



HAL
open science

Electrochemical advanced oxidation processes using novel electrode materials for mineralization and biodegradability enhancement of nanofiltration concentrate of landfill leachates

Marwa El Kateb, Clément Trelu, Alaa Darwich, Matthieu Rivallin, Mikhael Bechelany, Sakthivel Nagarajan, Stella Lacour, Nizar Bellakhal, Geoffroy Lesage, Marc Heran, et al.

► **To cite this version:**

Marwa El Kateb, Clément Trelu, Alaa Darwich, Matthieu Rivallin, Mikhael Bechelany, et al.. Electrochemical advanced oxidation processes using novel electrode materials for mineralization and biodegradability enhancement of nanofiltration concentrate of landfill leachates. *Water Research*, 2019, 162, pp.446-455. 10.1016/j.watres.2019.07.005 . hal-02278696

HAL Id: hal-02278696

<https://hal.umontpellier.fr/hal-02278696>

Submitted on 25 Oct 2021

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.



Distributed under a Creative Commons Attribution - NonCommercial 4.0 International License

Electrochemical advanced oxidation processes using novel electrode materials for mineralization and biodegradability enhancement of nanofiltration concentrate of landfill leachates.

Marwa El Kateb^{1,2,4}, Clément Trellu^{1,3,*}, Alaa Darwich¹, Matthieu Rivallin¹, Mikhael Bechelany¹, Sakthivel Nagarajan¹, Stella Lacour¹, Nizar Bellakhal⁴, Geoffroy Lesage¹, Marc Héran¹, Marc Cretin^{1,*}

¹ IEM, Univ Montpellier, CNRS, ENSCM, Montpellier, France

² Université de Tunis El Manar, Faculté des Sciences de Tunis, 2092 Tunis, Tunisie

³ Laboratoire Géomatériaux et Environnement, LGE – Université Paris-Est, EA 4508, UPEM,
77454 Marne-la-Vallée, France

⁴ Université de Carthage, Institut National des Sciences Appliquées et de Technologie,
Laboratoire d’Echo-Chimie, 1080 Tunis, Tunisie

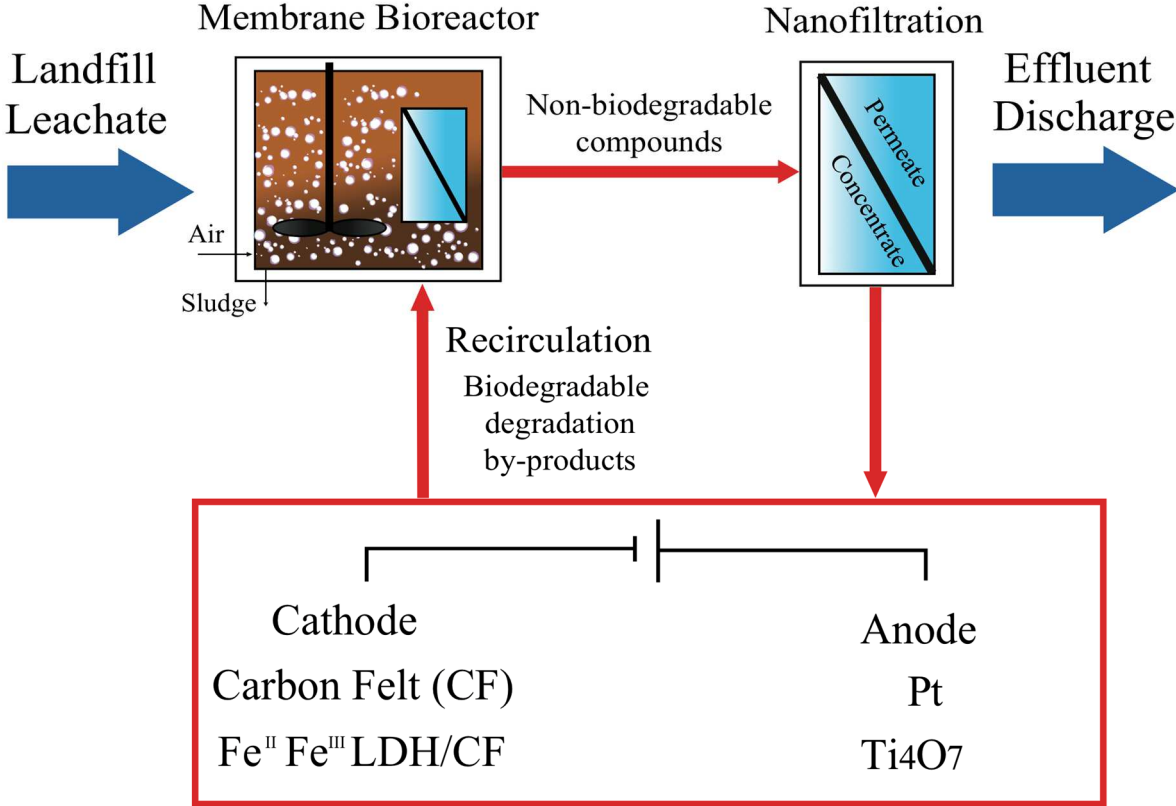
Manuscript submitted to Water Research for consideration

* Corresponding Author:

clement.trellu@u-pem.fr

+33 1 49 32 90 42

Graphical abstract



Electrochemical Advanced Oxidation Processes
Homogeneous/Heterogeneous electro-Fenton
Anodic Oxidation

