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# Dynamic Pore-Scale Dissolution by CO<sub>2</sub>-Saturated Brine in Carbonates: Impact of Homogeneous Versus Fractured Versus Vuggy Pore Structure

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# Water Resources Research

## RESEARCH ARTICLE

10.1029/2019WR026112

### Key Points:

- The impact of limestone pore structure on reactive transport was analyzed experimentally and computationally
- The dissolution patterns of homogeneous, fractured, and vuggy carbonate rocks are quantified and discussed in detail
- Increasing complexity of the pore geometry leads to more nonuniform dissolution

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## Dynamic Pore-Scale Dissolution by CO<sub>2</sub>-Saturated Brine in Carbonates: Impact of Homogeneous Versus Fractured Versus Vuggy Pore Structure

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**Abstract** Limestone pore structure strongly influences dissolution and associated reactive transport. These effects are critical in limestone diagenesis and but also in engineering operations such as carbon capture and storage (CCS). However, detailed studies on how CO<sub>2</sub>-enriched (acidic) brine changes this pore structure at relevant reservoir storage conditions are very limited. Thus, to provide further quantitative information and more fundamental understanding about these key processes, we studied the dissolution patterns of a homogeneous, a fractured, and a vuggy limestone when flooded with CO<sub>2</sub>-saturated brine at representative storage conditions. The pore structure of these limestones showed drastically different responses to the acidic brine flood. As such, preferential channels surrounded by branched channels were formed in the homogeneous sample, while fractures became the main flow path in the fractured sample. In contrast, only one dominant channel formed in the vuggy sample, which resulted in a sharp permeability increase. These dissolution patterns reflect the associated Damköhler number, which is significantly lower in the homogeneous, representing uniform dissolution. However, after injecting sufficient reactive fluid (1,000 PV), this uniform dissolution pattern transformed into a single preferential channel growth. Moreover, we conclude that increasing complexity of the pore geometry leads to more nonuniform dissolution. These dissolution patterns indicate the effect of initial pore structure on preferential channel growth and reaction transport. Our work provides key fundamental data for further quantifying limestone dissolution patterns in CCS, indicating that the CO<sub>2</sub> injection may cause the reactivation of geological faults and damage around wellbore, thus aids in the implementation of industrial-scale CCS.

## 1. Introduction

Carbon capture and storage has been recognized as the primary solution for mitigating anthropogenic CO<sub>2</sub> emissions (Bachu, 2000; Lackner, 2003), and deep saline aquifers are the largest potential target reservoirs (Celia et al., 2015; Zhang et al., 2019). Dissolution trapping is one of the secure capture mechanisms where the buoyant CO<sub>2</sub> dissolves in the formation brine; this acidified brine sinks deep into the reservoir and leads to the partial dissolution of carbonate reservoirs (Celia et al., 2015; Iglauer, 2011; Miao et al., 2020). In order to assess CO<sub>2</sub> storage capacity and security, a micrometer pore-scale analysis of the reactive transport processes is required to account for the complex phenomena occurring in these limestone formations (Li et al., 2006; Li et al., 2008; Li et al., 2014; Menke et al., 2015; Noiriél & Daval, 2017).

Carbonate dissolution is controlled by many factors including physical heterogeneity at multiple scales, chemical heterogeneity (spatial distribution of minerals), temperature, pressure, initial pH, and flow rates. In this paper, the dissolution pattern of three carbonate rocks with typical pore structure (homogeneous, fractured, and vuggy) are discussed.

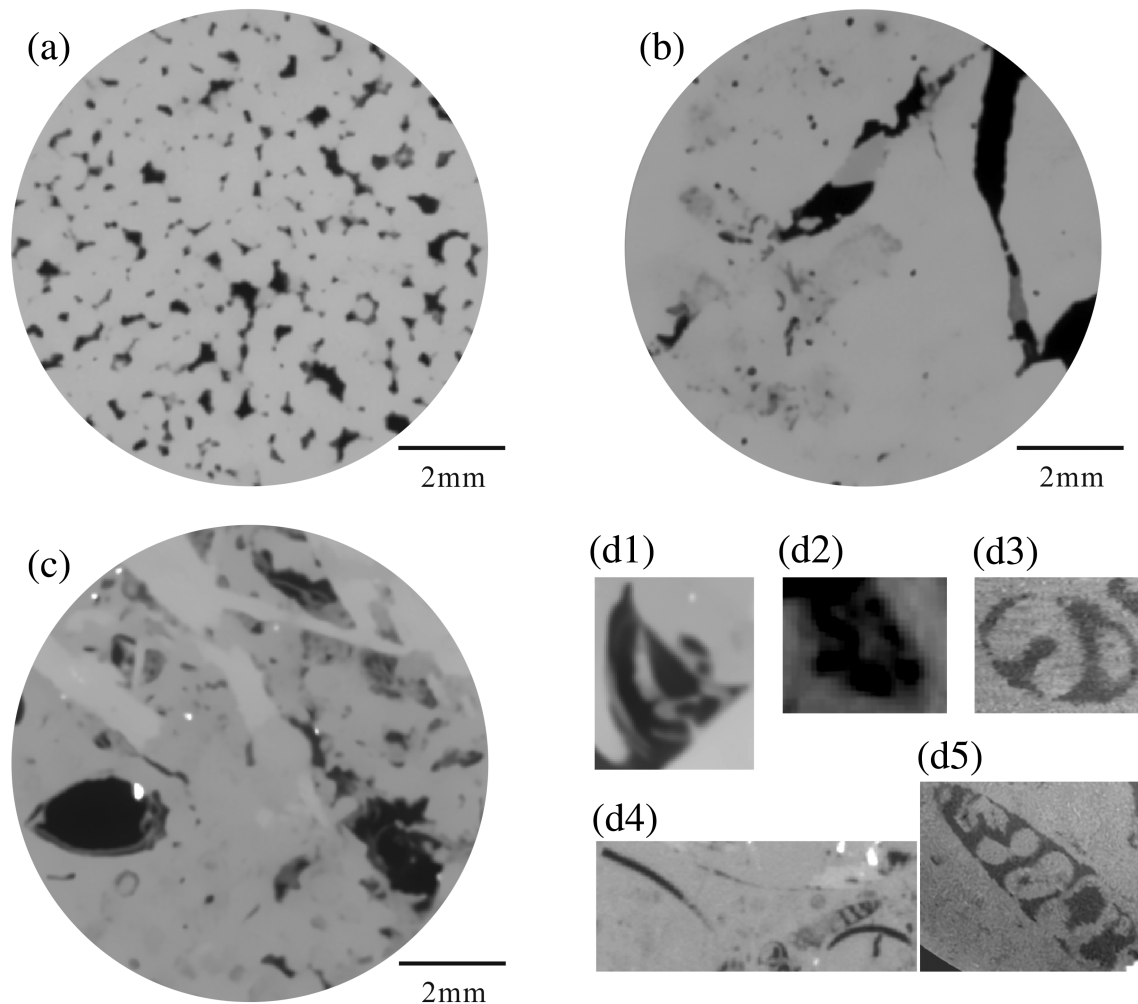
As such a number of previous studies investigated pore-scale reactive transport (Gharbi et al., 2013; Lebedev et al., 2017; Menke et al., 2016; Noiriél et al., 2004; Noiriél et al., 2005). It was shown that

dissolution is uniform in homogeneous limestone, while channels (“wormholing”) are formed in heterogeneous carbonate (Al-Khulaifi et al., 2019; Maheshwari et al., 2013). Especially in single mineral systems, the dissolution regime is structure dependent. In the works of Menke et al. (2015), the dissolution of a millimeter-scale limestone with a relatively simple pore structure at reservoir conditions was captured by microcomputed tomography (micro-CT) imaging and they found that the dissolution along the sample axis was uniform. Menke et al. (2016) examined the dissolution of two heterogeneous carbonates to investigate the impact of initial pore structure and found a new dissolution regime: The lower porosity sample experienced fast channel formation, in this case, the formation conditions were dependent on initial pore topology rather than flow rates. Thus far, most single-mineral imaging of CO<sub>2</sub> acidic brine dissolution is featured by homogeneous or heterogeneous limestones. Luquot et al. (2014) performed a set of four reactive flow experiments to describe the mass transfer process and the exponent of the permeability-porosity relationship. Smith et al. (2017) observed stable dissolution fronts in the more homogeneous dolomites and preferential flow paths in heterogeneous dolomites; they believed the porosity-permeability relationship is sensitive to pore space heterogeneity. However, the dissolution processes in fractured and vuggy carbonates are not fully understood, although limestones are typically fractured and contain large pores (“vugs”). Hence, the main goal of this study is to investigate how fractured and vuggy limestone pore structure impact on the reactive transport processes; this is then compared with the homogeneous benchmark case.

With the deepening of research, factors other than initial pore structure are added to the experiments of reactive transport. Luquot and Gouze (2009) reported nonuniform dissolution features near the injection well, while uniform dissolution was observed far from the injection well. However, the main factor controlling the two dissolution patterns is reactive brine with different partial pressures of CO<sub>2</sub> rather than pore structure. Menke et al. (2017) described reaction in both homogeneous and heterogeneous carbonates for multiple brine pH and found that channeling was faster with increasing heterogeneity and decreasing pH. In another work, Al-Khulaifi et al. (2018) combined pore-scale experimental observation and modeling to characterize reactive transport in two heterogeneity groups dolomite under two flow rates. Al-Khulaifi et al. (2017) imaged the dissolution of the composite core to examine the effect of physical and chemical heterogeneity on dissolution regime and effective reaction rate. In recent work, Al-Khulaifi et al. (2019) studied the impact of physical and chemical heterogeneity on the reaction rates in a multimineral carbonate.

Several studies have investigated dissolution in fractured carbonates (Andreani et al., 2008; Gouze et al., 2003; Szymczak & Ladd, 2009) and flow behavior in vuggy rocks (Arbogast et al., 2004; Kamath et al., 1998; Moctezuma-Berthier & Fleury, 2000; Zhang et al., 2005). They found that fracture dissolution regimes depend on fracture morphology, flow patterns, and mineralogical composition of the rock matrix. In the experiments of Detwiler (2008), Dijk et al. (2002), Ellis et al. (2011), and Noiriél et al. (2007), the dissolution of fractured carbonate rocks by flowing CO<sub>2</sub>-saturated brine was captured with nuclear magnetic resonance and micro-CT imaging. The intent was to identify the correlations between dissolution patterns, the fracture morphology, and mineralogical composition of the rock matrix by assessing spatial variability of dissolution rates and reactive surface changes. It was shown that at high Damköhler numbers, the dissolution patterns were dominated by the flow path. A large-scale fast channel was formed in the areas with little dissolution and slow closure of the fracture surfaces. By contrast, the low Damköhler experiment exhibited relatively uniform dissolution across the width of the fracture, with locally enhanced dissolution in small aperture regions. The dissolution patterns depended mainly on the mineral composition of the rock matrix. Finally, carbonate dissolution rate decreased due to reorganization of clay minerals. Izgec et al. (2010) explored the effects of large-scale vugs on carbonate acidizing with micro-CT imaging and numerical simulations. They observed that acid spreads in wormholes through vuggy rocks faster than those in homogeneous rocks and suggested that the local pressure drop created by vug was more important in wormhole flow path formation than the chemical reaction at the pore scale.

