling discovery is especially pertinent to the π - π stacking interactions, which corresponds to a very small binding energy between the aromatic backbone of the RhB dye molecules and the large delocalized π -electrons over the entire hexagonal skeleton of carbon graphite and carbon nanotubes [56–59]. Moreover, the observed porosities emanate from the pores available for RhB dye uptake, which further enhances the absorption capacity in the carbon composites. These pores provide excellent surfaces for trapping and thus favor the RhB dye molecules' adsorption [60–62].

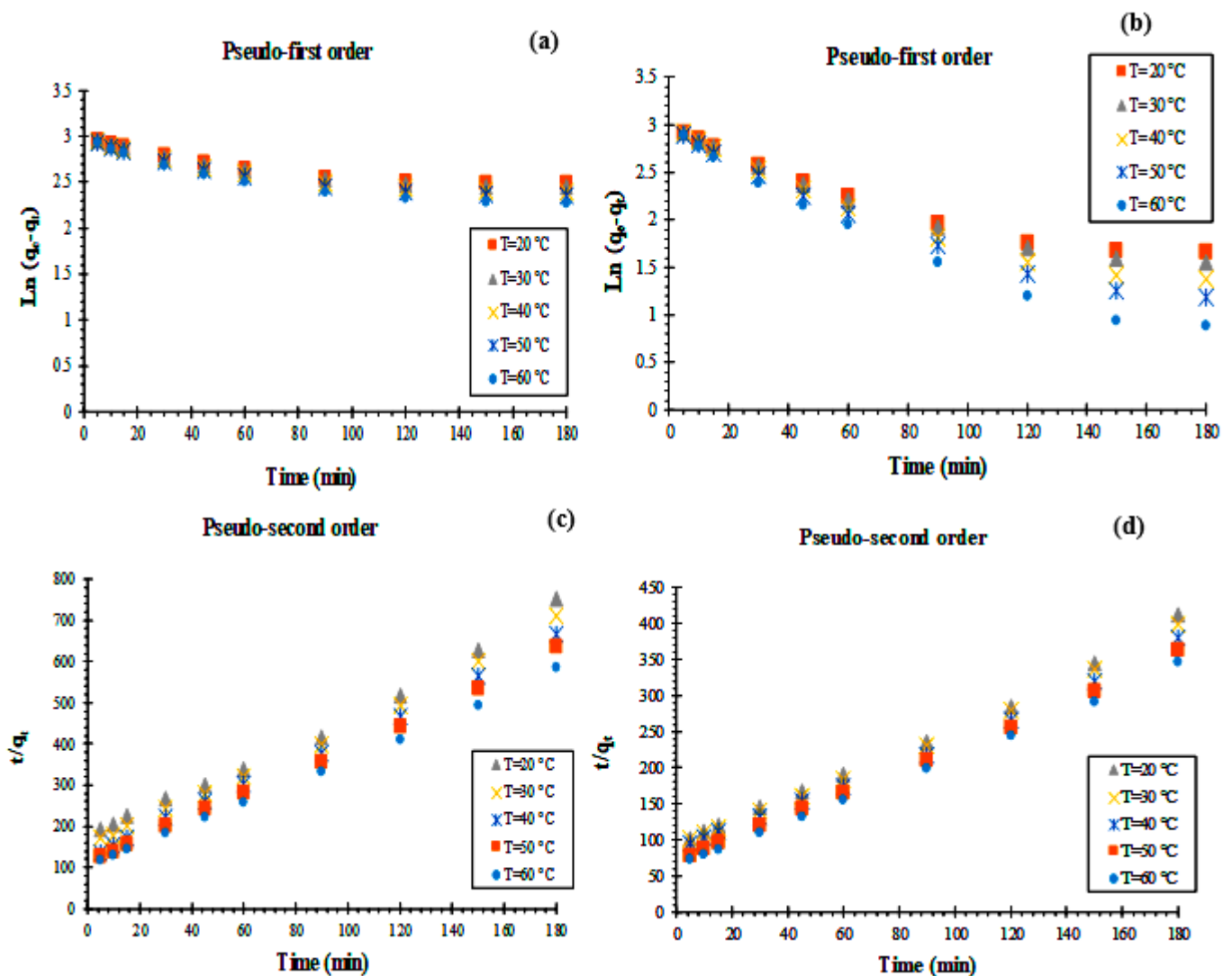


Figure 10. Pseudo-first-order and pseudo-second-order models plot for RhB at different temperatures onto (a,c) KS44-0 and (b,d) KS44-20.

