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HAL Id: hal-02488924
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Submitted on 6 Jul 2020

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Perfluoropolyalkylether Maleimides for Protection From Oxygen Inhibition and Surface Modification of Photoinitiator-Free UV-Cured Polymers

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Maleimides are attractive systems for photopolymerize for two major reasons: (1) they follow a radical mechanism without requiring a photoinitiator and (2) their rate of polymerization corresponds similarly to acrylates, which are commonplace in the industry. In this work, bismaleimide polypropylene oxide was cured under UV light forming thin films. Their surface properties were modified by copolymerization them with fluorinated comonomers. To this goal, perfluoropolyalkylethers (PFPAEs) with maleimide groups were synthesized, varying their chain structure, their functionality degree and consequently their intrinsic viscosity. These PFPPAE comonomers were highlighted to segregate at the surface, assuring omniphobic properties and acting as a protective layer against oxygen inhibition. These phenomenon were observed even when added at a concentration \( \leq 5\% \) w/w with respect to the main polypropylene oxide monomer. XPS analyses confirmed the segregation of the fluorine atoms at the surface during the UV-curing process of the coatings.

Keywords: UV-curing, maleimides, fluoropolymer, perfluoropolyalkylether, surface properties

INTRODUCTION

Photoinduced polymerizations are green processes suitable for the sustainable synthesis of polymers (Bowman and Kloxin, 2008; Yagci et al., 2010; Corrigan et al., 2019). In fact, they are bulk processes (without solvent) which require low amount of energy as they occur at room temperature. Indeed these photoinduced reactions are very fast: the reaction time for achieving quantitative conversion is of the order of seconds or minutes. Most photoinduced polymerizations are chain reactions, initiated by a proper photoinitiator able to absorb light, reach an excited state and originate reactive species; if the functionality of the monomers is \( \geq 2\), cross-linked networks are formed. Thus, from a liquid resin, a solid material can be easily obtained after a few seconds of light irradiation. Currently, UV and LED sources are used for these methods but the amount of curing depends on the absorbance of light, thus monomers are preferably...
processed in the form of thin films. Consequently, since the start of its industrial development, UV-curing finds application in fields such as paints, adhesives or protective coatings. At the present time, cutting-edge technologies such as 3D-printing and medical implants benefit from these processes: in fact they can be controlled in time and space as the monomers react only in the irradiated area and the polymerization stops when the light is switched off (Ligon et al., 2017). Commonly, monomers employed in the photoinduced processes used for the above applications, are acrylic, methacrylic and thiol-ene systems (Hoyle and Bowman, 2010; Crivello and Reichmanis, 2014; Du et al., 2017a,b). Maleimides represent an attractive alternative as they polymerize almost as fast as acrylates. Their photopolymerization in bulk and in solution was reported for the first time by Yamada et al. (1968); the mechanism of initiation was found to be based on an electron transfer able to provide a radical anion (von Sonntag et al., 1999; von Sonntag and Knolle, 2000). In the last decade, homopolymerization under UV-light of maleimides containing different backbones such as ethylene glycol (Vázquez et al., 2009), perfluoroalkyl chains (Soules et al., 2008), or silicones (Pozos Vázquez et al., 2010) was explored. Results confirmed that the reaction is very efficient even in the absence of a photoinitiator (Decker et al., 2004; Vázquez et al., 2009). This is of major relevance in fields where photopolymers are applied, as the presence of photoinitiators often poses safety issues due to their potential leaching. For this reason, maleimides are also proposed as photoinitiators for different types of unsaturated resins such as acrylates, vinyl ethers, etc. (Hoyle et al., 1997; Jönsson et al., 1997).

