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Electrochemical redox processes involved in carbon-coated KVPO₄F for high voltage K-ion batteries revealed by XPS analysis

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Abstract

K-ion batteries need efficient positive electrode materials with open structural frameworks to accommodate the large ionic radius of K⁺. In that direction, polyanionic compounds are of great interest. Among them, KVPO₄F is the most studied one. Its electrochemical redox processes still remain unclear, especially at high voltage. To tackle this issue, X-ray photoelectron spectroscopy was used to reveal the electrochemical redox processes of K_xVPO₄F. First, a carbon coating was performed and allowed increasing the overall electrochemical performance while mitigating electrolyte degradation at high voltage. Then, XPS analysis showed a high reversibility of the redox processes although the K⁺ extraction (and insertion) from x=0.5 to 0

was hindered, possibly by structural constraints while electrolyte degradation occurred mostly above 4.5 V.

Introduction

The development of new electrochemical energy storage systems as alternatives to the Li-ion technology remains a key challenge due to the growth of energy storage demand. Alternative batteries based on abundant and thus low cost raw materials are required. Na-ion and K-ion batteries have attracted increasing interest during the last decade. Thanks to the low potential of the K^+/K redox couple versus the other alkali metals in non-aqueous solvents, high voltage K-ion batteries (KIBs) could be developed.¹ Moreover, K^+ has the lowest Lewis acidity and desolvation energy compared to Na^+ and Li^+ ,² and thus higher ionic conductivity and faster electrode/electrolyte interface diffusion kinetics that would allow developing high power KIBs.

The high reactivity of K metal requires considering results obtained in half-cells with caution, especially the interaction between the electrolyte and the electrode surface. Indeed, migration of electrolyte degradation products can occur from the K metal to the working electrode.³ Among negative electrode materials,⁴ graphite is of great interest as it reversibly intercalates K^+ to form KC_8 .^{5,6,7} Promising electrochemical performance is obtained, with a reversible capacity of $\sim 250 \text{ mAh.g}^{-1}$ up to hundreds of cycles and even for high practical loading ($> 20 \text{ mg/cm}^2$).^{7,8,9,10,11,12,13} Regarding the safety issue, the average discharge potential of $\sim 0.3 \text{ V vs. } K^+/K$ makes graphite even more attractive for the application.

In fact, the major drawback of K-ion batteries is the need of positive electrode materials with open structural frameworks to accommodate the large ionic radius of K^+ (1.38 \AA) compared to Na^+ (1.02 \AA) and Li^+ (0.76 \AA),^{1,14} which is detrimental to the volumetric energy density.¹⁵ Polyanionic compounds are of great interest as they are suitable host

structures for the reversible K^+ intercalation as well as to deliver high voltage.¹⁶ For instance, $KVPO_4F$ (KVPF) showed reversible capacity up to $\sim 105 \text{ mAh.g}^{-1}$ with an average discharge potential of $\sim 4.3 \text{ V vs. } K^+/K$ and excellent rate performance.^{17,18,19,20,21,22} In a full-cell configuration using graphite, the energy density has even been evaluated at 300 Wh.kg^{-1} , suggesting a promising material for further full-cell studies.²³ Regarding its electrochemical behaviour as a positive electrode material, KVPF shows 4 distinct potential domains in both charge and discharge, corresponding to successive phase transitions with an interesting limited lattice volume change of $\sim 6.5 \%$,¹⁹ and the formation of intermediate phases K_xVPO_4F at $x=0.75$, $x=0.625$ and $x=0.5$. The reaction occurring at high voltage, between the composition $K_{0.5}VPO_4F$ and VPO_4F still remains unclear, however, with simultaneous active material redox reaction and large electrolyte degradation. It may explain why the theoretical capacity of KVPF (131 mAh.g^{-1}) was never reached so far. Therefore, it is of great interest to better understand the electrochemical processes involved in KVPF positive electrode materials in KIBs in order to improve its performance.

The present work proposes to investigate the electrochemical redox processes involved in KVPF upon cycling in K metal batteries using X-ray photoelectron spectroscopy (XPS). The impact of a carbon-coating on the electrochemical performance of KVPF was first evaluated, regarding its ability to improve the overall performance and mitigate the electrolyte degradation.