Table 6. Determined activation energies of the adsorption of RhB onto KS44-0 and KS44-20.

Composites	Activation Energy (E_a) (kJ·mol ⁻¹)	R ²
KS44-0	7.126	0.903
KS44-20	11.124	0.977

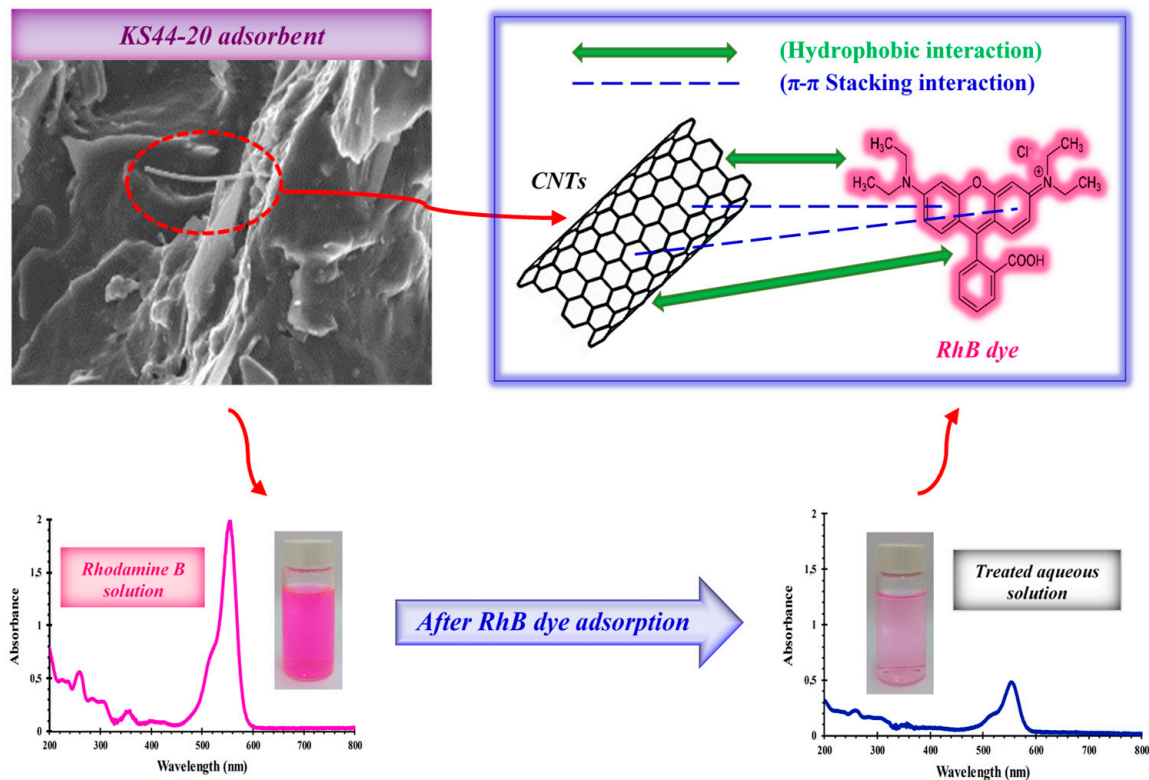


Figure 11. Mechanism of adsorption for the removal of RhB dye onto carbon composites (RhB) = 10 mg.L⁻¹, T = 20 °C, pH = 3 and time = 3 h.

3.4. Comparison of Performances with Other Adsorbents

The results of the adsorption capacities and thermodynamic parameters for the adsorption of RhB dye onto carbon composites were compared with other adsorbents with the same dye reported in the literature, as displayed in Table 7. Many of the reported research works have investigated the removal of RhB using carbon-based materials as adsorbents, such as terpyridine functionalized magnetic nanoparticles (MNP-Tppy), clay–cellulose composites, fly ash, commercial activated carbon (CAC), activated carbon obtained from pericarp of rubber fruit (PrAC), graphene oxide (GO) and multi-walled carbon nanotubes (MWCNTs). However, it is obvious that the adsorption capacities of the as-prepared carbon graphite/carbon nanotubes composite regarding RhB have higher RhB removal abilities than the previously mentioned adsorbents. The above study reveals that KS44-20 is a promising adsorbent for the removal of RhB from aqueous solutions due to the presence of carbon nanotubes, high surface area and high adsorption capacity. Therefore, the carbon graphite/carbon nanotube composite material has great potential for wastewater treatment and has been proven to be better than other adsorbents reported in the literature.