As seen above, there are extensive experimental and numerical studies of dissolution processes in limestone, such as wormholing in matrix and channeling in homogeneous, fractured, and vuggy carbonate (including limestone) rocks. However, these studies only carry out experiments or flow simulations in samples with single pore structure. Even if the effect of pore structure heterogeneity was discussed, this heterogeneity cannot represent fractured and vuggy carbonate rocks. Moreover, few studies have compared the dissolution



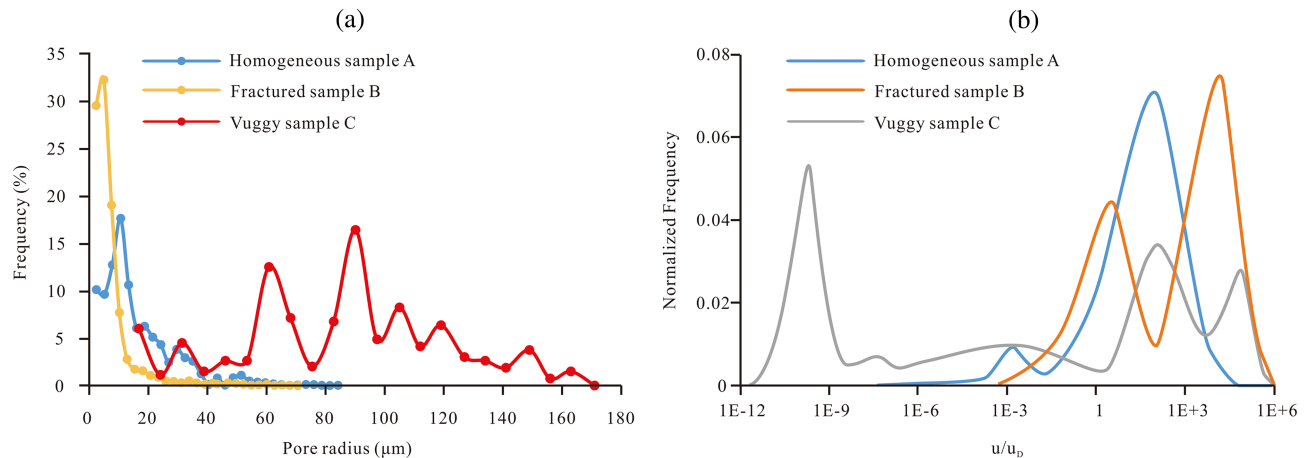
**Figure 1.** Two-dimensional slices through homogeneous carbonate rock (a), fractured carbonate rock (b), vuggy carbonate rock (c), and biodebris including gas-tropods (d1 and d4), coral (d2), spotted algae (d3), and foraminifera (d5) at 21.14  $\mu\text{m}$  voxel resolution.

patterns, effective reaction rates, porosity and permeability relationship, and the distribution of Damköhler and Péclet numbers in the three typical carbonate rocks.

X-ray microtomography has been successfully used to study the dynamic pore-scale changes induced by  $\text{CO}_2$ -acidified brine injection (Al-Khulaifi et al., 2017; Al-Khulaifi et al., 2018; Gouze & Luquot, 2011; Lebedev et al., 2017; Luquot & Gouze, 2009) and pore space characterization (Song et al., 2019; Yang et al., 2019; Yang et al., 2020). In this investigation, we combined experimental and numerical simulation (Lattice Boltzmann) methods and examined the detailed dissolution processes in homogeneous, fractured, and vuggy carbonate rocks. Experimentally, high-resolution in situ X-ray microtomography images were acquired before and after  $\text{CO}_2$ -saturated brine flooding. In addition, for each experimental group, the permeability was simulated for different displacement stages, and associated transport and reaction parameters

**Table 1**  
*Petrophysical and Mineralogical Rock Properties*

Sample	Initial pore structure	Porosity (%)	Permeability ( $\times 10^{-3} \mu\text{m}^2$ )	Calcite (%)	Quartz (%)	Anhydrite (%)
A	Homogeneous	9	1392.83	98	2	0
B	Fractured	15	77.75	96	4	0
C	Vuggy	10.65	16.53	75	1	24



**Figure 2.** The pore radius distributions (a) at 4.26  $\mu\text{m}$  resolution and initial velocity distributions (b) of each sample. Sample A had more homogeneous pore structure, while B and C had more heterogeneous pore structure.

(i.e., Péclet and Damköhler numbers, and effective reaction rates) were calculated to quantify the dissolution processes for homogeneous, fractured, and vuggy carbonate rocks. Thus, our study provides dissolution patterns, permeability changes, advection-diffusion effects, and  $Pe$ - $Da$  distributions for various pore structure. This information is required for accurate predictive modeling, which can determine how the  $\text{CO}_2$ -acidified brine moves in the subsurface. This again is the basis for carbon storage capacity predictions and risk assessments.

## 2. Materials and Methods

### 2.1. Carbonate Samples

Three samples were selected for the reactive transport experiments, these include homogeneous, fractured and vuggy carbonate rocks, respectively. All samples were cylindrical, and 2 cm long with a 9 mm diameter. Figure 1 shows X-ray microtomography (micro-CT) grayscale image slices of each carbonate tested. Permeability and porosity were measured by a QKY-II instrument using nitrogen on standard samples and mineral composition was measured by X-ray diffraction instrument (cf. Table 1). Sample A was acquired from the Feixianguan formation deposited at Triassic age. Sample B and Sample C were collected from Northeastern Turkmenistan (both deposited at Jurassic age). Sample B contained natural fractures, and Sample C contained vugs, that is, intergranular dissolved pores. Note that Sample C contained rich biodebris (such as gastropods, foraminifera, brachiopod, spotted algae, and coral); this contributes to the formation of vugs. The petrophysical characteristics of each sample are listed in Table 1. Samples A and B were essentially chemically homogeneous (with minor parts of quartz which is effectively nonreactive), while Sample C was a composite and contained three minerals (note that anhydrite does not react with  $\text{CO}_2$ -acidified brine but dissolves in trace amounts).

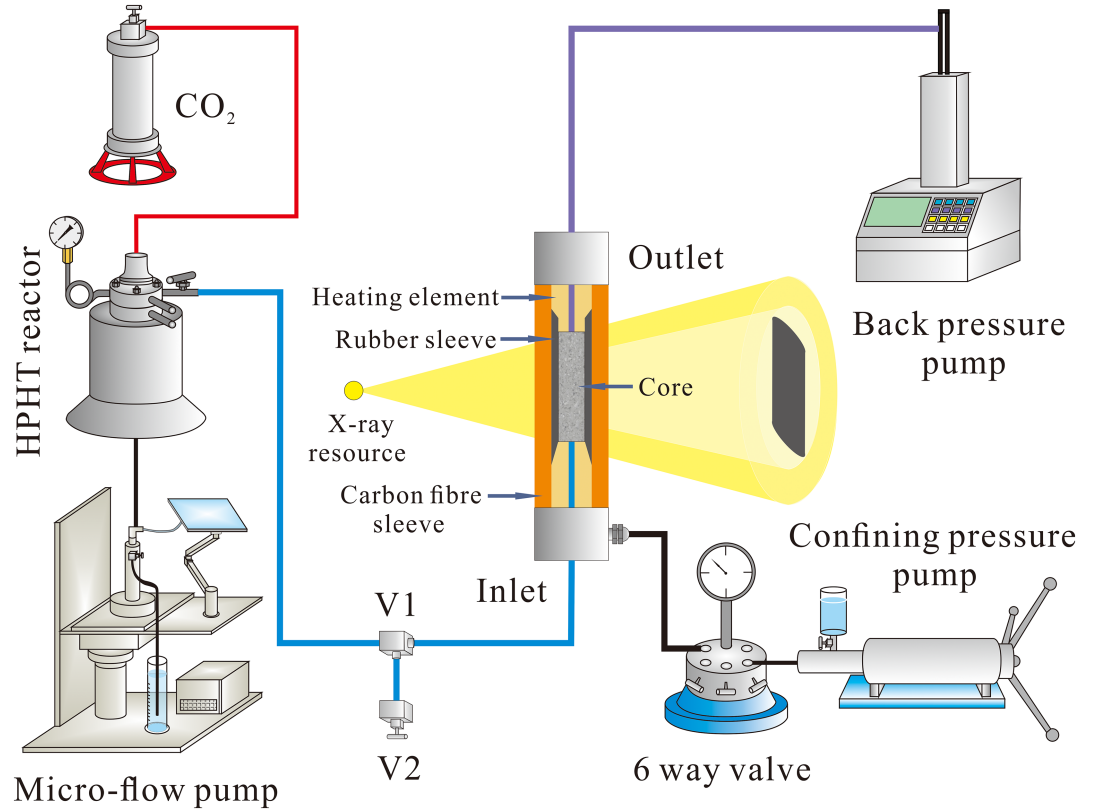
### 2.2. Sample Heterogeneity Characterization

Lattice Boltzmann method was used to simulate the associated fluid flow (Sun et al., 2017; Zhang et al., 2019). Thus, to calculate the permeability of the samples, the binary data file of the digital core was imported into a D3Q19 lattice model (Qian et al., 1992; Zhu et al., 2019). Here the distribution function  $f_i(x)$  represents the average movement of fluid particles. The left part of equation 1 represents the migration of particles, while the right represents the collision of particles, and  $\tau$  is the relaxation parameter (equation 2):

$$f_i(x + e_i \Delta t, t + \Delta t) - f_i(x, t) = -\frac{1}{\tau} [f_i(x, t) - f_i^{eq}(x, t)] \quad (1)$$

$$\tau = \frac{\nu}{c_s^2 \Delta t} + 0.5 \quad (2)$$

where  $\nu$  is the kinematic viscosity of the fluid and  $c_s$  ( $c_s = \frac{1}{\sqrt{3}}$ ) represents the lattice pseudo-sound-speed.



**Figure 3.** The in situ micro-CT core flooding apparatus used consisted of three pumps, a HPHT reactor, and a core holder.

The  $f_i^{eq}(x)$  in equation 1 is the equilibrium distribution function of the incompressible D3Q19 model (He & Luo, 1997):

$$f_i^{eq}(x) = \omega_i \left[ \rho(x) + \rho_0 \left( 3(e_i \cdot u) + \frac{9}{2}(e_i \cdot u)^2 - \frac{3}{2}(u \cdot u) \right) \right] \quad (3)$$