Experimental

$KVPO_4F$ and carbon coated $KVPO_4F$ ($KVPO_4F\text{-C}$) were prepared by a two-step reaction. First, VPO_4 was obtained by mixing V_2O_5 (from Sigma-Aldrich, 99.6%) and $NH_4H_2PO_4$ (from Sigma-Aldrich, 99.99%) by high energy ball-milling (SPEX 8000M mixer/mill at 1425 rpm) for 1hr followed by a thermal treatment at 800°C for 7h under

Ar/H₂ (5%). Second, stoichiometric amounts of KF·2H₂O (from Alfa Aesar, 98.5%, dried overnight under vacuum at 250°C and stored in a glovebox under Ar) and VPO₄ were mixed without (KVPF) or with (KVPF-C) the addition of 10 wt.% sucrose (from Sigma-Aldrich, ACS reagent). The mixture was obtained by SPEX ball-milling for 1hr, pelletized and thermal treated at 650°C for 8h under Ar. Rietveld refinement of the X-ray diffraction patterns was performed using the FullProf software and considering an orthorhombic unit cell described in the Pna2₁ space group.²⁴ It confirmed the nature of the grey (KVPF) and black (KVPF-C) powders obtained (**Fig. S1**). The smaller cell parameters and especially volume cell obtained for KVPO₄F-C versus the KVPO₄F phase reported and characterized in-depth by Fedotov *et al.*¹⁷ suggest a partial O²⁻ substitution for F⁻ and thus the formation of KVPO₄F_{1-y}O_y. y was estimated to 0.1 from the study of the solid solution KVPO₄F - KVPO₄O (**Table S1**, and from ³¹P MAS NMR data not reported here). Note that the stoichiometry in K/V/P was determined by chemical analyses (**Fig. S1**). In addition, thermogravimetric analyses and scanning electron microscopy images confirmed the formation of a carbon coating (2.5 wt.% of C) for KVPF-C compared to KVPF (**Fig. S2**). Electrodes were prepared by mixing KVPF or KVPF-C, carbon black (Super P, BET=62 m²/g, Alfa Aesar) and poly(vinylidene fluoride) according to a 70:25:5 weight ratio, in N-methyl-2-pyrrolidone (32:68 solvent:material weight ratio), by ball-milling for 1h. The obtained slurry was cast on an aluminium current collector, dried for 24h under Ar, and finally electrodes of 1.5 mg_{KVPO₄F}/cm² were punch out and dried under vacuum at 80°C for 12h. 2032 coin-type cells (316L stainless steel) were assembled under Ar using a Whatman (grade GF/D) and a polypropylene membrane (Celgard) as separators, K metal (99.95%, Alfa Aesar) as negative electrode and 100 μL of 0.8M KPF₆ in a mixture of EC:DEC by volume as electrolyte (KPF₆, Sigma-Aldrich, ≥99% ; EC, Sigma-Aldrich, anhydrous,

99% ; DEC, Sigma-Aldrich, anhydrous, $\geq 99\%$). All potentials reported in the text are referred to K^+/K . Note that K metal polarization in symmetrical $K//K$ cell was found negligible during open circuit voltage (< 6 mV, **Fig. S3**) compared to the literature.²⁵ During plating/stripping at 0.04 mA/cm² (corresponding to the current used to cycle $K//KVPF$ cells), however, a polarization of 0.1 V was measured. Cycling tests were performed at $C/5$ (26 mA.g⁻¹, *i.e.* the exchange of $1K^+$ in 5 hours) and 20°C using a VMP3 (Biologic, SAS France). Two formation cycles were first performed between 3.5 - 5 V *vs.* K^+/K to minimize electrolyte degradation impact (**Fig. S4 and Fig. 3**), as discussed later. $KVPF$ -C cells were then stopped at different potentials during the 3rd cycle and electrodes were recovered from the cells and washed twice with 1 mL of DEC during 30 s. XPS analysis was performed using an Escalab 250 Xi spectrometer, with a monochromatized Al $K\alpha$ radiation ($h\nu = 1486.6$ eV). Electrodes were placed on a sample holder using uPVC insulation tape (3M part number 655), and transferred to Ar filled glove box connected to the spectrometer. Using the standard charge compensation mode, core spectra were recorded with a 0.15 eV step and a constant 20 eV pass energy. Using CasaXPS software, the binding energy scale was calibrated from the $KVPF$ oxygen peak at 530.0 eV. It is reminded that XPS probes about 5 to 10 nm depth of the sample surface, and here between 6 - 7 nm for the vanadium, *i.e.* not only the extreme surface atomic layers (0.5 - 1 nm) but also the deeper ones, which means a representative part of the bulk is also probed (**Fig. S5**). Moreover, at the extreme surface, the presence of defects may lead to vanadium with an oxidation state higher or lower than in the bulk. $V 2p_{3/2}$ core spectra of $KVPF$ electrodes were thus fitted using peak positions and line shapes constraints corresponding to different oxidation states as determined from reference samples with V^{n+} in a PO_4 environment (**Fig. S5 and Table S2**) in agreement with the literature.^{26,27,28,29,30} This thorough procedure allows reliably following the evolution of