1 **Abstract**

2 The objective of this study was to implement electrochemical advanced oxidation processes
3 (EAOPs) for mineralization and biodegradability enhancement of nanofiltration (NF)
4 concentrate from landfill leachate initially pre-treated in a membrane bioreactor (MBR). Raw
5 carbon felt (CF) or Fe^{II}Fe^{III} layered double hydroxides-modified CF were used for comparing
6 the efficiency of homogeneous and heterogeneous electro-Fenton (EF), respectively. The
7 highest mineralization rate was obtained by heterogeneous EF: 96% removal of dissolved
8 organic carbon (DOC) was achieved after 8 h of electrolysis at circumneutral initial pH (pH₀
9 = 7.9) and at 8.3 mA cm⁻². However, the most efficient treatment strategy appeared to be
10 heterogeneous EF at 4.2 mA cm⁻² combined with anodic oxidation using Ti₄O₇ anode (energy
11 consumption = 0.11 kWh g⁻¹ of DOC removed). Respirometric analyses under similar
12 conditions than in the real MBR emphasized the possibility to recirculate the NF retentate
13 towards the MBR after partial mineralization by EAOPs in order to remove the residual
14 biodegradable by-products and improve the global cost effectiveness of the process. Further
15 analyses were also performed in order to better understand the fate of organic and inorganic
16 species during the treatment, including acute toxicity tests (Microtox[®]), characterization of
17 dissolved organic matter by three-dimensional fluorescence spectroscopy, evolution of
18 inorganic ions (ClO₃⁻, NH₄⁺ and NO₃⁻) and identification/quantification of degradation by-
19 products such as carboxylic acids. The obtained results emphasized the interdependence
20 between the MBR process and EAOPs in a combined treatment strategy. Improving the
21 retention in the MBR of colloidal proteins would improve the effectiveness of EAOPs
22 because such compounds were identified as the most refractory. Enhanced nitrification would
23 be also required in the MBR because of the release of NH₄⁺ from mineralization of refractory
24 organic nitrogen during EAOPs.

25 **Keywords**

26 Electro-Fenton; Anodic oxidation; Modified carbon felt; Sub-stoichiometric titanium oxide;

27 Landfill leachate; Biodegradability.

28 **1. Introduction**

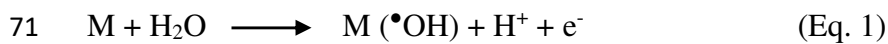
29 Rainwater percolation through waste layers of landfills generates leachates containing a
30 complex mixture of dissolved organic matter (DOM), inorganic compounds, heavy metals,
31 and xenobiotic organic substances (Kjeldsen et al., 2002), which represents a significant
32 hazard for the environment.

33 The implementation of a membrane bioreactor (MBR) followed by a nanofiltration (NF) step
34 is one of the most efficient treatment strategy currently used for management of landfill
35 leachates (Campagna et al., 2013; Amaral et al., 2016). However, NF is only a separation
36 process. Biorefractory organic pollutants are accumulated and concentrated. Thus, the
37 concentrate becomes an important residual issue for this treatment strategy (Van der Bruggen
38 et al., 2003; Zhang et al., 2009, 2013). Landfill discharge of the concentrate is commonly
39 performed. However, some national regulations do not allow such practice and it clearly does
40 not fix the long-term issue. In recent years, several processes have been investigated and
41 applied at industrial scale for the treatment of NF concentrate. For example, adsorption
42 processes are effective to remove organic matters in NF concentrate, but they are strongly
43 limited by the high organic charge of such effluents compared to the adsorption capacity of
44 adsorbent materials. Besides, membrane distillation and evaporation processes are
45 substantially limited by the high cost of equipment and energy consumption (Cui et al., 2018).

46 During the last two decades, electrochemical advanced oxidation processes (EAOPs) have
47 received great attention for efficient degradation of a large range of hazardous and
48 biorefractory organic compounds. They are based on *in situ* electrogeneration of hydroxyl
49 radicals ($\bullet\text{OH}$), a non-selective and powerful oxidizing agent ($E^\circ(\bullet\text{OH}/\text{H}_2\text{O}) = 2.80 \text{ V vs SHE}$)
50 (Brillas et al., 2009; Comninellis et al., 2008; Martínez-Huitle et al., 2015; Panizza and
51 Cerisola, 2009). EAOPs provide also several technical advantages such as high versatility,

52 easy operation, possibility for automation and low consumption of chemical reagents
53 (Radjenovic and Sedlak, 2015; Sirés et al., 2014). However, complete mineralization of
54 organic compounds requires high energy consumption. Therefore, the combination of EAOPs
55 with biological processes is currently more and more investigated as a feasible option for
56 improving the global cost-effectiveness of the process (Oller et al., 2011; Ganzenko et al.,
57 2014, 2018; Trellu et al., 2016a; Olvera-Vargas et al., 2015). In order to achieve reliable
58 conclusions, the accurate assessment of biodegradability enhancement by EAOPs requires the
59 use of proper measurement tools such as respirometric methods (Reuschenbach et al., 2003).