As previously mentioned, polymers obtained by photopolymerization often come in the form of film and are used as coatings, adhesives, inks where performance depends most importantly on surface properties (i.e., adhesion, wettability, self-cleaning, printability etc.). Tailoring the surface of UV-cured polymers while keeping the bulk properties unchanged is possible by introducing a small amount of a surface-active comonomer in the reactive formulation (Vitale et al., 2018; Wasser et al., 2018). This approach has been widely applied to acrylic, vinyl ether, epoxy systems using fluorinated comonomers (Vitale et al., 2015; Wasser et al., 2018). In this work, we describe the influence of fluorinated maleimides for the modification of photoinitiator-free maleimides-based coatings. In particular, three new fluorinated maleimides were prepared by functionalization of perfluropolyalkylether chains (PFPAEs) containing one or two of the following repeating units –(CF₂O)–, –(CF₃CF₂O)–, and –(CF(CF₃)CF₂O)–. These highly fluorinated building blocks were chosen as they are both hydro- and oleo-phobic. Moreover, they are known as non-toxic (Ameduri and Boutevin, 2004; Friesen and Ameduri, 2018) with respect to perfluoroalkyl chains (Wang et al., 2015) so that they can be employed in food applications and even in biomedical applications (Malinverno et al., 1996; Panatti, 2008; Zhang et al., 2018). The effect of mono- and bi-functional fluoromaleimides on the photopolymerization of a commercial bismaleimide polypropylene oxide was investigated. Then, their impact on the polymerization kinetics and on the surface properties of the coatings was examined.

### MATERIALS AND METHODS

#### Materials

The commercial bismaleimide poly(propylene oxide) (Mₚ∼400 g/mol) called BMI PPO was purchased from Specific Polymers (France).

The fluorinated monomaleimide MMI (see structure in Figure 1) was synthesized as previously reported (Bonneaud et al., 2019) by esterification of Krytox® Methylen Alcohol, kindly provided by Chemours Company (Wilmington, Delaware, USA).

Two fluorinated bismaleimides, labeled BMI C5 and BMI C10 (see structure in Figure 1), were synthesized starting from a commercial perfluoropolyalkylether diol Fluorolink® E10H (Mₚ≈1,800 g/mol) from Solvay Solexis (Bollate, Italy), purchased from Acotta (UK). The reaction was carried out using either 6-maleimidohexanoic acid (to obtain BMI C5) or 11-maleimidoundecanoic acid (to obtain BMI C10), 6-maleimidohexanoic acid, 11-maleimidoundecanoic acid, diclohexylcarbodiimide, dimethylamino pyridine, dichloromethane, hexadecane, and chloroform were purchased from Sigma Aldrich (St. Louis, Missouri, USA) and used as received. CsD₆ was purchased from Eurisotop (Cambridge, UK) and used in capillaries for NMR analyses. 1,1,1,3,3-pentafluorobutane was purchased from Alfa Aesar (Ward Hill, Massachusetts, USA).

#### Synthesis and Characterization of the Fluorinated Maleimides

The monomaleimide MMI was synthesized following a previous procedure (Bonneaud et al., 2019).

In a typical procedure, to a dichloromethane solution of 6-maleimidohexanoic acid or 11-maleimidoundecanoic acid (2.05 eq), dimethylaminopyridine (0.2 eq) and Fluorolink® E10H (2 eq) in 1,1,1,3,3-pentafluorobutane were added dropwise at 0°C. After 5 min, the ice bath was removed. After 1.5 h, the reaction was stopped. A first filtration was performed and the solvents were partially removed. A liquid-liquid extraction was performed using both 1,1,1,3,3-pentafluorobutane and saturated NaHCO₃ (aq). The fluorinated phase was then washed with saturated NaHCO₃ (aq) (2 ×) and brine (1 ×). A second filtration through silica/Celite® was performed. Hydroquinone was added (0.1% w/w) and the solvent was removed.

BMI C5 yield was 73%; BMI C10 yield was 21%. The products were characterized by ¹H NMR, ¹³C NMR, and FT-IR (ATR). NMR spectra were recorded on a Bruker AVANCE III 400 MHz spectrometer instruments using TopSpin 3.5 operating at 400.13 (¹H), 376.46 (¹³C), 100.62 (¹⁵N) MHz at room temperature except if specified. CsD₆ capillaries were used as internal references. The letters s, d, t, q, quint, sext, and spt stand for singlet, doublet, triplet, quartet, quintuplet, sextet, and septuplet, respectively.