the mean oxidation state upon cycling, especially considering that the mean oxidation state obtained by XPS (3.1) for the pristine KVPF-C material is in good agreement with the XRD analysis ($\text{KVPO}_4\text{F}_{1-y}\text{O}_y$ with $y=0.1$).

Results and discussion

The impact of the carbon coating was evaluated regarding its ability to optimize the electrochemical performance of the material in KIBs and to mitigate the KVPF active material surface reactivity against the electrolyte. Overall, both KVPF and KVPF-C showed a similar charge/discharge profile at C/5 with 4 potential domains (**Fig. S4**), in very good agreement with the literature.¹⁹ However KVPF-C showed a lower average polarization of ~230 mV compared to ~300 mV for KVPF (**Fig. S4**) and a better capacity retention after 70 cycles compared to KVPF (76 % vs 62 %, (**Fig. S6**)). For both systems, a significant high-potential hysteresis is observed, suggesting that the main capacity loss would be related to the high-potential domain. Since KVPF and KVPF-C have close particles size (**Fig. S2**), the carbon coating was shown – as expected – to promote better transport properties within the positive electrode, lower polarization of the cell, smaller electrolyte degradation at high voltage and thus optimized performances. XPS analysis confirmed a significantly lower solid electrolyte interphase coverage at the KVPF-C surface compared to KVFP after the first cycle (**Table S3**). Note that the lower electrolyte degradation is explained by a passivation effect of the carbon coating at the KVPF surface. Based on these results, KVPF-C was therefore selected to study the electrochemical redox processes involved in KVPO_4F upon cycling.

Fig. 1a shows the typical 3rd charge/discharge cycle obtained for K//KVPF-C cells and indicates at which potentials the prepared cells were stopped to perform the analysis of

the redox processes involved in KVPF-C. Note that charge/discharge profiles obtained for all the cells showed very close polarization and capacities (**Fig. S7**), highlighting the great reproducibility of the electrochemical results so that reliable XPS analysis is expected on materials recovered *ex situ* from this series of cells. **Fig. 1b** shows the V 2p_{3/2} XPS core spectra of K_xVⁿ⁺PO₄F-C electrodes as recovered from K//KVPF-C cells during the third cycle at C/5, as function of the K content. For more information, the full V 2p core spectra can be found in **Fig. S8**. XPS analysis revealed that the pristine KVPF-C showed the presence of V²⁺, V³⁺ and V⁴⁺ (**Fig. 1b**) although only V³⁺ and V⁴⁺ were expected based on the KVPO₄F_{0.9}O_{0.1}-C formula. Note that V²⁺ was also observed for the VPO₄ precursor (**Fig. S5**). However, considering the very low amount (< 1 wt.%) of VPO₄ impurity in the KVPF-C material (**Fig. S1**), the presence of V²⁺ is likely to originate from the KVPF-C extreme surface. This was supported by the decrease of the V²⁺ contribution during oxidation as discussed thereafter.