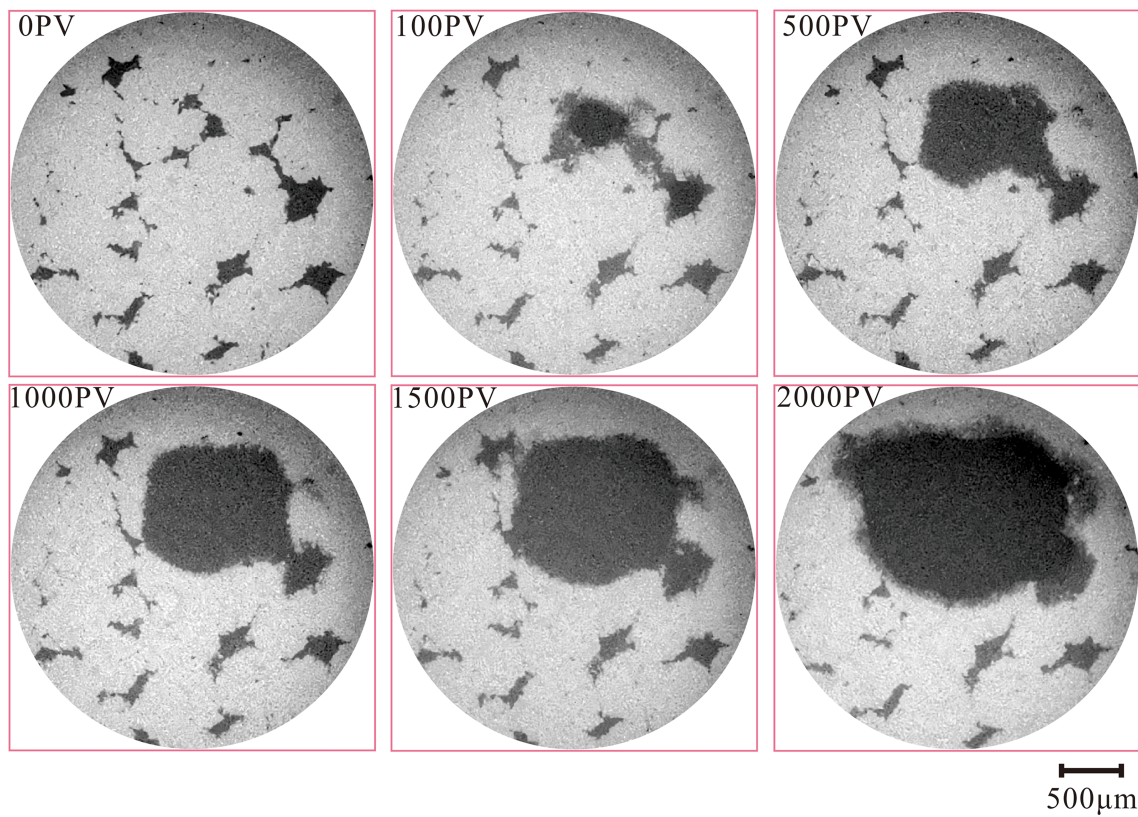
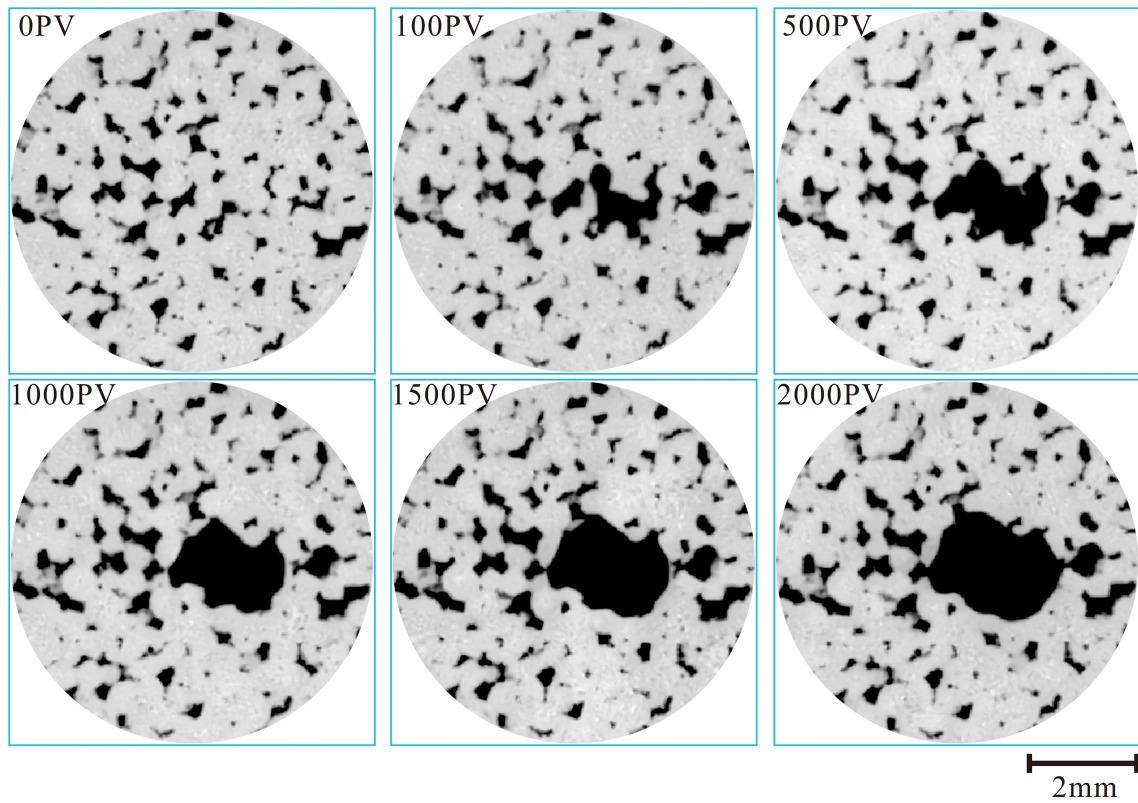
where  $\rho_0$  is the average fluid density and  $\omega_i$  is the weight factor in  $i$ th direction  $\omega_i$  for D3Q19 is identified as

$$\omega_i = \begin{cases} 1/3, & i = 0 \\ 1/18, & i = 1, 2, 3, 4, 5, 6, \\ 1/36, & i = 7, 8, 9, 10, 11, 12, 13, 14, 15, 16, 17, 18 \end{cases}$$

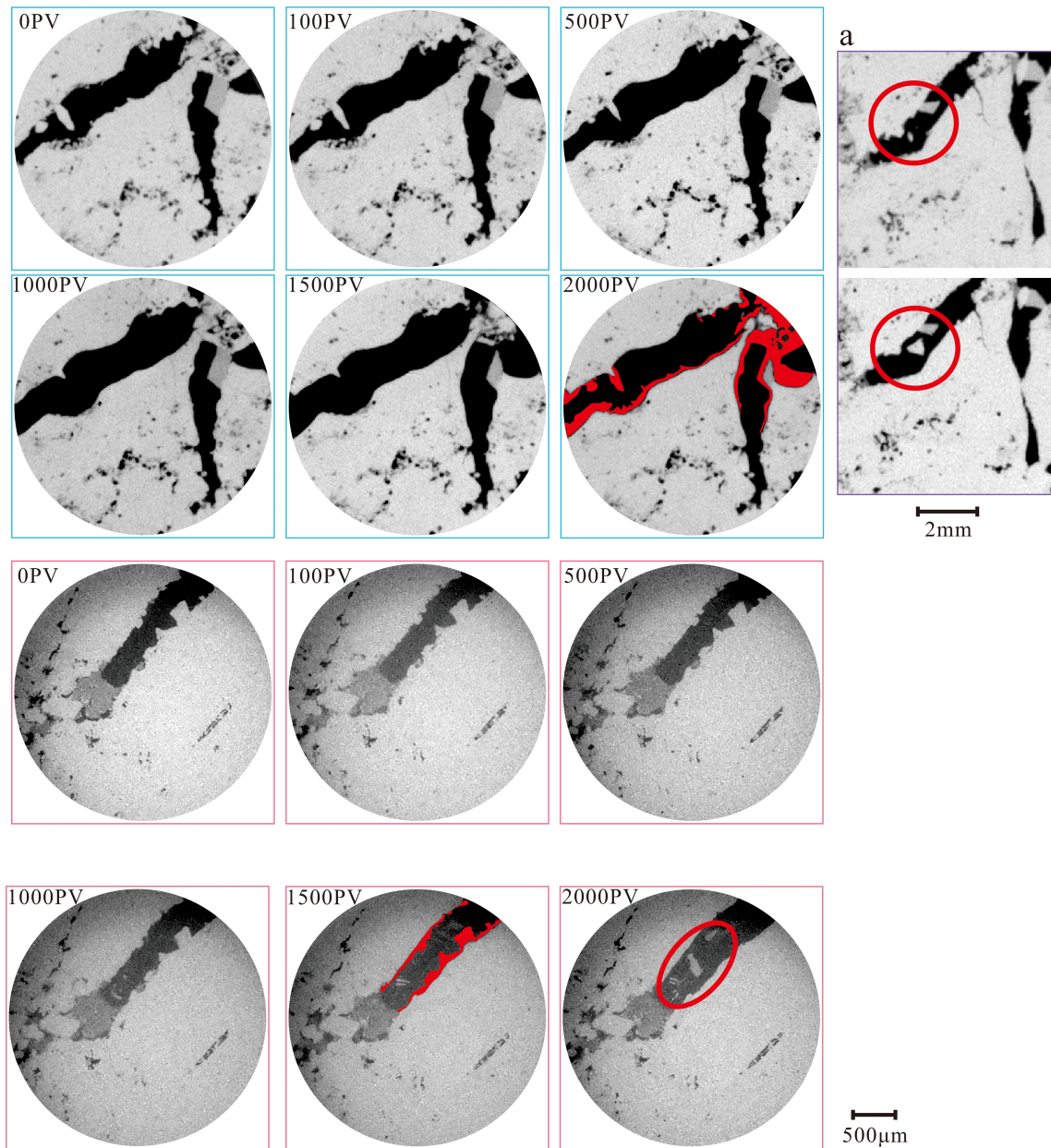
Besides, the macrodensity of the fluid ( $\rho$ ), lattice velocity ( $u$ ), and macroscopic pressure ( $p$ ) at each lattice unit were calculated as  $\rho(x) = \sum_{i=0}^{18} f_i(x)$ ,  $u = \frac{1}{\rho(x)} \sum_{i=0}^{18} f_i(x) e_i$ , and  $p(x) = c_s^2 \rho(x)$ .

Finally, the flow rate  $Q$  ( $m^3/s$ ) was obtained as  $Q = \sum_{j=1}^n u(x_j) A(x_j)$  where  $u(x_j)$  (m/s) is the velocity at each outlet node and  $A(x_j)$  is the area of each lattice at the outlet. The Darcy velocity was calculated as  $u_D = Q / \sum_{j=1}^n A(x_j)$ . From these simulations, the velocity distribution was obtained for the lattice space. Note that the initial velocity distribution can characterize the physical heterogeneity.

In this simulation, single-phase fluid was driven by differential pressure. The driven acceleration is  $5 \times 10^{-5}$  in  $x$  direction, and all units are lattice units (Yang et al., 2020; Zhao et al., 2016). Absolute permeability is independent of the type of fluid; thus, the boundary is static solid boundary. For this boundary, the commonly used method is to bounce back the particles on the boundary, which is called the standard bounce-back scheme. We used this scheme to convert the walls of the porous media to



**Figure 4.** Two-dimensional micro-CT slices through Sample A at 21.14 (blue boxes) and 4.26 μm (red boxes) resolution.



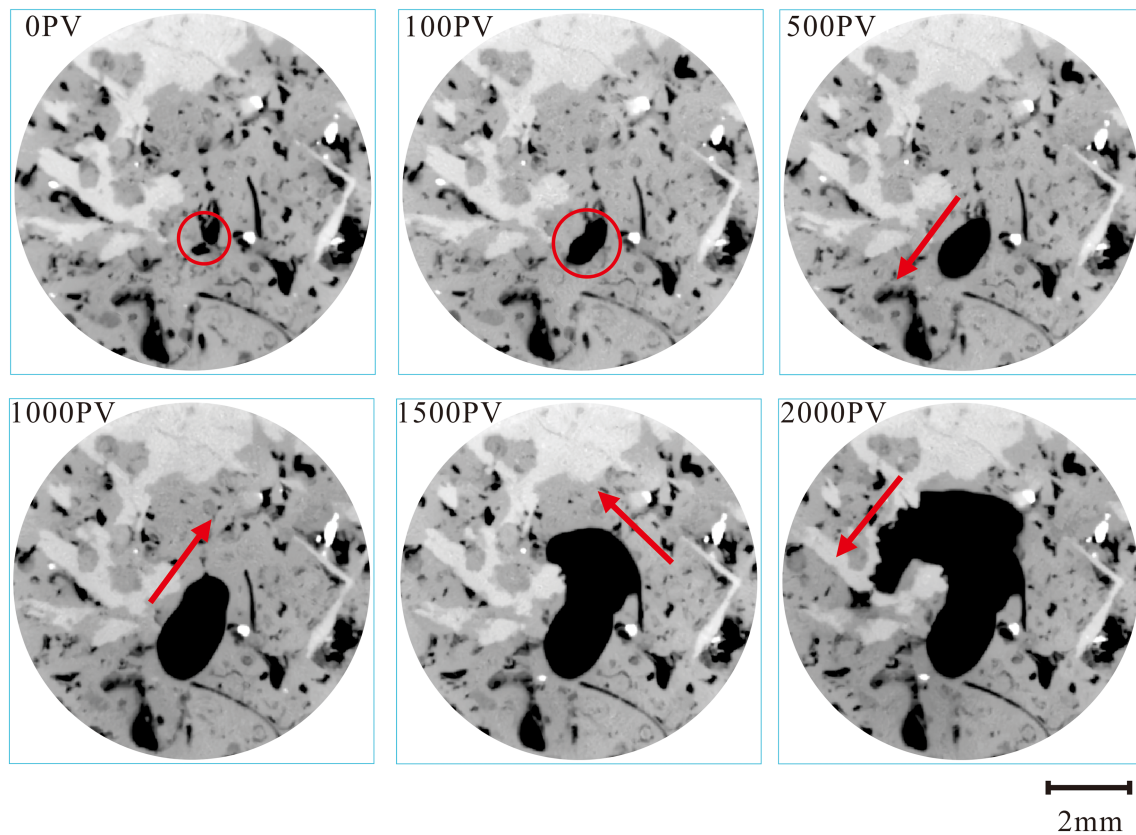
**Figure 5.** Two-dimensional micro-CT slices through Sample B at 21.14  $\mu\text{m}$  (blue boxes) and 4.26  $\mu\text{m}$  (red boxes) resolution. Red areas indicate reacted (dissolved) voxels; migrated fines appear in the red circle.