60 Anodic oxidation (AO) and electro-Fenton (EF) are the most popularized EAOPs (Brillas et
61 al., 2009; Panizza and Cerisola, 2009; Martínez-Huitle et al., 2015; Oturan et al., 2015). AO is
62 based on the generation of hydroxyl radicals ($\bullet\text{OH}$) via water oxidation at the surface of
63 anodes (M) with high overvoltage for oxygen evolution reaction (Eq. 1) (Panizza and
64 Cerisola, 2009; Trellu et al., 2017; Özcan et al., 2008), while $\bullet\text{OH}$ are generated
65 homogeneously in the bulk during the EF process through the Fenton's reaction (Eq. 2)
66 (Brillas et al., 2009; Ma et al., 2016; Zhang et al., 2007). H_2O_2 and iron (II) are continuously
67 electrogenerated at the cathode by reduction of dissolved oxygen (Eq. 3) and iron (III)
68 reduction (Eq. 4), respectively. External oxygen supply is required and an iron source must be
69 either initially added at catalytic amount to the treated solution (homogeneous EF) or
70 embedded onto suitable electrode materials (heterogeneous EF).



75 Homogeneous EF requires external iron source and acidic pH (*i.e.* pH 2.5 – 3.5) that prevents
76 iron precipitation (Brillas et al., 2009). Besides, heterogeneous EF using for example pyrite
77 (Ammar et al., 2015) or iron loaded sepiolite (Iglesias et al., 2013) as iron source has been
78 developed in order to operate the process over a wide pH range (Ganiyu et al., 2018; Poza-
79 Nogueiras et al., 2018). Innovative electrodes have been also synthesized and studied as both
80 heterogeneous catalyst source and cathode materials (Zhang et al., 2012; Wang et al., 2013;
81 García-Rodríguez et al., 2016; Ganiyu et al., 2017a). Particularly, the modification of raw
82 carbon felt (CF) with CoFe or Fe^{II}Fe^{III}-layered double hydroxide (LDH) also led to an
83 increase of the electroactive surface area, which in turn improved the generation of H₂O₂ and
84 the global efficiency of the heterogeneous EF process (Ganiyu et al., 2017a, 2018). However,
85 to the best of our knowledge, none study focused on the application of such promising
86 electrodes for the treatment of real wastewaters.

87 As regards to anode materials, sub-stoichiometric titanium oxides (especially Ti₄O₇) recently
88 received great attention for application in wastewater treatment by AO (Guo et al., 2016;
89 Ganiyu et al., 2016; Trelu et al., 2018b). Ti₄O₇ anode is able to generate large amounts of
90 physisorbed hydroxyl radicals (Ti₄O₇(•OH)) for the degradation and mineralization of organic
91 contaminants. Besides, this material has the potential to become a low-cost anode compared
92 to the well known boron-doped diamond anode (Ganiyu et al., 2017b; Trelu et al., 2018b).
93 However, to the best of our knowledge, such anode material has also still not been applied for
94 the treatment of real effluents.

95 The objective of this study was to investigate the application of these novel electrode
96 materials for the treatment of a NF concentrate of landfill leachate initially pre-treated in a
97 MBR, which represents an important challenge for environmental engineering. Various
98 configurations (*i.e.* homogeneous EF, heterogeneous EF, heterogeneous EF/AO) and
99 operating conditions were studied. The efficiency for mineralization of organic compounds

100 was compared. A particular attention was also given to the understanding of mineralization
101 mechanisms by using various analytical tools: (i) DOM was characterized by three-
102 dimensional excitation and emission matrix fluorescence (3DEEM), (ii) degradation by-
103 products such as short-chain carboxylic acids were identified and quantified by ion-exclusion
104 HPLC, and (iii) inorganic ions released during the mineralization process were identified and
105 quantified by ion chromatography. Moreover, acute toxicity of the effluent was assessed by
106 Microtox[®] analysis and the possibility to use such EAOPs as a pre-treatment before
107 recirculation towards the MBR was assessed by using respirometric method under similar
108 conditions than in the real industrial MBR. Finally, recommendations were given by taking
109 into consideration the interdependence of MBR process and EAOP in a combined treatment
110 strategy.

111

112 **2. Materials and methods**

113 **2.1 Chemicals**

114 For the preparation of Fe^{II}Fe^{III}-LDH modified CF, iron III nitrate nonahydrate Fe(NO₃)₃.9H₂O
115 (CAS 7782-61-8, 98% purity), iron II sulfate heptahydrate FeSO₄.7H₂O (CAS 7782-63-0,
116 >99% purity), urea CO(NH₂)₂ (CAS 57-13-6) and ammonium fluoride NH₄F (CAS 12125-01-
117 8, 99% purity) were supplied by Sigma Aldrich. Ultra-pure water (Millipore Mill-Q system,
118 resistivity >18 MΩ.cm at 25 °C) was used for the preparation of all solutions.

119

120 **2.2 Landfill leachate (NF concentrate)**

121 The NF concentrate was collected from a landfill leachate wastewater treatment plant
122 (WWTP) in the south of France. The raw landfill leachate was initially treated in a MBR, then

123 followed by a NF step. The concentrate from the NF step was collected and stored in a
124 refrigerator at 4 °C.