Table 7. Comparison of the Rhodamine B (RhB) dye adsorption capacities and thermodynamic parameters on different carbon adsorbents.

Carbon-Based Materials	Conditions				Q _e (mg.g ⁻¹)	Kinetic Model	Isotherm Model	ΔH [°] _{ads} (kJ.mol ⁻¹)	ΔG [°] _{ads} (kJ.mol ⁻¹)	ΔS [°] _{ads} (kJ.mol ⁻¹)	E _a (kJ.mol ⁻¹)	Ref
	(RhB) (mg.L ⁻¹)	pH	T (°C)	Time (h)								
MNP-Tppy	10	6	20	4	4.742	Pseudo-second-order	Langmuir	0.937	-1.220	3.69	-	[63]
Cellulose-clay	10	2	30	2	2.789	Pseudo-second-order	Redlich–Peterson	0.911	-3055.14	10.086	-	[64]
Fly ash	10	7	30	3	0.334	Pseudo-first-order	Langmuir and Freundlich	-24.97	-9.49	51.05	-	[65]
CAC	150	4	30	12	0.428	Pseudo-second-order	Langmuir	4.01	-	-	-	[8]
PrAC	300	4	30	36	0.110	Pseudo-second-order	Langmuir	14.81	-	-	-	[8]
GO	67	7	25	24	0.043	-	Freundlich	4.18	-1.77	5.95	-	[66]
MWCNTs	10	1	25	0.5	0.148	-	-	-	-	-	-	-
	10	7	25	0.5	0.109	Pseudo-second-order	Langmuir and Freundlich	-	-	-	-	[67]
	10	10	25	0.5	0.127	-	-	-	-	-	-	-
KS44-0	10	3	20	3	0.239	Pseudo-second-order	Freundlich	9.079	-1.165	4.010	7.126	This work
KS44-20	10	3	20	3	0.437	Pseudo-second-order	Freundlich	16.747	-1.158	41.583	11.124	

4. Conclusions

Porous carbon composite materials (KS44-0 and KS44-20) have been used for the absorption of Rhodamine B (RhB) dye from aqueous solution. The removal effectiveness of carbon nanotubes as adsorbents for RhB from aqueous solution, as well as cationic dyes, was described in this study, which aimed to limit the environmental effect of synthetic product dyes in water and wastewater. Eventually, the results showed that the surface characteristics of carbon materials were substantially altered by the addition of ferrocene content, related to the presence of carbon nanotubes (CNTs) within the composite material (KS44-20). The presence of CNTs that grew on the surface and inside the pores gave rise to more active sites and surface area for the interaction between the cationic dye, Rhodamine B (RhB). Hence, this higher adsorption for the adsorbent (KS44-20) relative to (KS44-0) can be explained by the fact that carbon nanotubes have been investigated as interesting and high-performing adsorbent materials to be used in water treatment. The optimization of the adsorption process depends on large parameters (time, pH, temperature, the concentration of the pollutant and the nature of the adsorbent), which govern its effectiveness. In such a way, by keeping the pH constant at 3.0 and varying the solution temperatures at 20 °C, 30 °C, 40 °C, 50 °C and 60 °C, one could find a significant influence on the adsorption process for RhB dye. The best Rhodamine B dye removal effectiveness was found in almond shells (85%), which took place within 3 h for graphite carbon/carbon nanotube composites. The equilibrium adsorption data fitted well to the Freundlich adsorption isotherm, better than the Langmuir isotherm. The pseudo-second-order kinetics model was observed to fit the adsorption data. At all temperatures, negative standard Gibbs free energy change values ($\Delta G^{\circ}_{\text{ads}}$) indicated that the adsorption process was endothermic, spontaneous and practicable. The values of standard enthalpy change ($\Delta H^{\circ}_{\text{ads}}$) and activation energy (E_a) indicated that the adsorption process is physical sorption. As a result, RhB dye is a planar molecule that is readily adsorbed due to the π - π interactions between the dyes' aromatic backbones and the hexagonal skeleton of graphite carbon and carbon nanotubes. Thus, it is believed that carbon composites will play a major role in organic pollutant reduction. The prepared adsorbent efficiency was compared to other adsorbents regarding the same dye removal, and we may conclude that its action is enhanced compared to other adsorbents reported in the literature.