bounce-back boundary conditions (Maier & Bernard, 2010; Ziegler, 1993). In this scheme, the particle will return to the fluid node along the original path when it touches a wall node. The speed and density of the fluid in the porous medium reaches a steady state until the difference between adjacent times is less than  $10^{-6}$ . According to Darcy's law, the equivalent permeability of each core was calculated as

$$K = -\mu \frac{u_D}{\nabla p} \quad (4)$$

Probability density functions can be used to characterize the physical heterogeneity (Bijeljic et al., 2013, 2013). From the simulations we obtained the lattice velocity  $u$  (the velocity in each





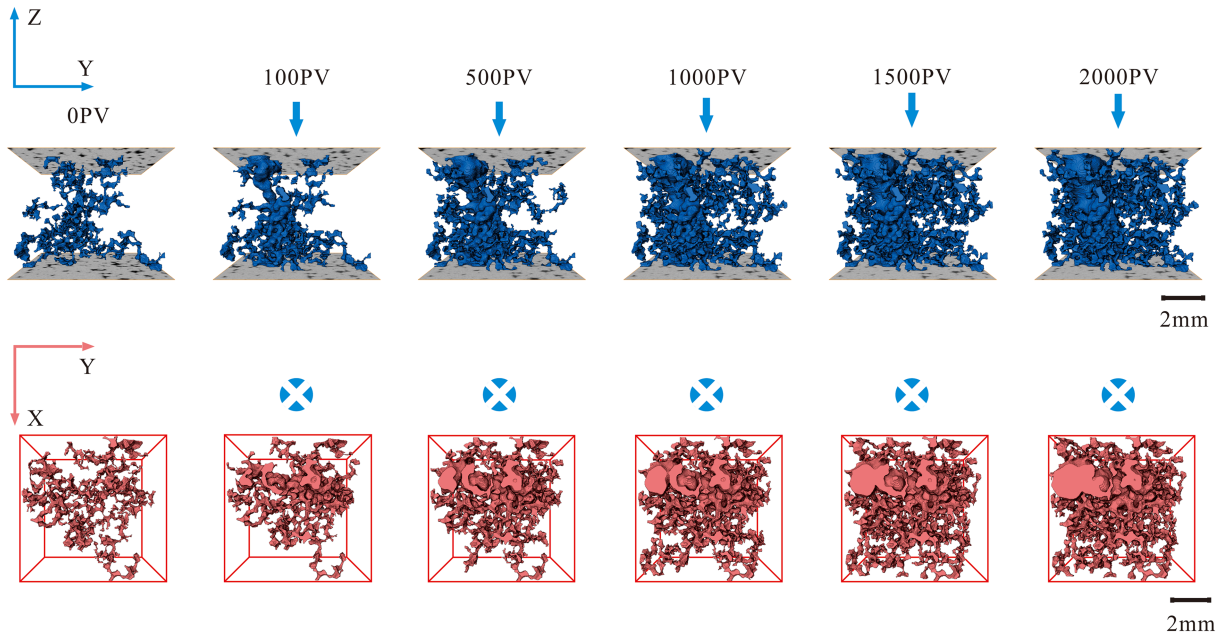
**Figure 6.** Two-dimensional micro-CT slices through the Sample C at 21.14  $\mu\text{m}$  resolution. Red circles represent dissolved hole, and the arrows represent the direction in which the vug grows.

lattice unit), which were divided by the Darcy velocity  $u_D$ , and lattice is generally regarded as a voxel of 3-D image. Through LBM simulation, Darcy velocities of Samples A, B, and C were  $1.02\text{E}-15$ ,  $1.60\text{E}-14$ , and  $3.13\text{E}-19$  m/s, respectively. Figure 2 shows the velocity distributions of the initial state of the three samples. Moreover, we extracted 3-D pore network model and calculated the pore radius distribution of each sample (Figure 2). Sample A had both single peaks in the distributions of pore radius and initial velocity, which indicates that Sample A is a homogeneous carbonate (Yang et al., 2019; Yang, Wang, et al., 2019). However, the two more heterogeneous distributions had multiple peaks, and the velocity distribution of cavernous sample was broader. The multiple peaks in the pore radius distribution curve reveal that Sample C has strong physical heterogeneity.

### 2.3. Experimental Methodology

A core was placed in a rubber sleeve and sandwiched between two heating elements (which were added upstream and downstream and maintained the core at 50 °C). This assembly was inserted into a carbon fiber core holder and a confining pressure of 2 MPa was applied through a separate deionized water pump, Figure 3. The core holder was then connected to the flow system, and the core and flow tubes were vacuumed for 6 hr. The  $\text{CO}_2$ -saturated brine was prepared by adding  $\text{CO}_2$  into a HPHT (high pressure and high temperature) reactor, which had been partially filled with 3 wt% NaCl brine. A pressure of 10 MPa and a temperature of 50 °C was applied and equilibrated brine was taken from the HPHT reactor (open the two-way valve V2) at the beginning of each experiment (El-Maghraby et al., 2012). The pH of  $\text{CO}_2$  acidified brine was verified to be 3.5 after testing with PHB-4 portable pH measuring instrument.

Initially the dry samples were imaged at two resolutions (4.26  $\mu\text{m}$  and 21.14  $\mu\text{m}$ ) under a confining pressure of 2 MPa. The core samples were then gradually pressurized to 8 MPa pore pressure by a back pressure pump while confining pressure was always 2 MPa higher than the pore pressure. Subsequently,  $\text{CO}_2$ -saturated brine was injected at 0.1 mL/min with a high-precision syringe pump. The cores were then imaged again



**Figure 7.** Three-dimensional visualizations ( $300 \times 300 \times 236$  voxels =  $6.34 \text{ mm} \times 6.34 \text{ mm} \times 4.99 \text{ mm} = 200.66 \text{ mm}^3$ ) in  $XY$  and  $YZ$  direction of the main pore space in the homogeneous limestone. The arrow represents the direction of  $\text{CO}_2$  acidified brine injection, and the cross represents the injection direction perpendicular to the  $XY$  plane ( $21.14 \mu\text{m}$  resolution images).

multiple times during the flood, after injecting 100 PV (pore volumes), 500, 1,000, 1,500, and 2,000 PV at the same high resolutions ( $4.26$  and  $21.14 \mu\text{m}$ ) to analyze the pore structural change (Iglauer & Lebedev, 2018). The raw tomograms were filtered with a 3-D nonlocal means filter to denoise the images (Buades et al., 2005; Yang et al., 2019), and the phases were segmented with a watershed algorithm based on different gray values (Schlüter et al., 2014; Yang et al., 2020). The pore phase was then extracted from the image for quantitative analysis, and the pore volume fraction (i.e.,  $\phi_{CT}$ ) was measured by counting the pore voxels and dividing them through the rock voxels.

#### 2.4. Determination of Effective Reaction Rate, Péclet, and Damköhler Number

Péclet number ( $Pe$ ) is defined as the ratio between the rate of mass transport of the solute to its molecular diffusion rate. We express  $u_{avg}$  as the average pore velocity of the fluid flow, calculated by dividing the Darcy velocity by  $\phi_{CT}$ ;  $\phi_{CT}$  is the porosity measured on the segmented images.  $L_c$  (m) is a characteristic length given by  $L_c = \pi/S$  where  $S$  ( $\text{m}^{-1}$ ) is the specific surface area measured on micro-CT images (Mostaghimi et al., 2012); and  $D_m$  is the molecular diffusion coefficient which has a value of  $7.5 \times 10^{-10} \text{ m}^2/\text{s}$  for  $\text{Ca}^{2+}$  (Luquot & Gouze, 2009) for these pressure and temperature conditions (10 MPa,  $50^\circ\text{C}$ ).

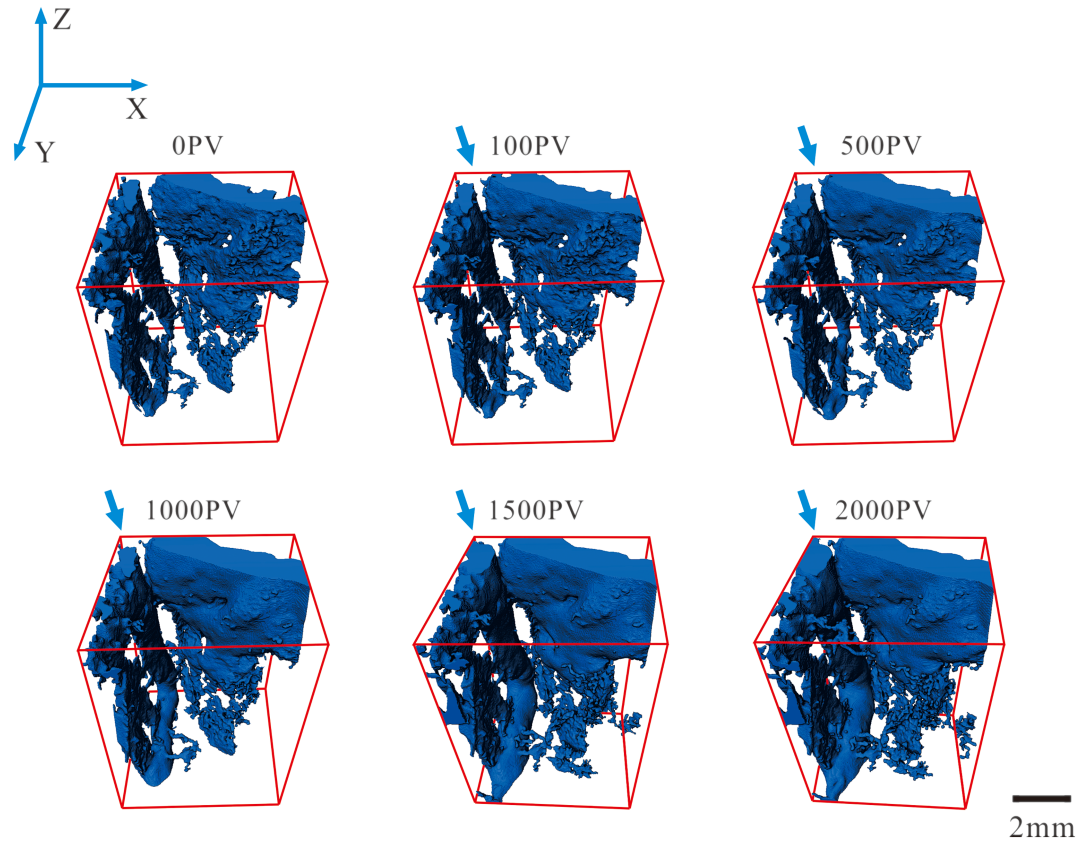
$$Pe = \frac{u_{avg} L_c}{D_m} \quad (5)$$

$$u_{avg} = \frac{u_D}{\phi_{CT}} \quad (6)$$

Moreover, Damköhler numbers ( $Da$ ) for the three samples were obtained.  $Da$  reflects the relationship between the chemical reaction rate and the mass transport rate, and it can be expressed as (Al-Khulaifi et al., 2019):