After charge of the cell to 4.10 V, *i.e.* during the 1st oxidative potential domain, the peak associated with V⁴⁺ increased while the V³⁺ and V²⁺ ones decreased (respectively +7%, -4% and -3%). At almost the end of the 2nd potential domain at 4.30 V (*i.e.* at the composition of about K_{0.625}VPO₄F), the peak associated with V⁴⁺ increased while the peaks associated with V³⁺ and V²⁺ decreased (respectively +15%, -12% and -3%). At the end of the 3rd potential domain at 4.46 V (*i.e.* for approx. K_{0.5}VPO₄F), a new peak, attributed to V⁵⁺ appeared (+13%) while the V⁴⁺, V³⁺ and V²⁺ peaks decreased (-11%, -1% and -1%, respectively). This phenomenon further continued after the 4th potential domain at 5.00 V, so that the V⁵⁺, V⁴⁺, V³⁺ and V²⁺ peaks represent 20%, 50%, 23% and 7%, respectively. Note that despite the remaining V³⁺ and V²⁺ even at the end of charge, the corresponding mean oxidation states were in good agreement with the theoretical/electrochemical ones, as discussed thereafter. Note that remaining of V³⁺ and

V^{2+} at the end of charge should arise from specific vanadium environment at the KVPF surface. During discharge, opposite phenomena were observed (**Fig. 1b**), indicating a high reversibility of the redox processes involved in KVPF-C upon cycling in the potential window 3.5-5V vs. K^+/K .

The evolution of the vanadium mean oxidation state as the K content changed during the 3rd cycle for $KVPO_4F_{0.9}O_{0.1}C$ and determined from both the charge/discharge capacity (DO_{elec}) and XPS analysis (DO_{xps}) is reported in **Fig. 2**.

Until the end of the 3rd potential domain, *i.e.* at 4.46 V, a very good agreement was observed between the vanadium mean oxidation states determined from the electrochemistry and from XPS analysis. On the contrary, at the end of the 4th potential domain, *i.e.* at 5.00 V, DO_{elec} was +3.9 while DO_{xps} was only +3.8. This discrepancy indicates that about 25% of the electrons exchanged in the last potential domain are not related to K^+ deintercalation from KVPF-C but to electrolyte degradation. This is in agreement with the high irreversible capacity observed for the 4th potential domain between the charge and the next discharge (**Fig. 1a**). Moreover, the evolution of the capacities associated with the different potential domains for the first 20 cycles showed that electrolyte degradation is mainly associated with the 4th potential domain as shown in **Fig. 3** by the large difference in height between the full and hatched black bars. Nevertheless, during the first two cycles, irreversibility was observed throughout the reaction, the latter being much more pronounced at high voltage (about 60%, **Fig. 3**). During the next cycles, a stabilization was observed with an excellent efficiency below 4.6 V (same height for the full and hatched blue and green bars), and thus a capacity loss almost exclusively associated with the high voltage domain (about 90%) and the composition range $K_{0.5}VPO_4F-C - VPO_4F-C$. This highlights the relevance of using the

3rd cycle to perform the redox processes study instead of the 1st cycle. During discharge, DO_{xps} decreased back to +3.1 (**Fig. 2**), indicating a complete reversibility of the redox processes. This result fully supports the limitation of the participation of the electrolyte degradation to high voltage, only after the 2nd cycle as shown in **Fig. 3**.

Importantly, only 0.7 K^+ were extracted from $KVPO_4F_{0.9}O_{0.1}C$ (**Fig. 2**) instead of 0.9 theoretically. Especially, the capacity associated with the 1st, 2nd and 3rd potential domains nearly reached the theoretical ones and remained almost constant from cycles 3 to 20 (**Fig. 3**). On the contrary, the 4th potential domain capacity continuously decreased to about half the theoretical one from cycle 3 to 20 (**Fig. 3**). Interestingly, no polarization increase (due to the accumulation of electrolyte degradation products) was observed (**Fig. S9**). Thus, the 4th potential domain capacity decrease is not explained by the cut-off potential being reached before the end of the domain. Instead, these results likely originate from structural constraints hindering the extraction (and insertion) of K^+ from $x=0.5$ to $x=0$ as proposed previously.^{19,31} To evaluate the possibility to further extract the remaining K^+ , a cell was hold at 5 V for 18h during the 3rd cycle and the corresponding DO_{xps} was +3.9 (**Fig. S10**), in agreement with the evolution of the cell parameters determined by XRD (not shown). This result highlights that despite the electrolyte degradation process, it is still possible to extract the potassium but the process is much slower, more likely due to a structural constraint such as cell volume reduction. This structural constraint could increase the K^+ migration barrier, inducing a decrease in the K^+ diffusion coefficient,³¹ slowing the global extraction process.