BMI C5 ¹H NMR (400 MHz, CsD₆, 25°C):

δ = 1.60 (m, -N(CH₂)₂CH₂CH₂(OH)₂, 2H), 1.87 (m, -NCH₂CH₂CH₂CH₂CH₂CH₂(OH)₂, 4H), 2.57 (q, JH-H = 7.4 Hz, -CH₂CH₂CH(OH)₂, 2H), 3.77 (t, JH-H = 7.2 Hz, -NCH₂-, 2H), 3.93-4.48 (m, -C(CH₂)₂OCH₂CH₂CH₂(OH)₂, 9H), 6.88 (s, -C(O)CH=CH(OH)₂, 2H)
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Preparation and Methods of Characterization of the Coatings

Photopolymerization kinetics were monitored by Real-time Infrared Spectroscopy performed on a Thermo Scientific Nicolet 6700 FTIR apparatus using attenuated total reflectance (ATR) methods controlled by OMNIC software. The conversion was monitored by following the disappearance of the vibrational band at 826 cm\(^{-1}\) corresponding to the maleimide group; the conversion \(\tau\) was estimated through the univariate method (\(\tau = (1-(A/A_0))^n\)). A mercury lamp (OmniCure S2000) was used as UV-light source. An OmniCure R2000 Radiometer was used to control the light output: the intensity at the sample surface was 10 mW/cm\(^2\). One drop of sample, with no photoinitiator, was deposited on the ATR unit of the infrared spectrometer (thickness \(\sim 10-30 \mu m\)) for each measurement. A polypropylene film (6 \(\mu m\) ) was used as air protector and each experiment was conducted four times.

For the fabrication of films, the reactive monomers (without photoinitiator) were poured onto glass slides previously cleaned with acetones. The coating of the films was done by bar coating (thickness around 25 \(\mu m\)) with a film applicator bar coater (Sheen 1117/100 mm). Irradiation was performed for 5 min with a mercury lamp (Dymax) with an intensity of 150 mW/cm\(^2\). One of these films (6 \(\mu m\) ) was used as air protector and each experiment was conducted four times.

The gel content was determined as \([w_1/w_0] \times 100\), where \(w_0\) is the initial weight of the sample (mg) and \(w_1\) is the weight after extraction (mg). The extraction solvent was chloroform for the BMI PPO-based polymers and 1,1,1,3,3-pentafluorobutane for the fully fluorinated coatings. The films were weighed and then immersed in the extracting solvent for 24 h (three repeats). After the extraction, the films were dried at 45°C under vacuum (10\(^{-2}\) mbar) for 24 h and reweighed.

Thermogravimetric analyses (TGA) were carried out with a NETZSCH TG209F1 apparatus, at a heating rate of 20°C min\(^{-1}\). Approximately 10 mg of sample were placed in an alumina crucible and heated from room temperature to 600°C under air (40 mL min\(^{-1}\)).

Differential scanning calorimetry (DSC) was used to determine the glass transition temperatures of the polymers, employing a NETZSCH DSC200F3 calorimeter. Constant
calibration was performed using indium, n-octadecane and n-octane standards. Ten to fifteen mg of sample were placed in an aluminum pan and the thermal properties were recorded between −150 and 100°C at 20°C min−1. Nitrogen was used as the purge gas.

Contact angle measurements were performed by using a contact angle system OCA20 coupled with a CCD-camera from DataPhysics Instrument using the software SCA20 4.1. The measurements were made in air at room temperature by the sessile drop technique 10 s after the deposition of the drop. Distilled water and hexadecane previously dried were the testing liquids. Between three and six repeats were made on three different samples previously irradiated. Their difference in the average value was no more than 3°.