125

126 **2.3 Electrochemical setup and electrode materials**

127 Experiments were conducted in an undivided cylindrical glass containing 220 mL of NF
128 concentrate at room temperature (25 °C). The electrochemical cell was similar to the one used
129 in several previous studies (Ganiyu et al., 2016; Trelu et al., 2016b). Either raw CF (for
130 homogeneous EF) or Fe^{II}Fe^{III}-LDH modified CF (for heterogeneous EF) was employed as
131 cathode (20 x 6 cm; 120 cm²), positioned on the inner wall of the cylindrical cell. CF (99.0%,
132 6.35 mm thick) was provided by Alfa Aesar. Fe^{II}Fe^{III}-LDH modified CF was prepared by *in-*
133 *situ* solvothermal process as reported elsewhere (Ganiyu et al., 2018). LDH coating was 0.62
134 ± 0.04 mg cm⁻², which was also in agreement with what has been reported previously (Ganiyu
135 et al., 2018). For homogeneous EF experiments, 0.2 mM of Fe²⁺ was added to the solution
136 and initial pH was adjusted at 3 (values usually reported as optimal). For heterogeneous EF
137 experiments, none Fe²⁺ was externally added and pH was not initially adjusted.

138 The anode was either a 24 cm² (3 x 8 cm) platinum mesh (for homogeneous and
139 heterogeneous EF) or a 32 cm² (4 x 8 cm) Ti₄O₇ thin film plasma deposited on Ti substrate
140 (for heterogeneous EF/AO) from Saint-Gobain Research Provence, France. Ti₄O₇ powder
141 used for plasma deposition was prepared by carbothermal reduction of TiO₂ as already
142 reported by our group (Ganiyu et al., 2016, 2017b). These rectuganlar-shaped anodes were
143 placed at the center of the cylindrical electrochemical cell, with an average interelectrode
144 distance of 3 cm.

145 Electrodes were connected to a DC power supply (CNB Electronique) with applied current set
146 at 1000 mA ($j = 8.3 \text{ mA cm}^{-2}$, calculated from the cathode surface, which is the working

147 electrode during the EF process) or 500 mA ($j = 4.2 \text{ mA cm}^{-2}$). A magnetic stirrer was used to
148 improve mass transport of chemical species toward/from the electrodes. The solution was
149 saturated with O_2 by bubbling compressed air through a glass frit 10 min before starting the
150 experiments and all along the electrolysis. The conductivity of the NF concentrate was high
151 enough to ensure the electrolysis without any additional supporting electrolyte. For
152 comparison, a reference experiment was also carried out using raw carbon felt cathode, Pt
153 anode, without adding any source of iron and without initial pH adjustment.

154

155 **2.4 Dissolved organic carbon and chemical oxygen demand analysis**

156 Mineralization rate of NF concentrate was determined by total organic carbon (TOC) analyses
157 using the Shimadzu TOC-L analyzer based on the 680 °C combustion catalytic oxidation
158 method. All samples were filtrated through 0.45 μm regenerated cellulose (RC) membrane
159 filters. Therefore, results are reported as dissolved organic carbon (DOC).

160 Chemical oxygen demand (COD) was analyzed by the method AFNOR NFT 90-101 using
161 Hach COD kits.

162

163 **2.5 Respirometric method for the determination of biodegradability**

164 Biodegradability was assessed with a BM-T Advance Respirometer (SURCIS S.L, Spain),
165 which consists in a 1 L capacity vessel, provided with an oxygen probe (Hamilton) and
166 temperature control system. The activated sludge used in the bioassays was collected from the
167 aerated tank of the landfill leachate WWTP using MBR and was thus acclimatized to the
168 effluent. The samples were evaluated without pH adjustment because of the high buffer
169 capacity of the landfill leachate samples. Continuous aeration and agitation were applied to

170 ensure air saturation conditions. Temperature was maintained at 20 °C during the tests and the
171 standardization with sodium acetate method was applied (conversion rate adjustment
172 according to instructions given by the respirometer's manufacturer). In order to inhibit the
173 nitrification process and measure the sample effect only on the heterotrophic bacteria, 1.5 mg
174 gVSS⁻¹ of N-allylthiourea was added before the beginning of each trial.

175 For biodegradability assays, 700 mL of endogenous activated sludge and 300 mL of target
176 sample were introduced in the respirometer. The biodegradability of samples pre-treated by
177 EAOPs was assessed through R tests (Fig. SI 1). The ratio between biodegradable COD
178 (bCOD) and DOC (bCOD/DOC) was used for determination of the biodegradable character
179 of each sample. A biodegradability enhancement index (BE) was calculated by using Eq. 5.

$$180 \quad BE = \frac{\left(\frac{bCOD_t}{DOC_t}\right)}{\left(\frac{bCOD_0}{DOC_0}\right)} \quad (\text{Eq. 5})$$

181 Where bCOD_t and DOC_t are bCOD (in gO₂ L⁻¹) and DOC (in gC L⁻¹) after t hours of
182 treatment by EAOPs.

183

184 **2.6 Characterization of dissolved organic matter by three-dimensional excitation and** 185 **emission matrix fluorescence (3DEEM)**

186 Natural organic matter was characterized by 3DEEM fluorescence using a Perkin-Elmer LS-
187 55 spectrometer (USA). The dilution factor was 200 for all samples. The procedure reported
188 by Jacquin et al. (2017) was used for fluorescence spectra acquisition and data extraction.
189 Chen et al. (2003) divided fluorescence spectra into five different areas corresponding to
190 different groups of fluorophores, *i.e.* regions I and II for aromatic proteins, region III for
191 fulvic acid-like (FA-like) fluorophores, region IV for soluble microbial by-product-like
192 (SMP-like) fluorophores and region V for humic acid-like (HA-like) fluorophores (Chen et

193 al., 2003). Similarly to the study of Jacquin et al. (2017), 3DEEM results were analyzed by
194 taking into consideration only 3 different zones, *i.e.* zone I' for region I + II (aromatic
195 proteins), zone II' for region IV (SMP-like) and zone III' for region III + V (HA + FA-like).
196 Calculation of the volume of fluorescence in these different zones was achieved following the
197 method from Jacquin et al. (2017).