Conclusion

In conclusion, the carbon coating of the KVPF material was found beneficial to the overall electrochemical performance while mitigating electrolyte degradation at high

voltage. Then, by comparing the average oxidation state derived from electrochemical and XPS analysis, the electrochemical redox processes of $K_xVPO_{4.1}F_{0.9}-C$ were revealed. Interestingly, a high reversibility was found although the K^+ extraction (and insertion) from $x=0.5$ to $x=0$ was hindered, possibly by structural constraints. Moreover, the severe electrolyte degradation issue, observed mostly above 4.5 V, is currently under further investigation, as well as possible KVPF dissolution in order to successfully mitigate these irreversible degradation reactions using electrolyte additives and enable highly efficient KVPF based full K-ion cells.

Conflicts of interest

There are no conflicts of interest to declare.

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References

1. K. Kubota, M. Dahbi, T. Hosaka, S. Kumakura, and S. Komaba, *Chem. Rec.*, **18**, 1–22 (2018).
2. T. A. Pham, K. E. Kweon, A. Samanta, V. Lordi, and J. E. Pask, *J. Phys. Chem. C*, **121**, 21913–21920 (2017).
3. L. Madec, V. Gabaudan, G. Gachot, L. Stievano, L. Monconduit, and H. Martinez, *ACS Appl. Mater. Interfaces*, **10**, 34116–34122 (2018).
4. V. Gabaudan, L. Monconduit, L. Stievano, and R. Berthelot, *Front. Energy Res.*, **7** (2019).

5. Z. Jian, W. Luo, and X. Ji, *J. Am. Chem. Soc.*, **137**, 11566–11569 (2015).
6. J. C. Pramudita, V. K. Peterson, J. A. Kimpton, and N. Sharma, *Powder Diffr.*, **32**, S43–S48 (2017).
7. L. Fan, R. Ma, Q. Zhang, X. Jia, and B. Lu, *Angew. Chemie - Int. Ed.*, **58**, 10500–10505 (2019).
8. S. Komaba, T. Hasegawa, M. Dahbi, and K. Kubota, *Electrochem. commun.*, **60**, 172–175 (2015) <http://dx.doi.org/10.1016/j.elecom.2015.09.002>.
9. W. Luo, J. Wan, B. Ozdemir, W. Bao, Y. Chen, J. Dai, H. Lin, Y. Xu, F. Gu, V. Barone, and L. Hu, *Nano Lett.*, **15**, 7671–7677 (2015).
10. Y. An, H. Fei, G. Zeng, L. Ci, B. Xi, S. Xiong, and J. Feng, *J. Power Sources*, **378**, 66–72 (2018) <https://doi.org/10.1016/j.jpowsour.2017.12.033>.
11. X. Niu, L. Li, J. Qiu, J. Yang, J. Huang, Z. Wu, J. Zou, C. Jiang, J. Gao, and L. Wang, *Solid State Ionics*, **341**, 2–7 (2019).
12. R. A. Adams, A. Varma, and V. G. Pol, *J. Power Sources*, **410–411**, 124–131 (2019) <https://doi.org/10.1016/j.jpowsour.2018.11.007>.
13. M. Carboni, A. J. Naylor, M. Valvo, and R. Younesi, *RSC Adv.*, **9**, 21070–21074 (2019) <http://dx.doi.org/10.1039/C9RA01931F>.
14. C. Vaalma, D. Buchholz, and S. Passerini, *Curr. Opin. Electrochem.*, **9**, 41–48 (2018).
15. X. Wu, D. P. Leonard, and X. Ji, *Chem. Mater.*, **29**, 5031–5042 (2017).
16. T. Hosaka, T. Shimamura, K. Kubota, and S. Komaba, *Chem. Rec.*, **19**, 735–745 (2019).
17. S. S. Fedotov, N. R. Khasanova, A. S. Samarin, O. A. Drozhzhin, D. Batuk, O. M. Karakulina, J. Hadermann, A. M. Abakumov, and E. V. Antipov, *Chem. Mater.*, **28**, 411–415 (2016).
18. K. Chihara, A. Katogi, K. Kubota, and S. Komaba, *Chem. Commun.*, **53**, 5208–5211 (2017).
19. H. Kim, D. H. Seo, M. Bianchini, R. J. Clément, H. Kim, J. C. Kim, Y. Tian, T. Shi, W. S. Yoon, and G. Ceder, *Adv. Energy Mater.*, **8**, 1–12 (2018).
20. H. Kim, Y. Ishado, Y. Tian, and G. Ceder, *Adv. Funct. Mater.*, **29**, 1–8 (2019).
21. J. Liao, Q. Hu, X. He, J. Mu, J. Wang, and C. Chen, *J. Power Sources*, **451**, 1–8 (2020).
22. Z. Liu, J. Wang, and B. Lu, *Sci. Bull.* (2020) <https://doi.org/10.1016/j.scib.2020.04.010>.
23. T. Hosaka, K. Kubota, A. S. Hameed, and S. Komaba, *Chem. Rev.* (2020).
24. Juan Rodriguez-Carvajal, *Phys. B*, **192**, 55–69 (1993).
25. T. Hosaka, S. Muratsubaki, K. Kubota, H. Onuma, and S. Komaba, *J. Phys. Chem. Lett.* **10**, 3296–3300 (2019).
26. G. Silversmit, D. Depla, H. Poelman, G. B. Marin, and R. De Gryse, *J. Electron Spectros. Relat. Phenomena*, **135**, 167–175 (2004).
27. A. Benayad, H. Martinez, A. Gies, B. Pecquenard, A. Levasseur, and D. Gonbeau, *J.*