$$Da = \frac{\pi r_{\text{mineral}}}{u_{avg} n} \quad (7)$$

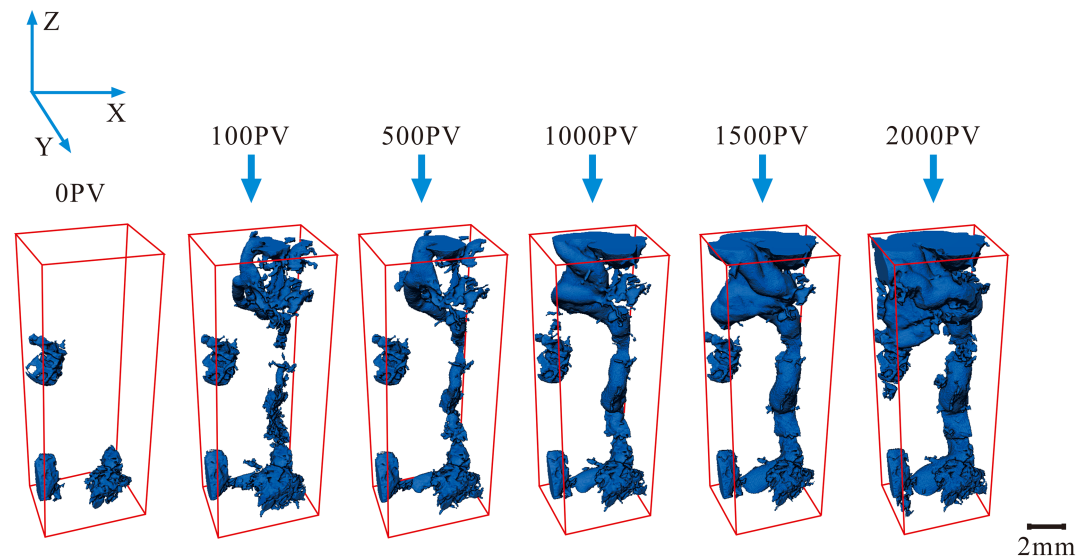
where  $r_{\text{mineral}}$  is the intrinsic reaction rate, which has a value of  $8.1 \times 10^{-4} \text{ mol}/\text{m}^2\text{s}$  for calcite (Peng et al., 2015; Peng et al., 2016), and  $n$  represents the molar number of minerals per unit rock volume.



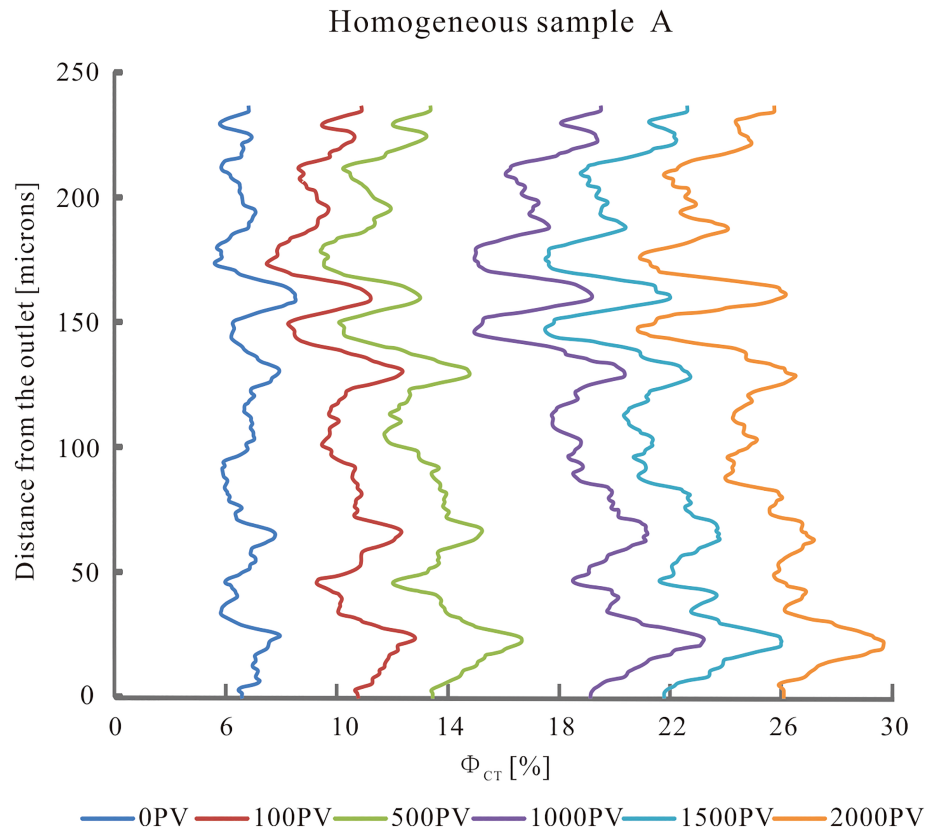
**Figure 8.** Three-dimensional visualizations ( $300 \times 300 \times 400$  voxels =  $6.34 \text{ mm} \times 6.34 \text{ mm} \times 8.46 \text{ mm} = 340.11 \text{ mm}^3$ ) of the pore space in the fractured limestone. The arrow represents the direction of  $\text{CO}_2$  acidified brine injection.

$$n = \frac{\rho_{\text{mineral}} f (1 - \phi_{CT})}{M_{\text{mineral}}} \quad (8)$$

where the density ( $\rho_{\text{mineral}}$ ) of calcite is  $2,710 \text{ kg/m}^3$ ,  $M_{\text{mineral}}$  is molecular mass of calcite ( $0.1001 \text{ kg/mol}$ ), and  $f$  is the percentage of calcite.



**Figure 9.** Three-dimensional visualizations ( $250 \times 250 \times 800$  voxels =  $5.29 \text{ mm} \times 5.29 \text{ mm} \times 16.91 \text{ mm} = 472.37 \text{ mm}^3$ ) of the pore space in the vuggy limestone. The arrow represents the direction of  $\text{CO}_2$  acidified brine injection.



**Figure 10.** Relationship between slice-averaged porosity and distance from the sample outlet, homogeneous Sample A.

Effective reaction rate is the reaction rate in core with irregular pore, rather than the intrinsic reaction rate under well-mixed conditions (Al-Khulaifi et al., 2017). Equation 9 thus captures an effective reaction rate ( $r_{eff}$ ) of the mineral dissolution, where  $\Delta t$  is the time interval between scans, and the mineral volume changes ( $\Delta\phi_{CT}$ ) were determined by segmenting consecutively scanned CT images. We assume that all the surface area contributes to the reaction with  $\text{CO}_2$  saturated brine although such surface area may be lower than the real due to the resolution limitation:

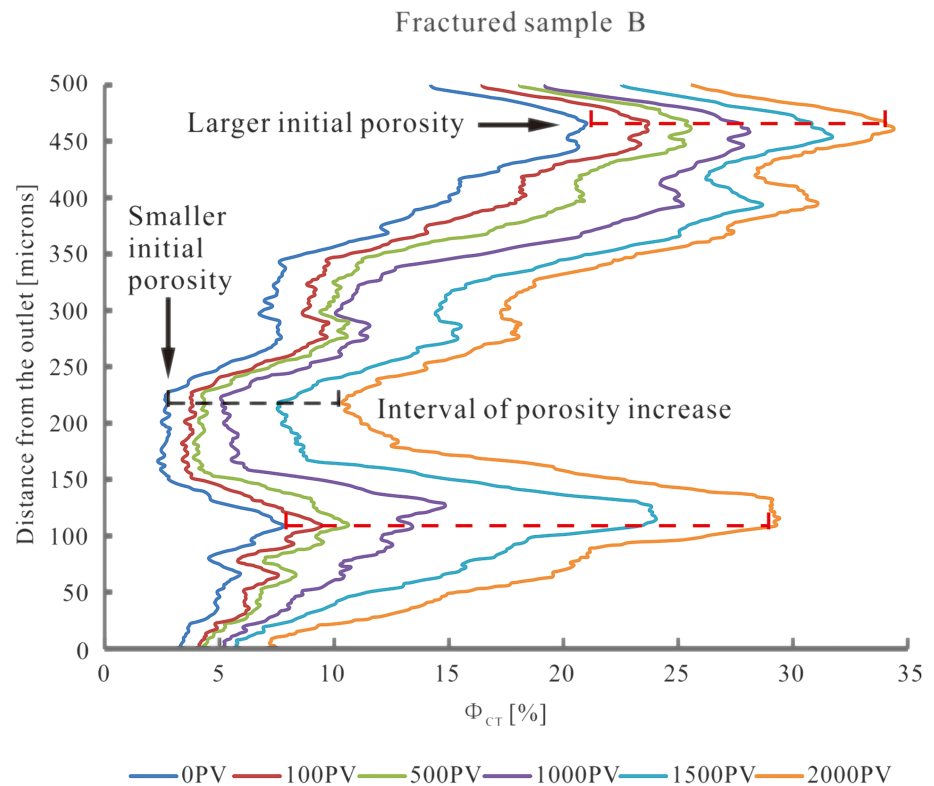
$$r_{eff} = \frac{\rho_{\text{mineral}}(1 - \phi_{CT}) \Delta\phi_{CT}}{M_{\text{mineral}}S \Delta t} \quad (9)$$

### 3. Results and Discussion

#### 3.1. Pore-Scale Limestone Dissolution

##### 3.1.1. Dissolution of Homogeneous Carbonate Rock

In the homogeneous carbonate (blue box in Figure 4, 21.14  $\mu\text{m}$ ), clearly, a single channel formed gradually. At higher resolution (4.26  $\mu\text{m}$ ), it could be observed that circular particles dissolved and pore space originally disconnected in 2-D plane became connected. The new pore space was formed adjacent to the original effective pore space, consistent with Lebedev et al. (2017) who showed that acidified brine partially flows through the high permeability part of the sample. However, Menke et al. (2017) found a uniform dissolution mechanism and several flow paths developed there throughout the rock due to the continuous exposure of the entire rock surface to the reactive fluid. Other areas of Sample A also dissolved, although they were spatially smaller but larger in number. This discrepancy was caused by the comparatively lower homogeneity of our sample, while uniform dissolution occurs only when grains are highly sorted and rounded.