X-ray photoelectron emission spectra were recorded using a monochromatised Al Kalpha (hv = 1486.6 eV) source on a ThermoScientific K-Alpha system. The X-ray Spot size was about 400 μm. The Pass energy was fixed at 20 eV with a step of 0.1 eV for core levels and 150 eV for surveys (step 1 eV). The spectrometer energy calibration was done using the Au 4f7/2 (83.9 ± 0.1 eV) and Ag3d5/2 (368.2 ± 0.1 eV) photoelectron lines. XPS spectra were recorded in direct mode N (Ec).

The viscosities of BMI C5 and BMI C10 were measured at 25°C on the AR-1000 rheometer (TA Instruments). A 25 mm diameter and 4° cone-plan geometry were used. The flow mode was used with a gradient from 10 to 0.1 s−1 and then the steady flow mode from 0.1 to 10 s−1. The viscosity value is the average value over 15 points of the "plateau."

RESULTS AND DISCUSSION

Photocopolymerization of Maleimides

Different maleimides containing a perfluoropolyalkylether chain (PFPAE) and possessing a different functionality were synthesized: the chemical structure of the monomers is reported in Figure 1.

The monofunctional fluorinated monomaleimide (MMI) containing hexafluoropropyleneoxide units was synthesized as reported in a previous paper (Bonneaud et al., 2019). Two telechelic difunctional bismaleimides (BMI C5 and BMI C10) were obtained by esterification reaction of maleimidocarboxylic acids of different molecular weight with a commercial fluorinated diol HO-RF-OH, where RF (reported in Figure 1) is a fluorinated chain based on −(CF₂O)−, −(CF₂CF₂O)−. Therefore, the fluorinated BMI are characterized by the same perfluoropolyalkylether structure, and have different alkyl spacers −(CH₂)n−, where n = 5 or 10 for BMI C5 and BMI C10, respectively.

First, the neat fluorinated maleimides were homopolymerized. As already mentioned, maleimides can initiate the polymerization on their own according to the mechanism of Figure 2, where the mechanism of initiation proposed is an electron transfer which provides a radical anion (von Sonntag et al., 1999; von Sonntag and Knolle, 2000).

MMI reaction was studied previously (Bonneaud et al., 2019): it was reported that MMI is a highly reactive monomer with a quantitative conversion observed even in 8 s in the absence of air and in 40 s in the presence of air (see also Table 1).

Having low viscosity, it could be easily coated on glass as a film and under irradiation exhibited high reactivity, giving rise to a linear polymer. As expected, also BMI C5 and BMI C10 were reactive even in the absence of a photoinitiator. The homopolymers obtained from the bifunctional fluorinated maleimides had quantitative conversion as indicated by the high gel content values (Table 1). However, the viscosity of BMI C10 was too high and did not allow a uniform coating of substrates, i.e., it is unsuitable for the preparation of coatings through a bulk process.

Therefore, the fluorinated monomers were preferably used in copolymerization with BMI PPO at a low concentration, i.e., 2 w/w% and 5 w/w% resulting in crosslinked polymers in the form of transparent films, 25 μm thick.

The photocopolymerization kinetics of the comonomeric mixtures of BMI PPO with each fluoromonomer, were monitored by Real-time Infrared Spectroscopy, by following the

![Figure 2](image-url)
disappearance of the band at 826 cm$^{-1}$ corresponding to the maleimide group. No photoinitiator was added. The irradiation was done both in air and in absence of air: protection from air was obtained covering the monomers with a polypropylene sheet according to common practice. The conversion data vs. time are plotted in Figure 3, comparing the reactions with the homopolymerization of BMI PPO.

Neat BMI PPO provided crosslinked films with a high gel content (Table 1). On Figure 3A, BMI PPO demonstrated that its photopolymerization kinetics was found to be faster combined with a higher yield in the absence of air. In the presence of air, the rate was significantly lower. Indeed, the final conversion never reached the value obtained in the absence of air. Thus, after 5 min of irradiation its value was 91.9% instead of 96.5%.