Electron Spectros. Relat. Phenomena, **150**, 1–10 (2006).

28. A. Benayad, H. Martinez, A. Gies, B. Pecquenard, A. Levasseur, and D. Gonbeau, *J. Phys. Chem. Solids*, **67**, 1320–1324 (2006).

29. B. Fleutot, H. Martinez, B. Pecquenard, J. B. Ledeuil, A. Levasseur, and D. Gonbeau, *J. Power Sources*, **180**, 836–844 (2008).

30. E. Hryha, E. Rutqvist, and L. Nyborg, *Surf. Interface Anal.*, **44**, 1022–1025 (2012).

31. S. S. Fedotov, N. D. Luchinin, D. A. Aksyonov, A. V. Morozov, S. V. Ryazantsev, M. Gaboardi, J. R. Plaisier, K. J. Stevenson, A. M. Abakumov, and E. V. Antipov, *Nat. Commun.*, **11**, 1484 (2020) <http://www.nature.com/articles/s41467-020-15244-6>.

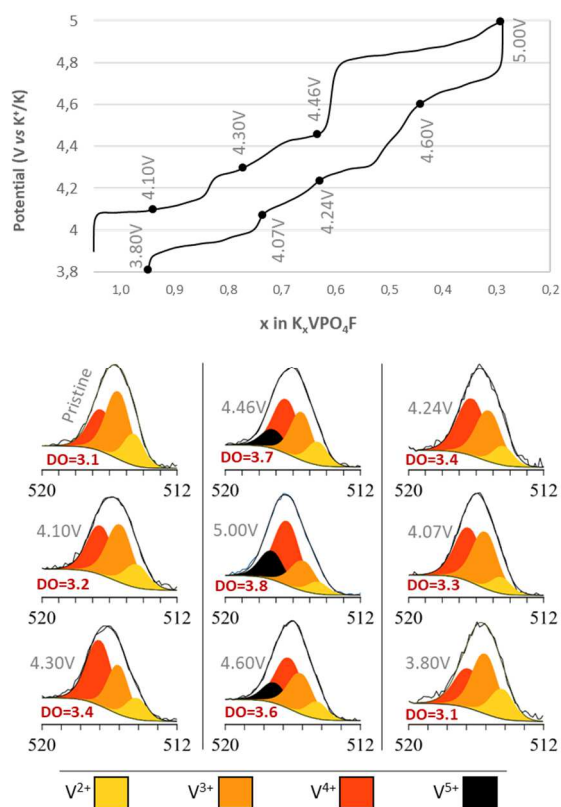


Fig. 1 a) Typical 3rd charge/discharge cycle (at C/5, 26 mA.g⁻¹) obtained for K//KVPF-C cells indicating at which potentials cells were stopped for the analysis of the redox processes involved in KVPF-C. b) Vanadium 2p_{3/2} XPS core spectra of K_xVⁿ⁺PO₄F_{0.9}O_{0.1}-C electrodes as recovered from K//KVPF-C cells during the 3rd cycle, as function of the potential.