**Figure 11.** Relationship between slice-averaged porosity and distance from the sample outlet, homogeneous Sample B.

### 3.1.2. Dissolution of Fractured Carbonate Rock

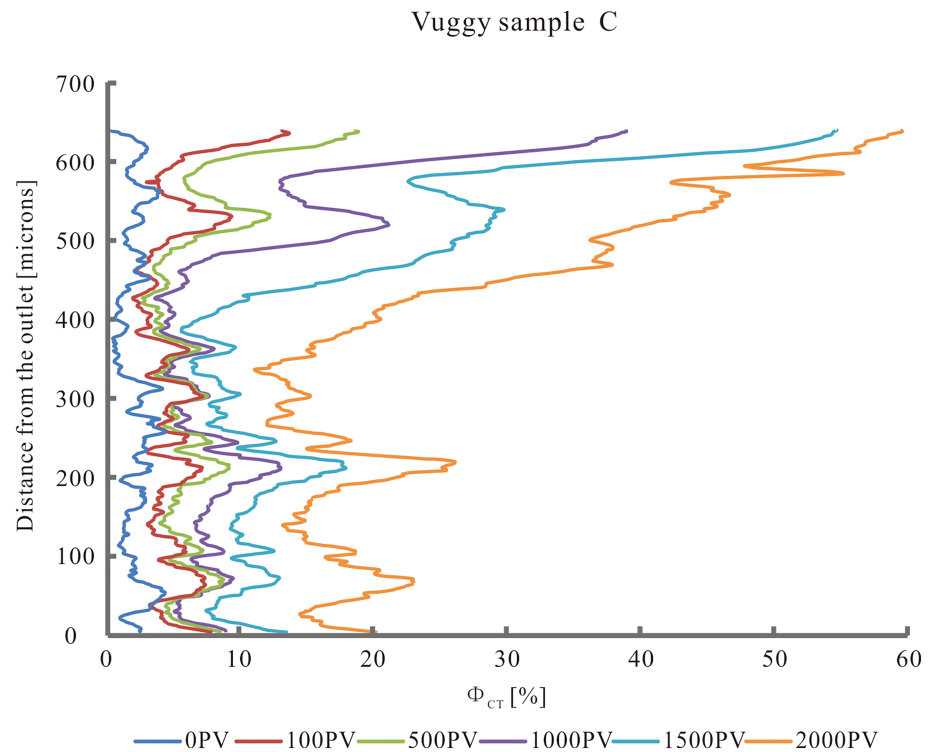
The fractured carbonate rock showed a distinctly different response. Clearly, the acidic brine flowed mainly through the fractures as expected, and dissolution thus occurred at the fracture walls, thus increasing the fracture aperture, Figure 5. Furthermore, some particles consisting of calcite crystals on the inner wall of the fractures were dissolved in the flooding process (red part of 2,000 and 1,500 PV in Figure 5). The inner wall of the fracture was also smoothed, while fracture aperture increased. At the same time, we also observed an interesting phenomenon as indicated by the red circles; that is, particles appeared in the fractures, apparently flowing downstream which further proves that the flow channel of the CO<sub>2</sub> acidified brine is confined to the main fractures.

### 3.1.3. Dissolution of Vuggy Carbonate Rock

Initially, this rock contained disconnected vugs; after injecting 100 PV of the acidic brine, one of these vugs enlarged (shown in the red circle in Figure 6) and eventually formed an irregularly shaped hole in the center of the visual field. Unlike dissolution in the homogeneous carbonate, dissolution proceeded in four directions. This is clearly related to the heterogeneous pore structure and chemical composition (this rock contains 24 wt% anhydrite, Table 1). Notably, the vug only grew into the calcite matrix (gray), but not into the anhydrite (white/light gray), during further injection of 1,500–2000 PV of acidic brine. Although anhydrite slightly dissolves in water at 50 °C, calcite reacts with acid to form Ca<sup>2+</sup>, which inhibits the dissolution of anhydrite. Therefore, we observed on 2-D slices that anhydrite was almost insoluble. The above described effect was, however, not observed in the high-resolution image (4.26 μm), as it occurred outside the field of view. This indicates that dissolution in vuggy carbonate is heterogeneous and localized, and thus, dissolution is strongly confined spatially, which is distinctly different to dissolution in homogeneous carbonate, see above. We conclude that areas outside the reactive vug are hardly affected by the reactive fluid.

## 3.2. Preferential Channel Growth and Porosity Analysis

In order to further quantify the dissolution induced changes in the pore morphology, we analyzed the 3-D spatial distribution of the largest flow channels and calculated dynamic permeabilities on the segmented images.



**Figure 12.** Relationship between slice-averaged porosity and distance from the sample outlet, homogeneous Sample C.

### 3.2.1. Analysis of the Connected 3-D Pore Space

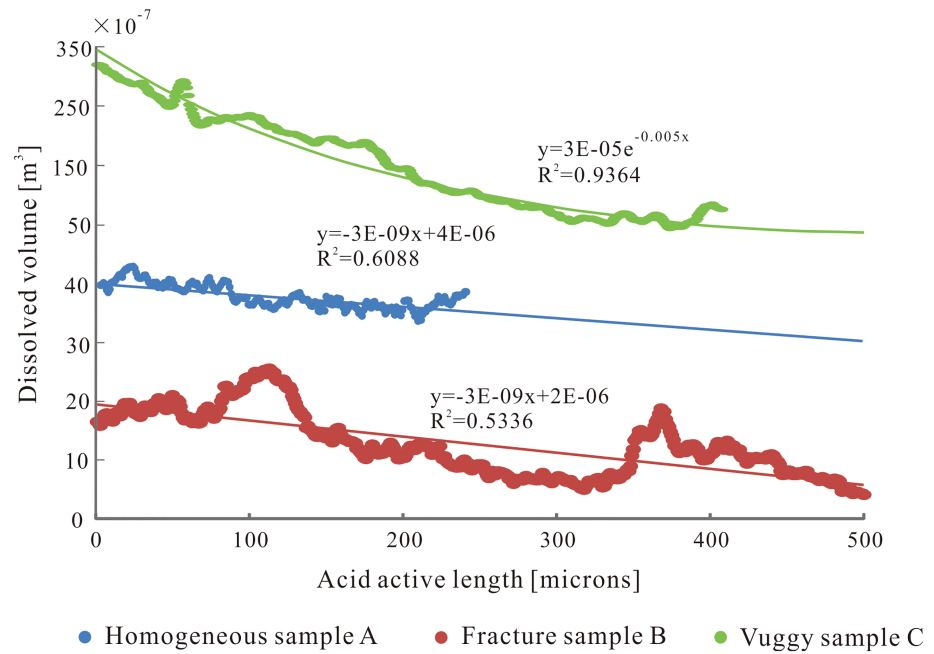
Figure 7 shows 3-D visualizations of the pores with a volume larger than  $10^8 \mu\text{m}^3$  for the homogeneous sample at different dissolution stages. A main channel formed in the XY and YZ directions, consistent with Luquot and Gouze (2009) and Al-Khulaifi et al. (2018). In addition, there were many branched channels around the main channel, which grew continuously until 1,500 PV of acid were injected, afterward these branches did almost not grow any longer. The reason is that the main channel became the dominant path of acid flow, which is confirmed by the permeability analysis in section 3.3: The permeability at 2,000 PV is below the power law curve; that is, the trend of permeability growth slowed down, indicating that the reactive fluid was confined to the dominant channel and did not react with other particles.

In the fractured sample, the original two fractures were clearly the preferential flow channels as expected, these fractures constantly widened during acid injection, while small branched channels developed at the outlet, Figure 8. However, the overall structure and shape of the pore space remained unchanged; thus, regions outside fractures hardly reacted with the  $\text{CO}_2$  acidified brine, consistent with the phenomena described in section 3.1.2. Note that  $\text{CO}_2$  acidified brine can weaken the limestone matrix mechanically (Lebedev et al., 2017). In fractured carbonate rock,  $\text{CO}_2$  acidified brine chose the original fractures as the preferential flow channels, so the original fractures are easier to be dissolved. We conclude that  $\text{CO}_2$  injection near geological faults increases the probability that faults will be reactivated.

Figure 9 shows the effective 3-D pore space ( $21.14 \mu\text{m}$  resolution; pore sizes larger than  $10^9 \mu\text{m}^3$  are shown) for the vuggy carbonate rock. Initially there were three large disconnected vugs, which connected after injection of 2,000 PV of the acidic brine. In contrast to the homogeneous sample, not many branched flow channels developed around this main flow channel. In fact, vuggy carbonate rock is highly heterogeneous; thus, our experimental result extends the application scope of the single channel dissolution pattern proposed by Menke et al. (2017).

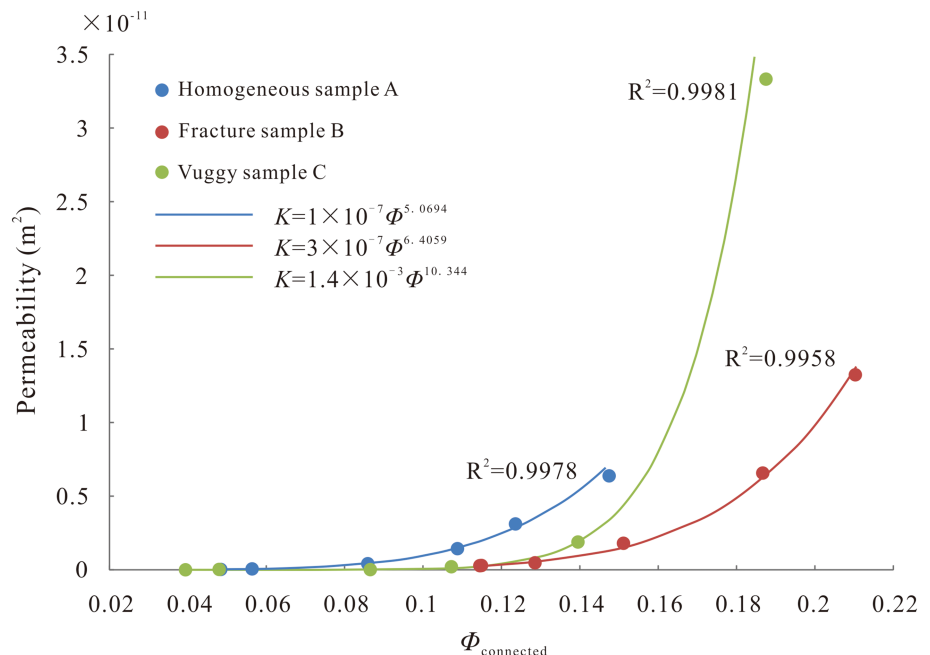
### 3.2.2. Porosity Changes and Dissolved Volume

Figures 10–12 depict the average porosity over each 2-D slice along the flow direction for the three carbonate rock samples. For homogeneous carbonate rock (Figure 10), the majority of the porosity increase occurred



**Figure 13.** Dissolved volume fraction calculated from micro-CT images versus core length. The dissolved volume of Sample A and C is calculated at 2,000 PV and that of Sample B at 500 PV.

during the first 1,000 PV injected, and porosity increased uniformly over the length of the core. However, for fractured carbonate rock (Figure 11), the porosity increase was lower where the initial porosity was low, and stronger where initial porosity was large. The porosity increase in vuggy carbonate rock was mainly localized in the entrance area (Figure 12) and increased continuously with injection time, consistent with the observations in Figure 9. We conclude that more attention should be paid to the damage around the



**Figure 14.** The power law relationships between permeability and connected porosity for the different samples studied. The porosity-permeability increase is dependent on dissolution pattern.