Figure 3A shows the influence of MMI on the conversion curves. In the absence of air, the addition of MMI slightly slowed down the reaction, the initial rate was lower while the final conversion was higher than the neat BMI PPO, and could be reached at a longer irradiation time. This reduced conversion rate could be ascribed to the lower average functionality of the system as MMI is monofunctional. Being MMI monofunctional, it could however enhance the mobility of the system, thus leading to higher conversion degrees (from 96.5% without MMI to 97.3% with 5 w% of MMI after 5 min of UV irradiation, as shown in Table 1). It is worth reporting that the system prior irradiation appeared white and opaque, indicating heterogeneity, while after an exposure of 5 min under UV-light, transparent coatings were formed. Thus heterogeneity was initially present and could have an effect on initial reaction rate: the phase separation may delay the photopolymerization process due to lower light absorption and hindering of reactive species diffusion at/though the interface.

Concerning the copolymerization in the presence of air, the final conversion was always improved by the fluorinated comonomer, as in the case of the reaction performed in the absence of air: from 91.9% without MMI to 96.2% with 5 w% of MMI after 5 min of UV irradiation, as shown in Table 1. More interestingly, in the presence of air, the addition of MMI enhanced the speed rate: the higher the percentage of MMI additive, the higher the initial rate. As the photopolymerization kinetics are slower in the presence of air due to the quenching of radicals by oxygen (Ligon et al., 2014), the addition of MMI seems to contrast this inhibition phenomenon. It is known that the fluorinated phase of methacrylate PFPAE with hydrogenated
copolymers segregate at the air side of UV-cured films and PFPAEs are able to dissolve oxygen (Vitale et al., 2013). Thus, it was expected that the fluorinated phase from MMI would migrate toward the air side, causing a surface segregation of the fluorinated domains. The fluorine-enriched phase on the air side then behaves as a protective layer against oxygen inhibition by protecting the reactive maleimide phase from the air. As a result, thanks to the surface segregation, the reaction rate was improved. Moreover, due to the maleimide substituent attached to the fluorinated chain, the fluorinated protective layer is covalently bonded to the network and cannot leak over time.

The photopolymerization kinetics profile of BMI PPO in the presence of the fluorinated additives having more than one functionality in their backbone, BMI C5 and BMI C10, was also studied. In Figure 3B, the effect of BMI C10 was examined on the polymerization conversion (similar trends were found for BMI C5). In the absence of air, both BMI C5 and BMI C10 revealed to slow the kinetics of the commercial BMI PPO, even when added in low amount. Nonetheless, the final conversion of the three systems was very similar (Table 1). In the presence of air, BMI C10 was also found to act as a protection against oxygen inhibition as discussed before for MMI, by significantly improving the photopolymerization rate of BMI PPO (Figure 3B). Similar results would be expected for BMI C5 with a speed increase when added in comparison to the neat BMI PPO product in the presence of air.

Figure 4 compares in details BMI C5 and BMI C10 when added to BMI PPO in different amount and cured in the absence of air. The change in reaction rate and conversion over an irradiation time of 30 s is negligible if the concentration of the fluorinated comonomer is 2 wt%. When the concentration increases to 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%. When the concentration of the fluorinated comonomer is 2 wt%. When the concentration is 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%. When the concentration increases to 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%. When the concentration increases to 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%. When the concentration increases to 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%. When the concentration increases to 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%. When the concentration increases to 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%. When the concentration increases to 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%. When the concentration increases to 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%. When the concentration increases to 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%. When the concentration increases to 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%. When the concentration increases to 5 wt%, the conversion showed a slower kinetics for the fluorinated comonomer is 2 wt%.