198

199 The procedures for ICP-MS analysis, toxicity test (Microtox[®]), identification/quantification of
200 inorganic ions and identification/quantification of short-chain carboxylic acids are provided in
201 Supplementary Information (SI)

202

203 **3. Results and Discussion**

204 **3.1 Characterization of the NF concentrate effluent**

205 The characteristics of the NF concentrate are presented in SI (Table SI 1). The effluent was a
206 dark brown liquid having a slightly alkaline pH and high organic charge. The conductivity
207 (3.1 mS cm^{-1}) was between 3 and 10 times lower than values reported in the literature ($10 -$
208 33 mS.cm^{-1}) (Li et al., 2015; Hu et al., 2018; Xu et al., 2017). This might be ascribed to the
209 high diversity of landfill leachate and to a lower retention of ionic species by the NF step. In
210 fact, the concentration factor of divalent (Mg^{2+} and Ca^{2+}) and monovalent (Na^+ and K^+)
211 inorganic cations was only 3.1 ± 0.3 and 1.7 ± 0.2 , respectively. Monovalent ions are less
212 retained by the NF membrane because of lower charge interactions (Van der Bruggen et al.,
213 2004). The COD content ($2.1 \text{ gO}_2 \text{ L}^{-1}$) of the NF concentrate was also in the low range of
214 values usually reported in the literature ($1.7 - 5.5 \text{ gO}_2 \text{ L}^{-1}$) (Li et al., 2015; Hu et al., 2018; Xu

215 et al., 2017) because of (i) operation of the NF process with a lower concentration factor
216 and/or (ii) lower initial organic loading rate of the effluent pre-treated by the MBR.

217 The ratio between bCOD and total COD or total DOC was low (bCOD/COD = 0.12 or
218 bCOD/DOC = 0.26), thus indicating the low biodegradability of the concentrate owing to the
219 presence of high-molecular weight and non-biodegradable compounds. A high concentration
220 of nitrate ($[\text{NO}_3^-] = 90 \text{ mg L}^{-1}$) was observed due to the complete nitrification of NH_4^+ in the
221 MBR and a partial denitrification in the anoxic tank. As regards to metals, Sr ion was the
222 most concentrated (6.0 mg L^{-1}) and high concentrations of As (0.38 mg L^{-1}) and Cr (0.67 mg
223 L^{-1}) ions were also reported. These metals are released by different wastes discarded in
224 landfills such as glass products (Ponthieu et al., 2007), fluorescent lights and ceramics
225 (Mahindrakar and Rathod, 2018). Besides, the high concentration of Sb (1.2 mg L^{-1}) indicates
226 an important contamination from plastic decomposition (Westerhoff et al., 2008).

227

228 **3.2 Mineralization efficiency by EAOPs**

229 Various experiments were performed in order to evaluate the efficiency of different EAOPs
230 (*i.e.* homogeneous EF, heterogeneous EF, heterogeneous EF/AO).

231 First, volatilization and adsorption phenomena were evidenced by bubbling oxygen into the
232 electrochemical reactor without any current applied. Under these conditions, DOC content
233 decreased and reached a plateau after 1 h with around $16 \pm 2\%$ removal due to volatilization of
234 volatile organic compounds and adsorption of hydrophobic compounds on CF. Besides, it was
235 also observed that initial pH adjustment at 3 decreased the DOC value by $15 \pm 5\%$ because of
236 the precipitation of humic acids.

237 Second, experiments were carried out under constant current density applied to the electrodes.

238 As shown in Fig. 1A, 96% of DOC removal was obtained after 8 h of electrolysis by

239 heterogeneous EF with an applied current density of 8.3 mA cm^{-2} . The excellent
240 mineralization efficiency achieved without any initial pH adjustment was ascribed to
241 heterogeneous EF reaction occurring on the surface of the $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH catalyst, which
242 catalyzed the decomposition of H_2O_2 to generate large amounts of hydroxyl radicals $\bullet\text{OH}$
243 (Ganiyu et al., 2018). By comparison, only 59% removal of DOC was obtained in the
244 reference experiment with analogous operating conditions but using raw CF as cathode
245 instead of $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH modified CF, thus highlighting the benefits obtained from the use of
246 the modified cathode. Heterogeneous EF was also even more efficient than conventional
247 homogeneous EF (90% removal of DOC after 8 h) with external addition of iron catalyst (0.2
248 mmol L^{-1}) and initial adjustment of pH at 3. Therefore, $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH modified CF appeared
249 as an ideal cathode material that enhances the mineralization efficiency of the EF process and
250 avoids both initial pH adjustment and external addition of iron. The stability and reusability of
251 the prepared $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH modified CF cathode was assessed by reusing the same electrode
252 for 4 successive electrolysis cycles (Fig. 2). The slight but continuous decrease of the
253 efficiency could be ascribed to (i) the initial loss of loosely bounded $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH at the
254 external surface of the CF substrate due to vigorous stirring and/or (ii) iron leaching at acidic
255 pH for which $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH coating is less stable. Indeed, a progressive reduction of pH to
256 values in the range 3 - 4 was noticed during electrolysis, which was explained by the
257 generation of short-chain carboxylic acids. This leaching phenomenon was confirmed by
258 comparison of SEM images of raw CF, $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH modified CF before use and $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -
259 LDH modified CF after 4 cycles of electrolysis (Fig. SI 2). While an extensive growth on the
260 CF of dense platelets of $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH with uneven and porous structure was initially
261 obtained, a partial degradation of the global $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH structure was observed after 4
262 cycles. Thus, in order to avoid the depletion of the catalyst, a continuous pH control