Supplementary Materials: The following supporting information can be downloaded at: <https://www.mdpi.com/article/10.3390/ma16031015/s1>, Figure S1: Calibration curve (at 555 nm) and chemical structure of Rhodamine B dye molecule (RhB).

Author Contributions: Conceptualization, I.J., M.C., S.C. and M.A.; Methodology, S.Z.; Validation, I.J. and M.A.; Investigation, M.C. and S.C.; Writing—original draft, S.Z.; Writing—review & editing, I.J. and M.A.; Visualization, M.C. and S.C.; Supervision, I.J. and M.A.. All authors have read and agreed to the published version of the manuscript.

Funding: This research received no external funding.

Institutional Review Board Statement: Not applicable.

Informed Consent Statement: Not applicable.

Data Availability Statement: The data presented in this study are available within the paper itself and the supplementary material. Any more data needed is available on request from the corresponding author because it is kept in the laboratory note books.

Conflicts of Interest: The authors declare no conflict of interest.

References

1. Pagga, U.; Brown, D. The degradation of dyestuffs. Part II. Behavior of dyestuffs in aerobic biodegradation tests. *Chemosphere* **1986**, *15*, 479–491. [[CrossRef](#)]
2. Chequer, F.D.; De Oliveira, G.R.; Ferraz, E.A.; Cardoso, J.C.; Zanoni, M.B.; De Oliveira, D.P. Textile dyes: Dyeing process and environmental impact. *Eco-Friendly Text. Dye. Finish.* **2013**, *6*, 151–176.
3. DEPA (Danish Environmental Protection Agency). Toxicity and Fate of Azo Dyes. In *Survey of Azo-Colorants in Denmark*; DEPA: Copenhagen, Denmark, 2000.
4. Khadhraoui, M.; Trabelsi, H.; Ksibi, M.; Bouguerra, S.; Elleuch, B. Discoloration and detoxification of a Congo red dye solution by means of ozone treatment for a possible water reuse. *J. Hazard. Mater.* **2009**, *161*, 974–981. [[CrossRef](#)]
5. Abdolrahimi, N.; Tadjarodi, A. Adsorption of rhodamine-B from aqueous solution by activated carbon from almond shell. *Proceedings* **2019**, *41*, 51.
6. Mohammadi, M.; Hassani, A.J.; Mohamed, A.R.; Najafpour, G.D. Removal of rhodamine B from aqueous solution using palm shell-based activated carbon: Adsorption and kinetic studies. *J. Chem. Eng. Data.* **2010**, *55*, 5777–5785. [[CrossRef](#)]
7. Inyinbor, A.A.; Adekola, F.A.; Olatunji, G.A. Adsorption of Rhodamine B dye from aqueous solution on Irvingia gabonensis biomass: Kinetics and thermodynamics studies. *S. Afr. J. Chem.* **2015**, *68*, 115–125. [[CrossRef](#)]
8. Hayeeye, F.; Sattar, M.; Tekasakul, S.; Sirichote, O. Adsorption of Rhodamine B on activated carbon obtained from pericarp of rubber fruit in comparison with the commercial activated carbon. *Songklanakarini J. Sci. Technol.* **2014**, *36*, 177–187.
9. Ganesan, P.; Kamaraj, R.; Sozhan, G.; Vasudevan, S. Oxidized multiwalled carbon nanotubes as adsorbent for the removal of manganese from aqueous solution. *Environ. Sci. Pollut. Res.* **2013**, *20*, 987–996. [[CrossRef](#)]
10. Gandhi, M.R.; Vasudevan, S.; Shibayama, A.; Yamada, M. Graphene and graphene-based composites: A rising star in water purification—a comprehensive overview. *Chem. Select.* **2016**, *1*, 4358–4385. [[CrossRef](#)]
11. Zghal, S.; Jedidi, I.; Cretin, M.; Cerneaux, S.; Abdelmouleh, M. One-step synthesis of highly porous carbon graphite/carbon nanotubes composite by in-situ growth of carbon nanotubes for the removal of humic acid and copper (II) from wastewater. *Diam. Relat. Mater.* **2020**, *101*, 107557. [[CrossRef](#)]