**Characterization of the UV-Cured Copolymeric Coatings**

The thermal properties of the UV-cured (co)polymers were assessed by TGA and DSC analyses. As reported in Table 2, the thermal stability of these polymers was very good thanks to the maleimide group. The temperature corresponding to 5% of weight loss (T5%) for the polymer obtained from the commercial resin BMI PPO was 381°C. In comparison, the fluorinated homopolymers MMI, BMI C5, and BMI C10, respectively had a T5% of 336, 326, and 325°C. In general, a decrease of the T5% value was observed when the coatings contained the fluorinated additives, however, its value was always 340°C or higher.

Concerning the glass transition temperature of the systems, the fluorinated homopolymers based on the difunctional BMI C5 and BMI C10 exhibited two Tg's (Table 2): the one at low temperatures, around −100°C, corresponded to the fluorinated phase (TgF) and the one above room temperature corresponded to the hydrogenated phase (TgH). The neat MMI polymer showed only one Tg at −66°C. These results suggest that the PFPAE chain present in MMI, characterized by pending –CF3 groups, has a significant effect on the glass transition temperature.

<table>
<thead>
<tr>
<th>Sample</th>
<th>BMI PPO (wt%)</th>
<th>PFPAE additive (wt%)</th>
<th>T5% (°C)</th>
<th>TgF (°C)</th>
<th>TgH (°C)</th>
</tr>
</thead>
<tbody>
<tr>
<td>BMI PPO – Neat</td>
<td>100</td>
<td>0</td>
<td>381</td>
<td>n/a</td>
<td>36</td>
</tr>
<tr>
<td>BMI PPO + MMI 2%</td>
<td>98</td>
<td>2</td>
<td>379</td>
<td>n/a</td>
<td>40</td>
</tr>
<tr>
<td>BMI PPO + MMI 5%</td>
<td>95</td>
<td>5</td>
<td>377</td>
<td>n/a</td>
<td>37</td>
</tr>
<tr>
<td>MMI – Neat</td>
<td>0</td>
<td>100</td>
<td>336</td>
<td>−66</td>
<td>n/a</td>
</tr>
<tr>
<td>BMI PPO + BMI C5 2%</td>
<td>98</td>
<td>2</td>
<td>379</td>
<td>n/a</td>
<td>61</td>
</tr>
<tr>
<td>BMI PPO + BMI C5 5%</td>
<td>95</td>
<td>5</td>
<td>340</td>
<td>n/a</td>
<td>49</td>
</tr>
<tr>
<td>BMI C5 – Neat</td>
<td>0</td>
<td>100</td>
<td>326</td>
<td>−100</td>
<td>63</td>
</tr>
<tr>
<td>BMI PPO + BMI C10 2%</td>
<td>98</td>
<td>2</td>
<td>357</td>
<td>n/a</td>
<td>62</td>
</tr>
<tr>
<td>BMI PPO + BMI C10 5%</td>
<td>95</td>
<td>5</td>
<td>360</td>
<td>n/a</td>
<td>56</td>
</tr>
<tr>
<td>BMI C10 – Neat</td>
<td>0</td>
<td>100</td>
<td>325</td>
<td>−105</td>
<td>68</td>
</tr>
</tbody>
</table>

**TABLE 2** | Thermal properties and gel content values of the different mixtures of BMI PPO and the fluorinated additives MMI, BMI C5, and C10.

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a Tg of the fluorinated phase; b Tg of the hydrogenated phase.

![FIGURE 4](https://example.com/figure4.png) | Effect of the amount of BMI C5 and BMI C10 on the photopolymerization conversion curves in absence of air.
restricted mobility, compared to the linear PFPAEs chain present in the BMI monomers.

As shown in Table 2, for all the copolymers, only one glass transition temperature (T_g) was detected. T_g was not influenced by the addition of MMI to BMI PPO, even though a slight increase of mobility, due to the pending fluorinated chains bonded only on one side to the polymer network, would be expected. Whereas, T_g was increased by the addition of the less mobile multifunctional additives BMI C5 and BMI C10, which could be due to an enhanced crosslinking of the system.