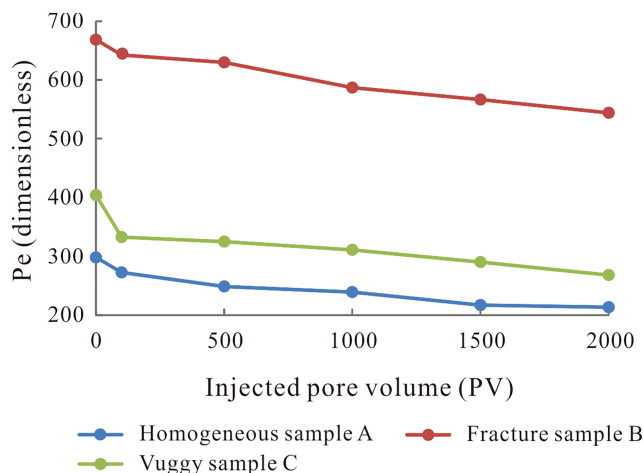
**Table 2**  
Parameters for All the Samples Over the Full Course of Experiments

Sample	Injected pore volume	$\phi_{CT}$	$S$ ( $m^{-1}$ )	$L$ (m)	$Pe$	$Da$ ( $10^{-5}$ )	$r_{eff}$ ( $mol/m^2 s$ )
A	0	0.068	28,082	$1.12 \times 10^{-4}$	298	3.07	0
	100	0.096	24,293	$1.29 \times 10^{-4}$	244	3.52	$5.23 \times 10^{-6}$
	500	0.131	19,744	$1.59 \times 10^{-4}$	220	4.35	$1.58 \times 10^{-6}$
	1,000	0.189	17,137	$1.83 \times 10^{-4}$	176	5.96	$1.41 \times 10^{-6}$
	1,500	0.219	15,544	$2.02 \times 10^{-4}$	167	7.57	$5.16 \times 10^{-7}$
	2,000	0.253	13,836	$2.27 \times 10^{-4}$	162	9.15	$4.71 \times 10^{-7}$
B	0	0.122	6,986	$4.50 \times 10^{-4}$	669	9.99	0
	100	0.131	6,773	$4.64 \times 10^{-4}$	642	10.84	$3.65 \times 10^{-6}$
	500	0.144	6,283	$5.00 \times 10^{-4}$	630	12.09	$1.12 \times 10^{-6}$
	1,000	0.168	5,778	$5.44 \times 10^{-4}$	587	14.52	$1.09 \times 10^{-6}$
	1,500	0.203	4,956	$6.34 \times 10^{-4}$	566	18.31	$1.18 \times 10^{-6}$
	2,000	0.229	4,574	$6.87 \times 10^{-4}$	544	21.35	$6.92 \times 10^{-7}$
C	0	0.079	17,863	$1.76 \times 10^{-4}$	404	12.79	0
	100	0.108	15,856	$1.98 \times 10^{-4}$	333	13.87	$7.25 \times 10^{-6}$
	500	0.127	13,802	$2.28 \times 10^{-4}$	325	15.48	$1.07 \times 10^{-6}$
	1,000	0.151	12,132	$2.59 \times 10^{-4}$	311	18.58	$7.47 \times 10^{-7}$
	1,500	0.18	10,909	$2.88 \times 10^{-4}$	290	23.44	$6.46 \times 10^{-7}$
	2,000	0.228	9,318	$3.37 \times 10^{-4}$	268	27.33	$8.84 \times 10^{-7}$

wellbore when injecting CO<sub>2</sub> into vuggy carbonate reservoirs, especially considering that the mechanical properties of limestone matrix are significantly weakened by CO<sub>2</sub> acidified brine.

These dissolution effects were further quantified; thus we analyzed the relationship between the dissolved volume and core length (Figure 13). These functions (at 2,000 PV) were linear for Sample A (homogeneous) and exponentially decreasing trend for Sample C (vuggy), while dissolved volume of Sample A varied little with core length, and dissolution was almost uniform throughout the core.

For fractured Sample B, the dissolved volume at 500 PV was a linear decay (Figure 13), which helps us to find a rule: Dissolution occurred mainly near the entrance when 500 PV reactive fluid was injected. Note that the dissolved volume at 2,000 PV was also analyzed, but there was no obvious correlation between it and core length. This irrelevancy was due to the rapid increase of the dissolved volume near the outlet in the later stage of dissolution (e.g., in Figure 11 the yellow curve of 2,000 PV, 100 μm from the outlet, you also find that the initial porosity of this location is larger). With the increase of CO<sub>2</sub> injection, the larger the initial porosity, the easier the dissolution occurred. Therefore, the dissolution of fractured carbonate rocks is not simply concentrated in the entrance area but related to initial porosity.



**Figure 15.** Péclet number ( $Pe$ ) versus the injected pore volumes.

This porosity evolution and dissolution rates can be explained by a higher reactive surface area in vuggy sample than in the homogeneous one. In fact, the initial permeability of Sample A was higher than for Sample C. This means that the connectivity was better in the homogeneous sample, which promotes the localization of the dissolution in the initial main branches, whereas in Sample C dissolution reaction should connect the vuggy pore together to increase the permeability and the dissolution rate is thus higher.

### 3.3. Porosity-Permeability Relationship

Time-dependent porosity-permeability relationships for the three samples were obtained via LBM (see section 2.2), Figure 14. These relationships can be fitted well ( $R^2$  of all samples is above 0.99) with a power law ( $K = a \phi_{connected}^b$ , where  $K$  is the dynamic permeability,  $\phi_{connected}$  is the porosity connected along the flow direction,  $a$  is a constant, and  $b$  is the power law exponent).



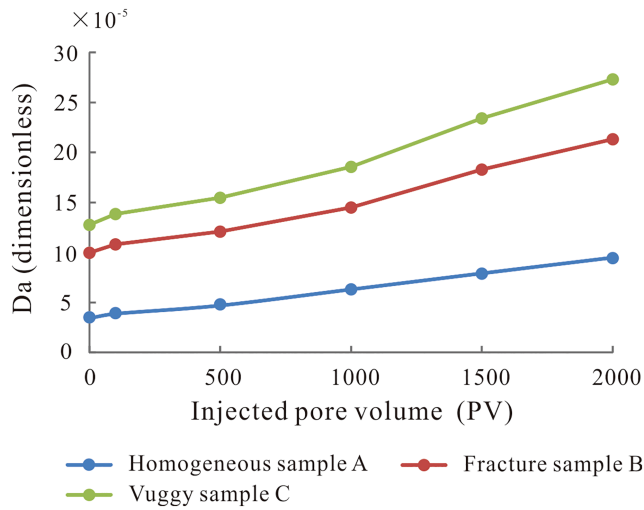


Figure 16. Damköhler number ( $Da$ ) versus injected pore volumes.

between acidified brine and rock particles after preferential channel is formed, which is manifested in a lower exponent (Al-Khulaifi et al., 2018; Luquot & Gouze, 2009; Menke et al., 2015; Smith et al., 2017).

### 3.4. Dynamic Péclet and Damköhler Number and Effective Reaction Rates

$Pe$  (the ratio of advection rate to diffusion rate) was calculated using the flow rate of 0.1 mL/min and the specific surface area obtained from the segmented images for the three samples (Table 2). In all experiments a high Péclet number prevailed; thus, fresh  $\text{CO}_2$ -acidified brine was readily available throughout the core since the spatially averaged advection rate was much higher than the  $\text{H}^+$  diffusion rate. Table 2 also contains  $Da$  and  $r_{eff}$  derived from porosity changes at different dissolution stages. In all experiments, the ratio of reaction rate to mass transfer rate ( $Da$ ) was significantly lower than 1, which indicates that the reaction rate was relatively slow when compared to both, the diffusion and advection rate. Furthermore, the initial effective reaction rate of vuggy Sample C was the highest, while it was lowest for the fractured Sample B.

Figure 15 depicts the evolution of Péclet numbers. As expected,  $Pe$  of Sample A (homogeneous, with high specific surface area) was lowest, while that of Sample B (fractured, with the lowest specific surface area) was the highest. For all samples,  $Pe$  decreased rapidly initially, but more gradually after 100 PV of injection. This indicates that there was a competition between preferential flow paths during channel formation,

which led to an advection-dominated transport in more pore space. Once the main channels were formed, the majority of the acidic brine flowed through them, and as a consequence the matrix pore space (outside the preferential channels) was dominated by diffusion. Furthermore, the high  $Pe$  number in Sample B indicates that advective transport in the fractured sample was strongest. Conversely,  $Pe$  of Samples A and C were smaller, which indicates a relatively strong impact of diffusion. The fluid is relatively stationary in the area where dissolution occurs, and there are obvious diffusive restrictions on transport and reaction (Menke et al., 2017).

$Da$  was always significantly higher than 1, and  $Da$  increased with the number of PV injected (Figure 16), consistent with the porosity increase, compare equation 7.  $Da$  of homogeneous Sample A was always lower than that of the other two samples, which indicates that the reaction rate is slower than the advection rate, consistent with homogeneous dissolution. Note that increasing  $Da$  reflects more nonuniform dissolution (Golfier et al., 2002), which is consistent with the preferential channel growth in homogeneous carbonate rock, compare section 3.2.1. With the increase of the pore volume, the competitive advantage of the preferential flow

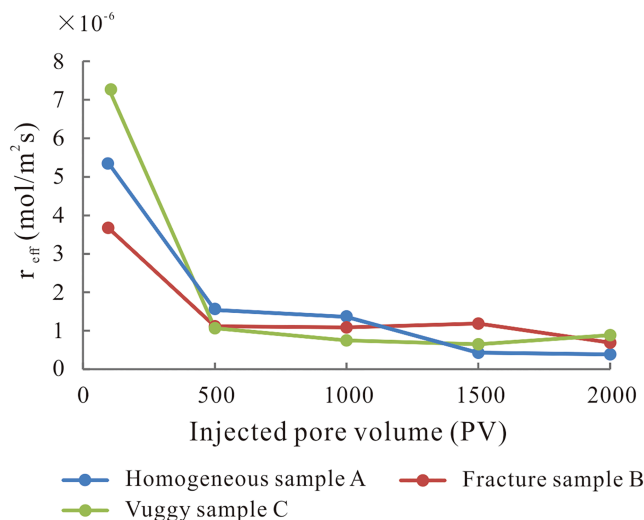
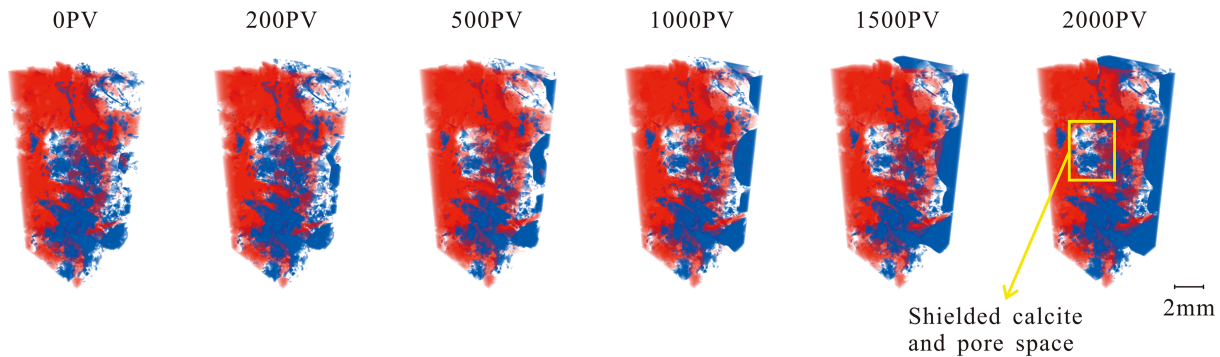


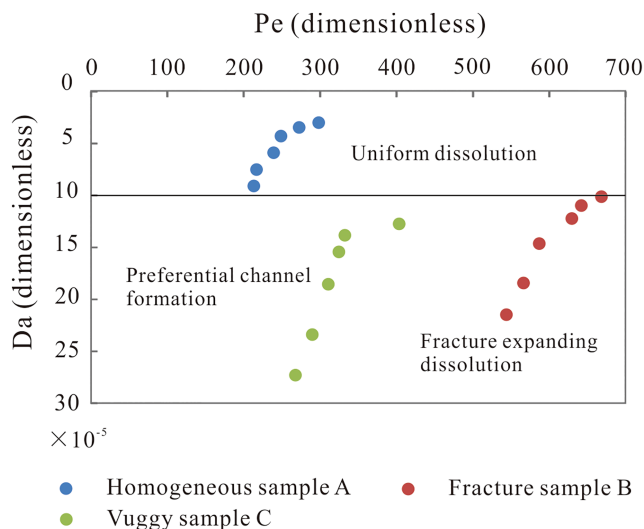
Figure 17. Effective reaction rates ( $r_{eff}$ ) versus injected pore volume.



**Figure 18.** Three-dimensional view of anhydrite (red) and pore space (blue) of vuggy Sample C. The blank part filled with calcite.