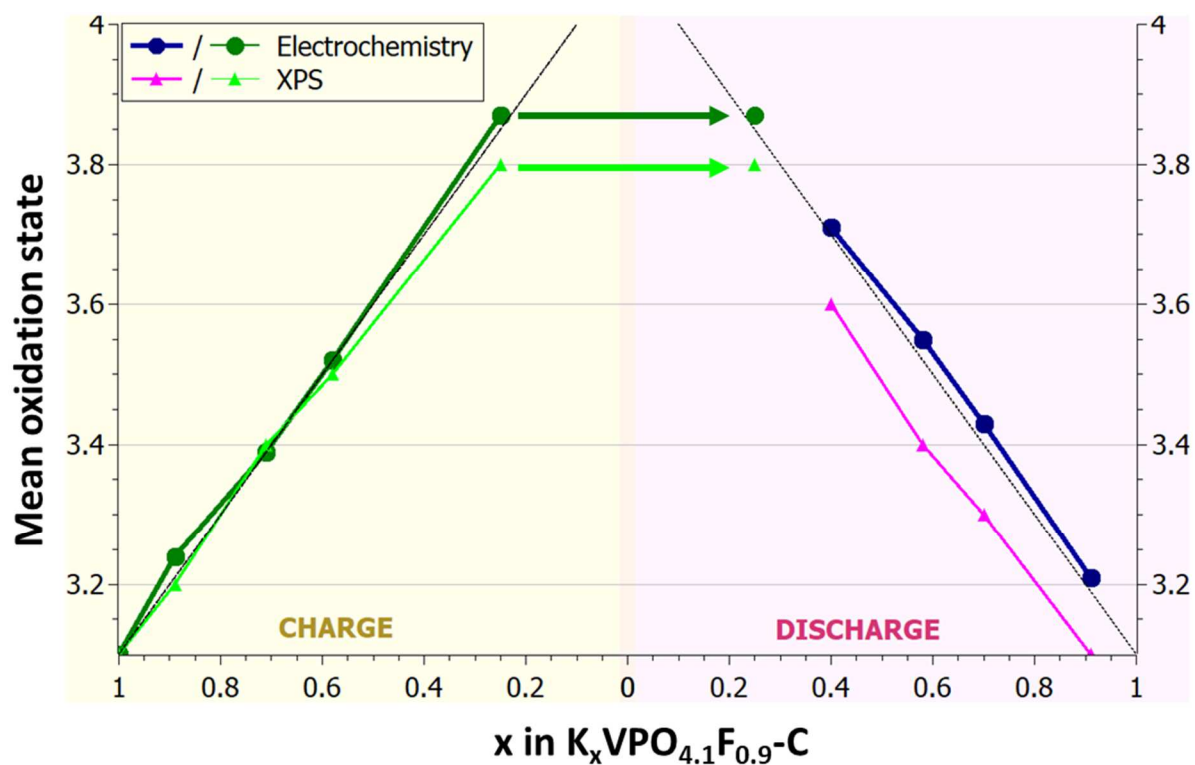


Fig. 2. Vanadium mean oxidation state (DO) of $KVPO_4F_{0.9}-O_{0.1}-C$, as the K content changed during the 3rd cycle, determined from the charge/discharge capacity on one side and from the analysis of the XPS spectra on the other side. Grey lines represent the theoretical DO based on the $KVPO_4F_{0.9}-O_{0.1}-C$ formula.