Contact angle measurements with water and hexadecane (θ_water and θ_hexadecane, respectively) were performed to determine the surface properties of the coatings (Table 3). After detachment from the substrate, measurements were made on the upper side (air side) and on the side in contact with glass (glass side). The commercial BMI PPO polymer showed a water contact angle of around 79° at the air side and a hexadecane contact angle of 21°.

The presence of fluorinated additive had a strong impact on the BMI PPO surface properties. Analyzing the results reported in Table 3 for the different copolymers, two mains features can be detected. Firstly, the addition of MMI, BMI C5 or BMI C10 increased the contact angle values with water and hexadecane on the air side: values were as high as 109° and 64°, respectively. The fluorinated monomers, although in low amount, imparted omniphobic properties to the coatings on the air side. Secondly, the contact angle values were always higher on the air side than on the glass side, either for water and hexadecane. This phenomenon was assumed to come from the segregation of the fluorinated phase on the air side due to its very poor affinity to glass and its good affinity with air, as reported for other UV-cured systems containing fluorinated comonomers (Vitale et al., 2015). No main difference between the fluorinated additives could be highlighted, as they all led to comparable contact angle values.

XPS analyses (Table 4) confirmed the segregation of the fluorinated monomers at the air surface as commonly found in literature (Casazza et al., 2002; Hu et al., 2008; Bongiovanni et al., 2012). Indeed, the percentage of fluorine was found to be much higher on the air side than on the glass side for the copolymeric coatings. Moreover, while the percentages of the fluorinated additives were low, i.e., around 0.01 mol% for the copolymers containing 5 wt% of comonomer, at the air surface the amount of fluorine greatly exceeded the calculated values (Table 4). An example, the sample containing MMI has a 100-fold increase of the atomic percentage of fluorine at the air surface. The values of the fluorine atomic percentage of the pure fluorinated comonomers, calculated on the basis of their chemical structure (Figure 1), are 51, 35, and 31% for MMI, BMI C5 and BMI C10, respectively. Therefore, it is demonstrated that at the air surface there is a fluorinated layer, which contains nearly exclusively the PFPAE comonomer. These results clearly explain the high contact angle values with water and hexadecane on the air side and the protection from oxygen inhibition.

### CONCLUSIONS

Different fluorinated PFPAE-based maleimides were synthesized with different degree of functionality. They were employed at a low content as a comonomer in the photopolymerization of bismaleimide poly(propylene oxide) resin (BMI PPO). The fluorinated maleimides displayed an interesting behavior under...
UV-light: in the presence of air they acted as a protective layer against oxygen inhibition, allowing an increase of the photopolymerization rate. The final coatings exhibited a good thermal stability due to the maleimide structure as well as an omniphobic behavior. Indeed, the surface properties of BMI PPO were modified thanks to the surface segregation of the fluorinated additives. These results are very encouraging in view of industrial application to overcome the oxygen inhibition effect and at the same time tune the surface properties of coatings.

DATA AVAILABILITY STATEMENT

The datasets generated for this study are available on request to the corresponding author.

AUTHOR CONTRIBUTIONS

JB was involved in the synthesis of the monomers. CB participated to the synthesis and investigated the photopolymerization of the new monomers as well as their characterization. GT and AV also did the characterization of the polymers. CJ-D led the conception of the work. CF and RB supervised the work as part of a H2020 project, contributed to the data analysis and interpretation. RB and CB prepared the manuscript. All authors contributed to its revision, read, and approved the submitted version.

FUNDING

This research project has received funding from the European Union’s Horizon 2020 research and innovation program under grant agreement No. 690917—PhotoFluo; Natural Sciences and Engineering Research Council of Canada (NSERC), Discovery Grants Program RGPIN-2015-05513 and Ministère de l’Enseignement Supérieure et de la Recherche. We would also like to thank Valérie Flaud and Jérôme Esvan for their efforts concerning the XPS analyses.

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Conflict of Interest: The authors declare that the research was conducted in the absence of any commercial or financial relationships that could be construed as a potential conflict of interest.

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