12. Zghal, S.; Jedidi, I.; Cretin, M.; Cerneaux, S.; Cot, D.; Lagerge, S.; Deabate, S.; Abdelmouleh, M. Electroactive adsorbent composites of porous graphite carbon/carbon nanotube for highly efficient organic dye removal. *Korean J. Chem. Eng.* **2022**, *39*, 2239–2251. [[CrossRef](#)]
13. Nazir, R.; Khan, M.; Rehman, R.U.; Shujah, S.; Khan, M.; Ullah, M.; Zada, A.; Mahmood, N.; Ahmad, I. Adsorption of selected azo dyes from an aqueous solution by activated carbon derived from *Monotheca buxifolia* waste seeds. *Soil Water Res.* **2020**, *15*, 166–172. [[CrossRef](#)]
14. Ju, J.W.; Lee, E.J.; Yoon, C.S.; Myung, S.T.; Sun, Y.K. Optimization of layered cathode material with full concentration gradient for lithium-ion batteries. *J. Phys. Chem. C* **2014**, *118*, 175–182. [[CrossRef](#)]
15. Acharya, J.; Kumar, U.; Meikap, B.C. Thermodynamic characterization of adsorption of lead (II) ions on activated carbon developed from tamarind wood from aqueous solution. *S. Afr. J. Chem. Eng.* **2013**, *18*, 70–76.
16. Chen, X. Modeling of experimental adsorption isotherm data. *Information* **2015**, *6*, 14–22. [[CrossRef](#)]
17. Vasudevan, S.; Lakshmi, J. Electrochemical removal of boron from water: Adsorption and thermodynamic studies. *Can. J. Chem. Eng.* **2012**, *90*, 1017–1026. [[CrossRef](#)]
18. Xing, X.; Qu, H.; Shao, R.; Wang, Q.; Xie, H. Mechanism and kinetics of dye desorption from dye-loaded carbon (XC-72) with alcohol-water system as desorbent. *Water Sci. Technol.* **2017**, *76*, 1243–1250. [[CrossRef](#)]
19. Kamaraj, R.; Vasudevan, S. Facile one-pot electrosynthesis of Al(OH)³⁻ kinetics and equilibrium modeling for adsorption of 2, 4, 5-trichlorophenoxyacetic acid from aqueous solution. *New J. Chem.* **2016**, *40*, 2249–2258. [[CrossRef](#)]
20. Kamaraj, R.; Davidson, D.J.; Sozhan, G.; Vasudevan, S. Adsorption of herbicide 2-(2, 4-dichlorophenoxy) propanoic acid by electrochemically generated aluminum hydroxides: An alternative to chemical dosing. *RSC Adv.* **2015**, *5*, 39799–39809. [[CrossRef](#)]
21. Kamaraj, R.; Pandiarajan, A.; Gandhi, M.R.; Shibayama, A.; Vasudevan, S. Eco-friendly and easily prepared graphenenanosheets for safe drinking water: Removal of chlorophenoxyacetic acid herbicides. *ChemistrySelect* **2017**, *2*, 342–355. [[CrossRef](#)]
22. Wu, F.C.; Tseng, R.L.; Juang, R.S. Initial behavior of intraparticle diffusion model used in the description of adsorption kinetics. *Chem. Eng. J.* **2009**, *153*, 1–8. [[CrossRef](#)]
23. Inyinbor, A.A.; Adekola, F.A.; Olatunji, G.A. Kinetics, isotherms and thermodynamic modeling of liquid phase adsorption of Rhodamine B dye onto *Raphia hookeri* fruit epicarp. *Water Resour. Ind.* **2016**, *15*, 14–27. [[CrossRef](#)]
24. Zahir, A.; Aslam, Z.; Aslam, U.; Abdullah, A.; Ali, R.; Bello, M.M. Paspalum notatum grass-waste-based adsorbent for rhodamine B removal from polluted water. *Chem. Biochem. Eng. Q.* **2020**, *34*, 93–104. [[CrossRef](#)]
25. Aljebori, A.M.K.; Alshirifi, A.N. Effect of different parameters on the adsorption of textile dye maxilon blue GRL from aqueous solution by using white marble. *Asian J. Chem.* **2012**, *24*, 5813.
26. Dridi, C.; Youcef, L. Removal of lead by adsorption on kaolin. *Larhyss J.* **2016**, *27*, 7–22.
27. Mittal, A.; Malviya, A.; Kaur, D.; Mittal, J.; Kurup, L. Studies on the adsorption kinetics and isotherms for the removal and recovery of Methyl Orange from wastewaters using waste materials. *J. Hazard. Mater.* **2007**, *148*, 229–240. [[CrossRef](#)]
28. Postai, D.L.; Demarchi, C.A.; Zanatta, F.; Melo, D.C.C.; Rodrigues, C.A. Adsorption of rhodamine B and methylene blue dyes using waste of seeds of *Aleurites Moluccana*, a low cost adsorbent. *Alex. Eng. J.* **2016**, *55*, 1713–1723. [[CrossRef](#)]