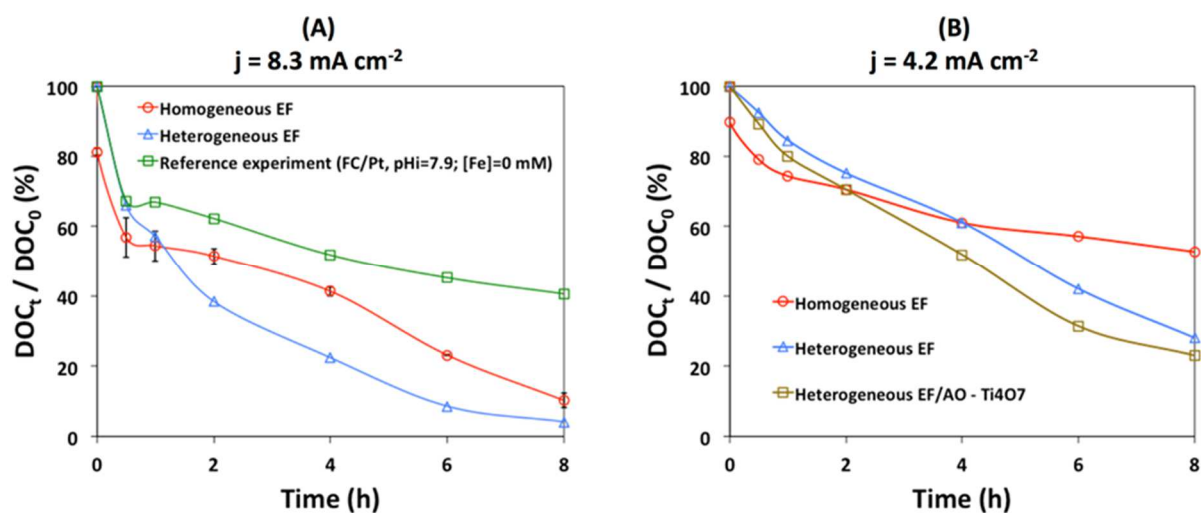
263 regulation would be recommended rather than a unique pH re-adjustment at the end of the
264 treatment.

265 Third, experiments were also performed at lower current density (4.2 mA cm^{-2} , Fig. 1B).
266 Lower DOC removal rates were obtained, e.g. 72% vs 96% removal of DOC after 8 h of
267 treatment by heterogeneous EF at 4.2 mA cm^{-2} and 8.3 mA cm^{-2} , respectively. Indeed, high
268 current density enhances the generation of hydrogen peroxide and regeneration rate of Fe(II).
269 Thus, further $\bullet\text{OH}$ are produced for oxidation and mineralization of organic compounds.

270 Fourth, the influence of the anode material was also investigated (Fig. 1B). After 8h of
271 treatment by heterogeneous EF/AO using Ti_4O_7 anode at 4.2 mA cm^{-2} , 77% removal of DOC
272 was achieved, compared to 72% by heterogeneous EF using Pt anode. The higher efficiency
273 of Ti_4O_7 anode was ascribed to the generation at the anode surface of physisorbed $\text{Ti}_4\text{O}_7(\bullet\text{OH})$
274 with great oxidation ability, because of the higher overvoltage for oxygen evolution reaction
275 ($>0.7 \text{ V}$), compared to Pt anode ($<0.4 \text{ V}$) (Trellu et al., 2018a; Ganiyu et al., 2018). Therefore,
276 by taking also into consideration the lower cost of Ti_4O_7 , this anode material appeared as a
277 suitable electrode to combine AO and EF processes.

278

279



280 **Figure 1** – DOC removal efficiency vs time during the mineralization of NF concentrate.

281 Comparison of different configurations and operating conditions.

282 **(A) $j = 8.3 \text{ mA cm}^{-2}$** : (○) Homogeneous EF with CF cathode, Pt anode, $[\text{Fe}^{2+}] = 0.2 \text{ mmol L}^{-1}$, $\text{pH}_0 = 3$; (△) Heterogeneous EF with $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH modified CF cathode, Pt anode, $[\text{Fe}^{2+}] =$
 283 0 mmol L^{-1} , $\text{pH}_0 = 7.9$; (□) Reference experiment with CF cathode, Pt anode, $[\text{Fe}^{2+}] = 0 \text{ mmol}$
 284 L^{-1} , $\text{pH}_0 = 7.9$. Three homogeneous EF experiments were performed in order to assess the
 285 reproducibility of the experimental procedure. Standard deviations are reported in Figure 1A.
 286