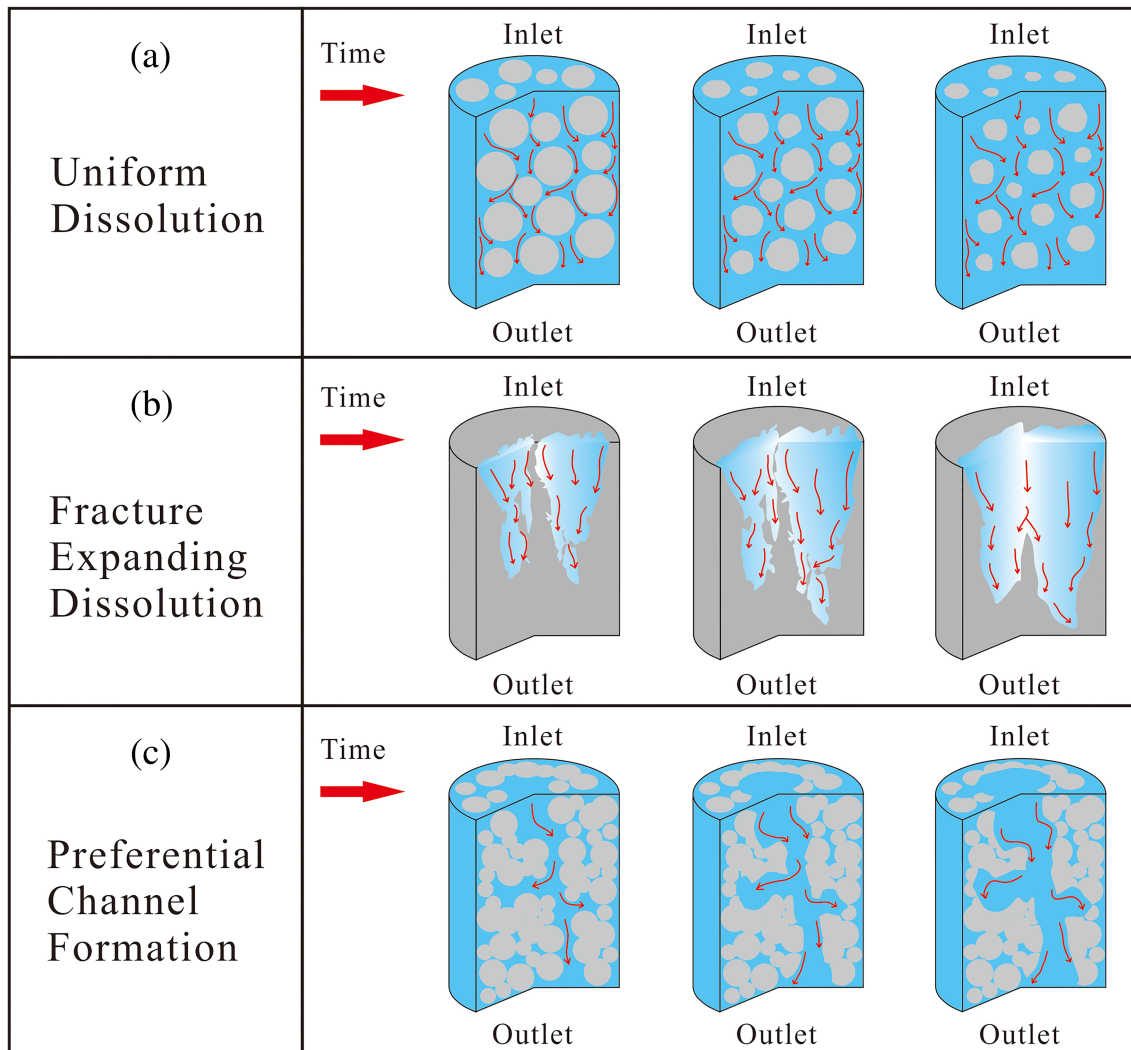
channel became more and more significant in the homogeneous Sample A; on the contrary, the branched channels hardly grew after 1,500 PV injection. For Sample B and Sample C, their  $Da$  remained at a high level, indicating that the dissolution is nonuniform from beginning to end, and the preferential channels dominate the flow in the early stage of dissolution.

The effective reaction rate ( $r_{eff}$ ) measured on the segmented images using equation 9 decreased with time at constant experimental conditions (Figure 17). We conclude that the preferential flow channels control the flow and lead to the decrease of the  $r_{eff}$ . Because the majority of acidic liquid passed through the center of preferential channels and reacted only with the channel walls without touching or reacting with other rock grains (Menke et al., 2017). In addition, the initial effective reaction rate of fractured Sample B was the lowest. Lower  $r_{eff}$  indicates that the reactive fluid cannot fully react with the rock grains. On the other hand, it can be understood that the fracture controls the flow from the beginning. Note that Al-Khulaifi et al. (2019) discussed the effect of chemical heterogeneity (they used a sample consisting of calcite and dolomite) on the effective reaction rate, and pointed out that the proximity of minerals to preferential flow channels determines the effective reaction rate and calcite was shielded by dolomite once the main flow channel formed. As our Sample C is also chemically heterogeneous (Table 1), we also extracted the spatial distribution of anhydrite and pore of vuggy sample during the dissolution process. Figure 18 represented only the area where the anhydrite is distributed, this is not representative of the entire sample (according to the results of mineral content analysis as shown in Table 1, the calcite content is more than 3 times the anhydrite content). Red represents anhydrite, blue is pore space, and the blank part is the filled calcite. Clearly, the anhydrite (on the left) shielded calcite from reacting, causing the preferential flow channel to localize on the right. Since the anhydrite hardly reacted with the  $CO_2$  acidified brine, the growth direction of the main flow channel changed four times from 2-D slices. Furthermore, we speculate that the shielding effect of anhydrite on calcite will limit the full core effective reaction rates.



**Figure 19.**  $Pe$ - $Da$  distribution diagram versus injected pore volume.

Different dissolution patterns correspond to different  $Pe$ - $Da$  combinations. It is obvious that the  $Pe$ - $Da$  distribution of three samples were different, and clustering of points was apparent, Figure 19. The  $Pe$ - $Da$  points were clustered around their basic dissolution pattern, but for all sets,  $Pe$  increased while  $Da$  decreased with increasing PV injected, consistent with Al-Khulaifi et al. (2017), Al-Khulaifi et al. (2018), Menke et al. (2016), Menke et al. (2017), and Menke et al. (2018). Golfier et al. (2002) provided the  $Pe$ - $Da$  distribution diagram for several dissolution patterns and found that uniform dissolution with the smaller the Damköhler number. Preferential channel formation and dissolution fracture expansion can be found in the range  $10^2 \leq Pe \leq 10^3$ . The demarcation line between uniform and nonuniform dissolution is  $Da = 10^{-4}$  for our three sets of experiments, which is different with Golfier's demarcation line (at  $Da \leq 8 \times 10^{-4}$ ; which was developed based on 2-D flat Hele-Shaw cells).



**Figure 20.** Conceptualized channel growth versus time during (a) uniform dissolution, (b) fracture expanding dissolution, and (c) preferential channel formation. Rock matrix (gray) is dissolved in advection-dominated regions (red arrows).

Note that the experimental results of Al-Khulaifi et al. (2018), Menke et al. (2017), and Menke et al. (2016) show that the  $Da$  value were distributed between  $1.88 \times 10^{-5}$  and  $9.28 \times 10^{-5}$  in the case of nonuniform dissolution, it is also beyond the range of nonuniform dissolution Goulet et al. (2002) offers. Actually, they simulated dissolution based on a simple 2-D flat model; while our study is based on real digital cores. Furthermore, a salt-filled transparent Hele-Shaw cell they used in experiment is different with natural limestones. We thus conclude that the range of Damköhler number in the same dissolution pattern is not constant, which is closely related to the complexity of pore geometry. As the complexity of the pore geometry increases, dissolution tends to be nonuniform.

### 3.5. Dissolution Pattern of Carbonate Rocks With Different Initial Pore Structure

The dissolution process for the three limestone samples were conventionalized in Figure 20. During uniform dissolution, the reactive fluid passes through pore space uniformly, thus most of the carbonate grains encounter and react with reactive fluid and eventually form multiple channels, rather than a single one. Such uniform dissolution can only occur when the pore structure of the rock is very uniform, that is, the size of the grains is similar and the degree of sphericity is very high (Menke et al., 2017). However, reservoir rock is not highly homogeneous. Furthermore, the preferential channel formed in Sample A is different from those formed in Samples B and C, where many surrounding branching channels were observed. In

fractured carbonate rock (Figure 20b), the reactive fluid flows through the original fractures so that the fractures extend and widen during dissolution, while the region far from the fractures hardly react. In a highly heterogeneous vuggy carbonate rock, the majority of the injected reactive fluid passes through the center of a fast flow path (largest vug); thus, it mainly reacts with the grains on the channel wall and a single preferential channel is formed. Moreover, the vugs around the preferential channel are further widened as they are also somewhat exposed to the acid.

#### 4. Conclusions

The physical heterogeneity has an important influence on limestone dissolution. Understanding the dissolution patterns of carbonate rocks with different pore structures can be applied to reservoir modeling and prediction of dissolution process. Based on a series of parametric analysis of micro-CT images, dissolution patterns of three different limestones were examined. Two-dimensional slices and 3-D images at different dissolution stages reveal that preferential channels were formed (or widened in case of fractures) throughout the whole core. Formation time and shape of the channels depended on the initial pore structure of the rock. These results were further quantified and compared with literature data. Following conclusions were drawn from the study:

1. The relatively homogeneous sample formed a large number of branched channels around a preferential main flow channel throughout the whole sample. However, after 1,500 PV of acidic brine were injected, the branched channels hardly grew, with the preferential channel now discharging most of the reactive fluid.
2. Fracture aperture increased continuously in the fractured carbonate rock, while the fracture wall smoothed. In addition, particles migrating downstream were observed, prove that the fracture is the main flow path of the acidic brine. However, the overall structure and shape of the pore space remained unchanged. Considering the weakening of limestone matrix by CO<sub>2</sub> acidified brine, the initial fractures were more likely to be dissolved. We should pay attention to reservoir instability caused by the reactivation of geological faults when injecting CO<sub>2</sub> near geological faults.
3. Channel growth was more complex in the vuggy limestone, while vugs grew in four directions successively. This growth was anisotropic in presence of anhydrite, which impeded dissolution (due to its lower chemical reactivity). In addition, formation of only one preferential channel was observed.
4. Porosity increased uniformly during the dissolution in the homogeneous sample, while porosity increase in the fractured sample was localized within the fractures themselves. In the vuggy sample, dissolution was dominant in the inlet area. We should pay more attention to wellbore instability caused by the weakening of near-wellbore matrix when injecting CO<sub>2</sub> into vuggy carbonate reservoirs.
5. The formation of a single preferential channel in the vuggy carbonate resulted in a sharp increase in permeability, which was manifested in the highest porosity-permeability power law exponent (10.344). In contrast in fractured and homogeneous carbonate, this exponent was smaller in uniform and fracture expanding dissolution patterns, permeability increased relatively slowly there.
6. We conclude that during channel formation, there is a competition between flow paths, which leads to an advection-dominated transport in more pore space. Once the preferential channels formed, the relative importance of diffusion became greater. This advection effect was strongest in the fractured sample, which had the highest Péclet number. Conversely,  $Pe$  of homogeneous (Sample A) and vuggy (Sample C) rocks were smaller, thus the fluid was relatively stationary in the area where dissolution occurred and there were diffusive restrictions on transport and reaction. The Damköhler number of homogeneous carbonate rock was always lower than that of the other two samples, which is required for homogeneous dissolution. The  $Da$  increase reflected the dissolution transformation from uniform to nonuniform. With the increase of the injected pore volume, the competitive advantage of the preferential channel became obvious. For fractured (Sample B) and vuggy (Sample C) rocks,  $Da$  remained high, indicating that the dissolution was always nonuniform.
7. Once preferential channels formed, the majority of reactive fluid passed through the center of these channels, which limited the effective reaction rate between the reactive fluid and the rock itself. Furthermore, the effective reaction rate of the fractured Sample B was the lowest initially, which illustrates that the fracture controlled the flow from the beginning.

8. The three dissolution patterns were analyzed in terms of their  $Pe$ - $Da$  distributions. The boundary between uniform and nonuniform dissolution was identified as  $Da = 10^{-4}$ . However, the dissolution tends to be nonuniform as the pore geometry complexity increases.

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