29. Saadi, R.; Saadi, Z.; Fazaeli, R.; Fard, N.E. Monolayer and multilayer adsorption isotherm models for sorption from aqueous media. *Korean J. Chem. Eng.* **2015**, *32*, 787–799. [[CrossRef](#)]
30. Rozanov, L.N. Kinetic equations of non-localized physical adsorption in vacuum for Freundlich adsorption isotherm. *Vacuum* **2021**, *189*, 110267. [[CrossRef](#)]
31. Ullah, M.; Nazir, R.; Khan, M.; Khan, W.; Shah, M.; Afridi, S.G.; Zada, A. The effective removal of heavy metals from water by activated carbon adsorbents of Albizia lebeck and Melia azedarach seed shells. *Soil Water Res.* **2020**, *15*, 30–37. [[CrossRef](#)]
32. Alberty, R.A. Standard Gibbs free energy, enthalpy, and entropy changes as a function of pH and pMg for several reactions involving adenosine phosphates. *J. Boil. Chem.* **1969**, *244*, 3290–3302. [[CrossRef](#)]
33. Maimaitiniyazi, R.; Zhou, X. Comment on the calculation of equilibrium constant and thermodynamic parameters by the distribution coefficient in Environmental Research 203 2022 111814. *Environ. Res.* **2022**, *210*, 112958. [[CrossRef](#)]
34. Vasudevan, S.; Sheela, S.M.; Lakshmi, J.; Sozhan, G. Optimization of the process parameters for the removal of boron from drinking water by electrocoagulation—a clean technology. *J. Chem. Technol. Biotechnol.* **2010**, *85*, 926–933. [[CrossRef](#)]
35. Lima, E.C.; Gomes, A.A.; Tran, H.N. Comparison of the nonlinear and linear forms of the van't Hoff equation for calculation of adsorption thermodynamic parameters (ΔS° and ΔH°). *J. Mol. Liq.* **2020**, *311*, 113315. [[CrossRef](#)]
36. Tran, H.N.; Lima, E.C.; Juang, R.S.; Bollinger, J.C.; Chao, H.P. Thermodynamic parameters of liquid–phase adsorption process calculated from different equilibrium constants related to adsorption isotherms: A comparison study. *J. Environ. Chem. Eng.* **2021**, *9*, 106674. [[CrossRef](#)]
37. Kamaraj, R.; Pandiarajan, A.; Jayakiruba, S.; Naushad, M.; Vasudevan, S. Kinetics, thermodynamics and isotherm modeling for removal of nitrate from liquids by facile one-pot electrosynthesized nano zinc hydroxide. *J. Mol. Liq.* **2016**, *215*, 204–211. [[CrossRef](#)]
38. Laidler, K.J. The development of the Arrhenius equation. *J. Chem. Educ.* **1984**, *61*, 494. [[CrossRef](#)]
39. Kyzas, G.Z.; McKay, G.; Al-Musawi, T.J.; Salehi, S.; Balarak, D. Removal of Benzene and Toluene from Synthetic Wastewater by Adsorption onto Magnetic Zeolitic Imidazole Framework Nanocomposites. *Nanomaterials* **2022**, *12*, 3049. [[CrossRef](#)]
40. Saha, P.; Chowdhury, S. Insight into adsorption thermodynamics. *Thermodynamics* **2011**, *16*, 349–364.
41. Chowdhury, S.; Mishra, R.; Kushwaha, P.; Saha, P. Removal of safranin from aqueous solutions by NaOH-treated rice husk: Thermodynamics, kinetics and isosteric heat of adsorption. *Asia-Pac. J. Chem. Eng.* **2012**, *7*, 236–249. [[CrossRef](#)]
42. Enoh, B.S.; Christopher, W. Adsorption of metal ions from carwash wastewater by phosphoric acid modified clay: Kinetics and thermodynamic studies. *Adsorption* **2015**, *7*, 1–8.
43. Mahmoud, M.A.; El-Halwany, M.M. Adsorption of cadmium onto orange peels: Isotherms, kinetics, and thermodynamics. *J. Chromatogr. Sep. Tech.* **2014**, *5*, 1000238. [[CrossRef](#)]