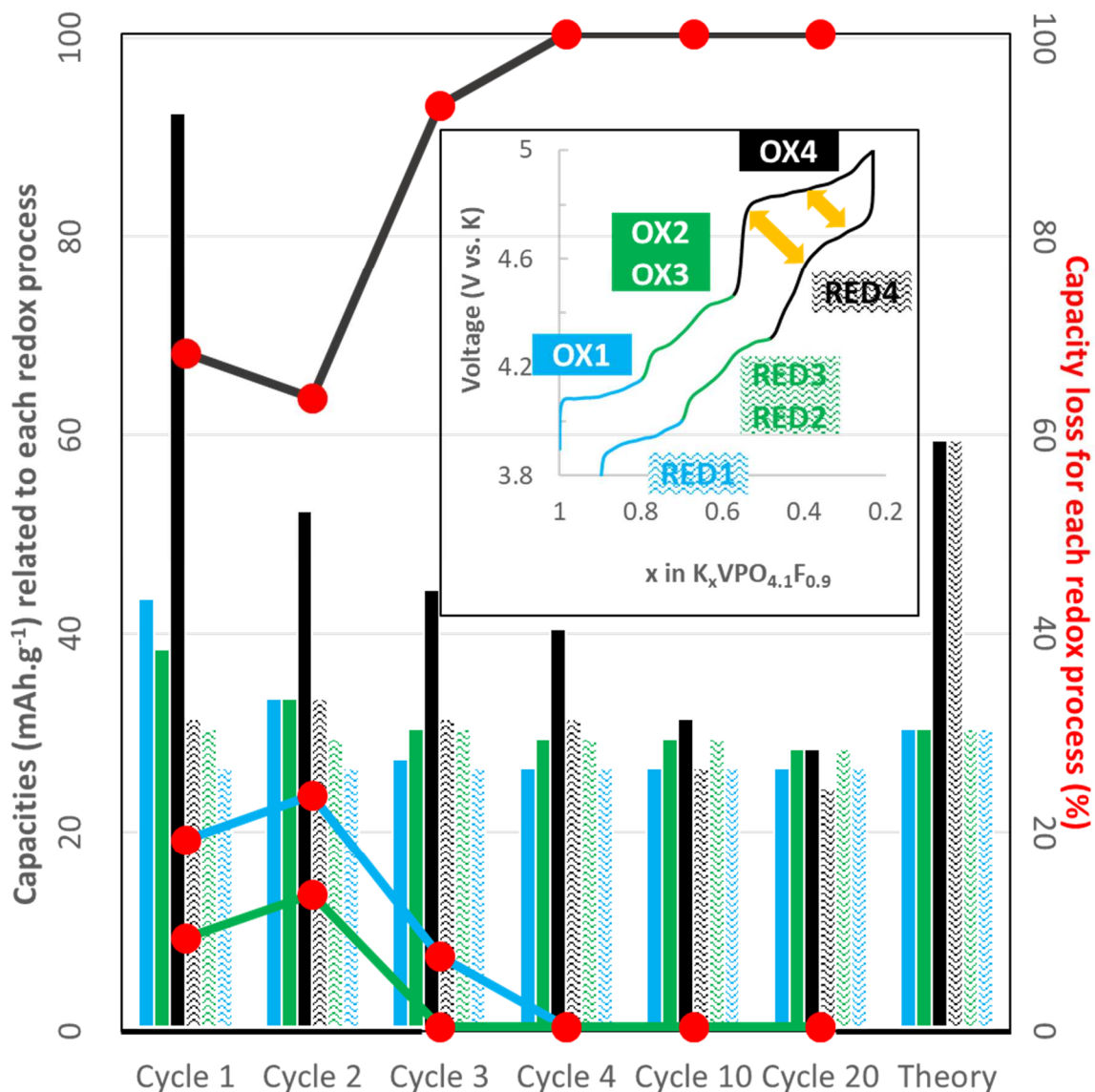


Fig. 3 Capacity (mAh.g⁻¹) related to the 1st, 2nd, 3rd and 4th redox processes as obtained from the derivative of the galvanostatic charge/discharge curves. Note that the 2nd and 3rd potential domains were combined as they cannot be perfectly dissociated after few cycles (not shown). Corresponding capacity loss (%) between oxidation and reduction is also reported by the curves with the red points. Theoretical capacity of the different redox processes based on the KVPO₄F_{0.9}-O_{0.1}C formula (118 mAh.g⁻¹) are also reported for comparison.