287 **(B) $j = 4.2 \text{ mA cm}^{-2}$** : (○) Homogeneous EF with CF cathode, Pt anode, $[\text{Fe}^{2+}] = 0.2 \text{ mmol L}^{-1}$, $\text{pH}_0 = 3$; (△) Heterogeneous EF with $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH modified CF cathode, Pt anode, $[\text{Fe}^{2+}] =$
 288 0 mmol L^{-1} , $\text{pH}_0 = 7.9$; (□) Heterogeneous EF/AO with $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH modified CF cathode,
 289 Ti_4O_7 anode, $[\text{Fe}^{2+}] = 0 \text{ mmol L}^{-1}$, $\text{pH}_0 = 7.9$.
 290

291

292 One of the main challenges for EAOPs is to reduce the energy consumption (EC). The EC
 293 was expressed as kWh per g of DOC removed and calculated from Eq. 6 (Brillas et al., 2009).

294

295

296 EC_{DOC} (kWh g^{-1} of DOC) = $\frac{E_{cell}It}{V\Delta(DOC)_{exp}}$ (Eq. 6)

297 where E_{cell} is the average cell voltage (V), I the applied current (A), t the duration of
298 electrolysis (h), V the volume of solution treated (L) and $\Delta(DOC)_{exp}$ the experimental decays
299 of DOC ($mgC L^{-1}$).

300 For example, 96% removal of DOC by heterogeneous EF at $8.3 mA cm^{-2}$ required 0.35 kWh
301 g^{-1} of DOC removed. By comparison, 77% DOC removal by heterogeneous EF/AO using
302 Ti_4O_7 anode at $4.2 mA cm^{-2}$ required only 0.13 kWh g^{-1} of DOC. This value decreased to 0.11
303 kWh g^{-1} of DOC for 45% removal of DOC by the same process (4 h of treatment instead of 8
304 h). In fact, high current density strongly increased energy consumption because of
305 concomitant rise in total cell voltage (from 7.2 to 10.4 V). Parasitic reaction such as hydrogen
306 evolution, $4 e^{-}$ reduction of O_2 to H_2O and $2 e^{-}$ oxidation of water to O_2 are also promoted at
307 higher current density. Moreover, at high removal rate of DOC and low residual DOC, the
308 current efficiency is strongly decreased by mass transport limitations. Therefore, in order to
309 achieve high DOC removal rate with low energy consumption, it was proposed to combine
310 EAOPs with a biological post-treatment, since EAOPs are able to transform biorefractory
311 organic compounds into by-products that are more biodegradable than initial compounds
312 (Ganzenko et al., 2014; Olvera-Vargas et al., 2015; Trellu et al., 2016a). In practice, that
313 would mean implementing the EAOP as a preliminary treatment before recirculation of the
314 NF concentrate towards the MBR. Thus, the next objective of this study was to monitor the
315 evolution of the biodegradability of the effluent, as well as to better understand the evolution
316 of organic and inorganic compounds during EAOPs.

317

318

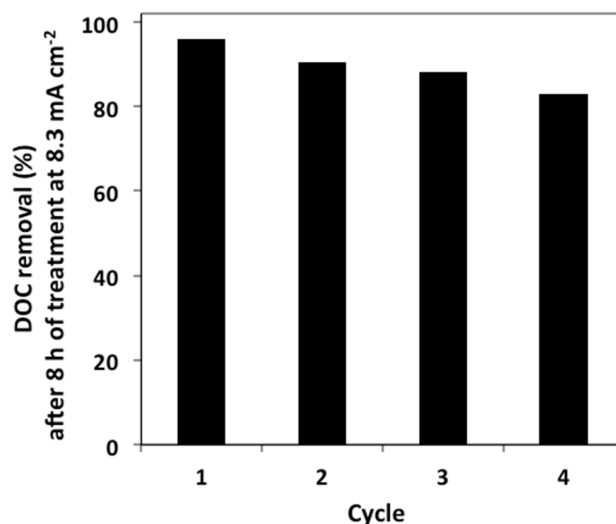


Figure 2 – DOC removal after 8h of electrolysis vs number of cycles for the heterogeneous EF treatment using Fe^{II}Fe^{III}-LDH/CF cathode and Pt anode at 8.3 mA cm⁻².