44. Kulczycki, A.; Kajdas, C. A new attempt to better understand Arrhenius equation and its activation energy. In *Tribology in Engineering*; 2013; pp. 47–62.
45. Kalaivanan, R.; Ganesh, N.V.; Al-Mdallal, Q.M. An investigation on Arrhenius activation energy of second grade nanofluid flow with active and passive control of nanomaterials. *Case Stud. Therm. Eng.* **2020**, *22*, 100774. [[CrossRef](#)]
46. Mahmoud, M.A. Kinetics and thermodynamics of aluminum oxide nanopowder as adsorbent for Fe (III) from aqueous solution. *Beni-Suef Univ. J. Basic Appl. Sci.* **2015**, *4*, 142–149. [[CrossRef](#)]
47. Erhayem, M.; Gaith, R.; Otman, O.E.; Frage, M.U. Adsorption, kinetic and thermodynamic study for removal of nickel ions by activated carbon from palm kernel. *Iran. J. Energy Environ.* **2020**, *11*, 339–350.
48. Nnaji, C.C.; Agim, A.E.; Mama, C.N.; Emenike, P.C.; Ogarekpe, N.M. Equilibrium and thermodynamic investigation of biosorption of nickel from water by activated carbon made from palm kernel chaff. *Sci. Rep.* **2021**, *11*, 1–20. [[CrossRef](#)]
49. Al-Harby, N.F.; Albahly, E.F.; Mohamed, N.A. Kinetics, Isotherm and Thermodynamic Studies for Efficient Adsorption of Congo Red Dye from Aqueous Solution onto Novel Cyanoguanidine-Modified Chitosan Adsorbent. *Polymers* **2021**, *13*, 4446. [[CrossRef](#)]
50. Hasani, N.; Selimi, T.; Mele, A.; Thaçi, V.; Halili, J.; Berisha, A.; Sadiku, M. Theoretical, Equilibrium, Kinetics and Thermodynamic Investigations of Methylene Blue Adsorption onto Lignite Coal. *Molecules* **2022**, *27*, 1856. [[CrossRef](#)]
51. Cantu, Y.; Remes, A.; Reyna, A.; Martinez, D.; Villarreal, J.; Ramos, H.; Trevino, S.; Tamez, C.; Martinez, A.; Eubanks, T. Thermodynamics, kinetics, and activation energy studies of the sorption of chromium (III) and chromium (VI) to a Mn₃O₄ nanomaterial. *Chem. Eng. J.* **2014**, *254*, 374–383. [[CrossRef](#)]
52. Rodrigues, L.A.; Maschio, L.J.; Evangelista da Silva, R.; Pinto da Silva, M.L.C. Adsorption of Cr (VI) from aqueous solution by hydrous zirconium oxide. *J. Hazard. Mater.* **2010**, *173*, 630–636. [[CrossRef](#)] [[PubMed](#)]
53. Musorrafiti, M.J.; Konek, C.T.; Hayes, P.L.; Geiger, F.M. Interaction of Chromium(VI) with the α -Aluminum Oxide-Water Interface. *J. Phys. Chem. C.* **2008**, *112*, 2032–2039. [[CrossRef](#)]
54. Boparai, H.K.; Joseph, M.; O'Carroll, D.M. Kinetics and thermodynamics of cadmium ion removal by adsorption onto nanozevalent iron particles. *J. Hazard. Mater.* **2011**, *186*, 458–465. [[CrossRef](#)]
55. Chen, X.; Xue, Z.; Yao, Y.; Wang, W.; Zhu, F.; Hong, C. Oxidation degradation of rhodamine B in aqueous by treatment system. *Int. J. Photoenergy* **2012**, *2012*, 1–5.
56. Rabti, A.; Hannachi, A.; Maghraoui-Meherzi, H.; Raouafi, N. Ferrocene-functionalized carbon nanotubes: An adsorbent for Rhodamine B. *Chem. Afr.* **2019**, *2*, 113–122. [[CrossRef](#)]
57. Koh, B.; Kim, G.; Yoon, H.K.; Park, J.B.; Kopelman, R.; Cheng, W. Fluorophore and dye-assisted dispersion of carbon nanotubes in aqueous solution. *Langmuir* **2012**, *28*, 11676–11686. [[CrossRef](#)]

58. Tournus, F.; Latil, S.; Heggie, M.I.; Charlier, J.C. π -stacking interaction between carbon nanotubes and organic molecules. *Phys. Rev. B* **2005**, *72*, 075431. [[CrossRef](#)]
59. Zhu, X.; Zhang, K.; Lu, N.; Yuan, X. Simultaneous determination of 2, 4, 6-trichlorophenol and pentachlorophenol based on poly (Rhodamine B)/graphene oxide/multiwalled carbon nanotubes composite film modified electrode. *Appl. Surf. Sci.* **2016**, *361*, 72–79. [[CrossRef](#)]
60. Bello, O.S.; Ahmad, M.A. Adsorptive removal of a synthetic textile dye using cocoa podhusks. *Toxicol. Environ. Chem.* **2011**, *93*, 1298–1308. [[CrossRef](#)]
61. Bello, O.S.; Alabi, E.O.; Adegoke, K.A.; Adegboyega, S.A.; Inyinbor, A.A.; Dada, A.O. Rhodamine B dye sequestration using Gmelina aborea leaf powder. *Heliyon* **2020**, *6*, e02872. [[CrossRef](#)]
62. Bello, O.S.; Adegoke, K.A.; Inyinbor, A.A.; Dada, A.O. Trapping Rhodamine B dye using functionalized mango (*Mangifera indica*) pod. *Water Environ. Res.* **2021**, *93*, 2308–2328. [[CrossRef](#)]
63. Ojemaye, M.O.; Okoh, A.I. Multiple nitrogen functionalized magnetic nanoparticles as an efficient adsorbent: Synthesis, kinetics, isotherm and thermodynamic studies for the removal of rhodamine B from aqueous solution. *Sci. Rep.* **2019**, *9*, 1–13. [[CrossRef](#)]
64. Kausar, A.; Shahzad, R.; Asim, S.; BiBi, S.; Iqbal, J.; Muhammad, N.; Sillanpaa, M.; Din, I.U. Experimental and theoretical studies of Rhodamine B direct dye sorption onto clay-cellulose composite. *J. Mol. Liq.* **2021**, *328*, 115165. [[CrossRef](#)]
65. Khan, T.A.; Ali, I.; Singh, V.V.; Sharma, S. Utilization of fly ash as low-cost adsorbent for the removal of methylene blue, malachite green and rhodamine B dyes from textile wastewater. *J. Environ. Prot. Sci.* **2009**, *3*, 11–22.
66. Sahara, J.; Naem, A.; Farooq, M.; Zareen, S.; Rahman, A.U. Thermodynamic studies of adsorption of rhodamine B and Congo red on graphene oxide. *Desalin. Water Treat.* **2019**, *164*, 228–239. [[CrossRef](#)]
67. Kumar, S.; Bhanjana, G.; Jangra, K.; Dilbaghi, N.; Umar, A. Utilization of carbon nanotubes for the removal of rhodamine B dye from aqueous solutions. *J. Nanosci. Nanotechnol.* **2014**, *14*, 4331–4336. [[CrossRef](#)]

Disclaimer/Publisher's Note: The statements, opinions and data contained in all publications are solely those of the individual author(s) and contributor(s) and not of MDPI and/or the editor(s). MDPI and/or the editor(s) disclaim responsibility for any injury to people or property resulting from any ideas, methods, instructions or products referred to in the content.