319

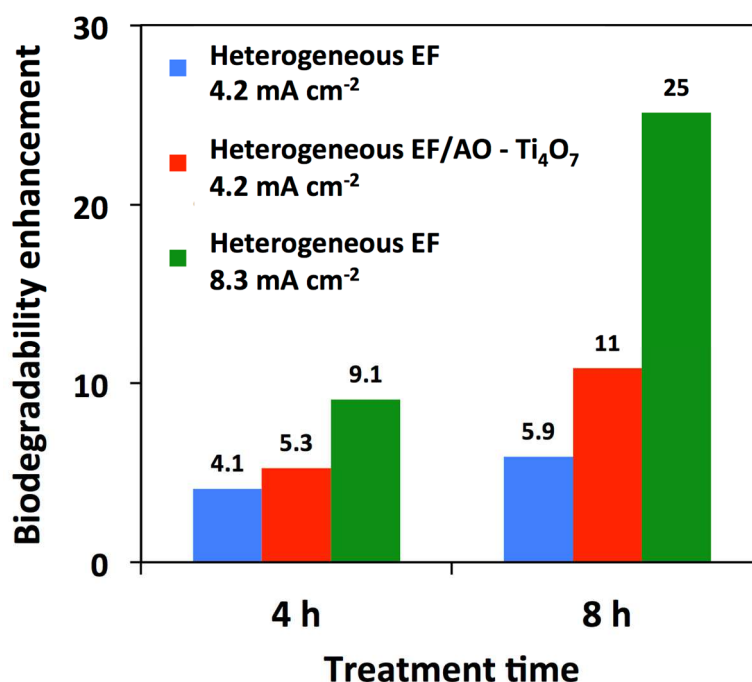
320

321 **3.3 Biodegradability enhancement**

322 Respirometric measurements performed have the advantage of being a direct and rapid
 323 biological assessment of aerobic degradation under similar conditions than in the real
 324 industrial MBR (Reuschenbach et al., 2003). The calculated biodegradability enhancement
 325 (BE) values are presented in Fig. 3 after 4 and 8h of treatment for different EAOP
 326 configurations and current densities applied.

327 In all cases, heterogeneous EF shows an increase in the biodegradability of the NF
 328 concentrate, which validate the potential positive role of EAOPs as a pre-treatment before
 329 recirculating the NF concentrate to the MBR. In general the more the mineralization rate is
 330 achieved the more the biodegradability enhancement is observed. After 8 h of treatment by
 331 heterogeneous EF at 4.2 mA cm⁻², the BE index reached 11 with Ti₄O₇ anode, compared to
 332 5.9 with Pt anode, which validate the positive effect of combining EF with AO. We could

333 state that a satisfactory biodegradability enhancement (between 4.1 and 9.1) was reached after
334 4 h with optimal energy consumption efficiency.



335 **Figure 3** – Biodegradability enhancement at two different treatment times for different
336 EAOPs and current densities.

337

338

339 **3.4 Acute toxicity of the effluent**

340 Toxicity evolution of the NF concentrate during its electrolysis by both heterogeneous and
341 homogeneous EF was evaluated by Microtox[®] standard method (Fig. 4). Initially (i.e. before
342 any treatment of the NF concentrate), around 50±5% of *V. fischeri* luminescence inhibition
343 was observed because of the presence of large amount of toxic trace metals and organic
344 pollutants. *V. fischeri* luminescence inhibition can be sensitive to various phenomena that can
345 not be controlled and studied separately during the treatment of such complex effluent, e.g.
346 the removal of toxic organic pollutants, the formation of toxic degradation by-products, the

347 formation of non-toxic degradation by-products (which could promote stimulation of bacterial
 348 luminescence), the release of inorganic compounds and the evolution of metal speciation.
 349 Therefore, it is difficult to draw reliable conclusions from these analyses. However, results
 350 show that EAOPs are not able to remove completely the acute toxicity from such complex
 351 effluent containing both organic and inorganic toxic compounds. Luminescence inhibition
 352 might even increase significantly, particularly after achieving high mineralization rates (for
 353 example during heterogeneous EF at 8.3 mA cm^{-2} and heterogeneous EF at 4.2 mA cm^{-2}).
 354 From the comparison of heterogeneous EF experiments performed with either Pt or Ti_4O_7
 355 anode, it seems that anodic oxidation participate to avoid the accumulation of toxic by-
 356 products in the bulk.

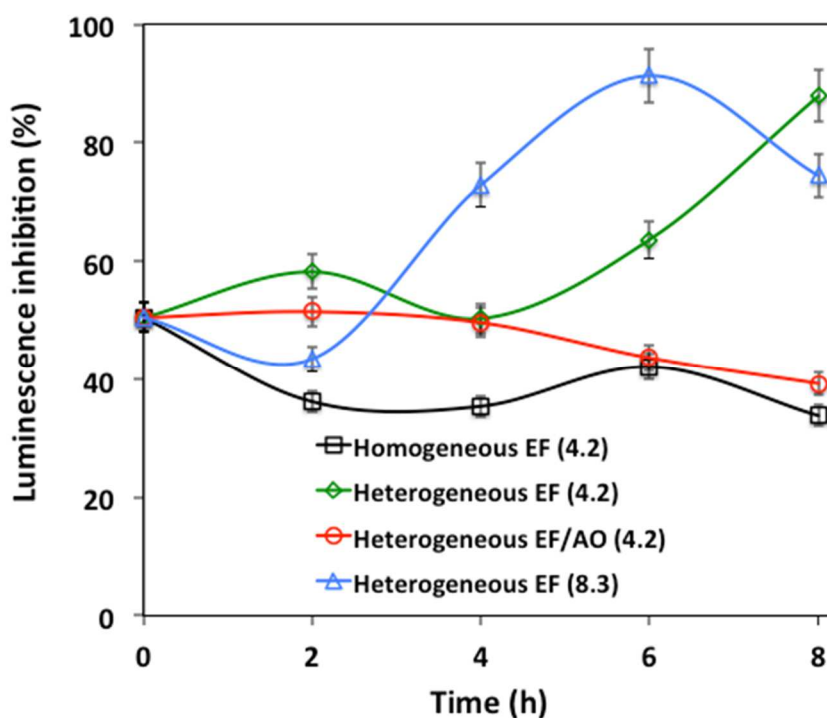


Figure 4 – Evolution of *Vibrio fischeri* luminescence inhibition (Microtox[®] test) vs electrolysis time according to EAOP configuration and current density (in brackets, mA cm^{-2})

357

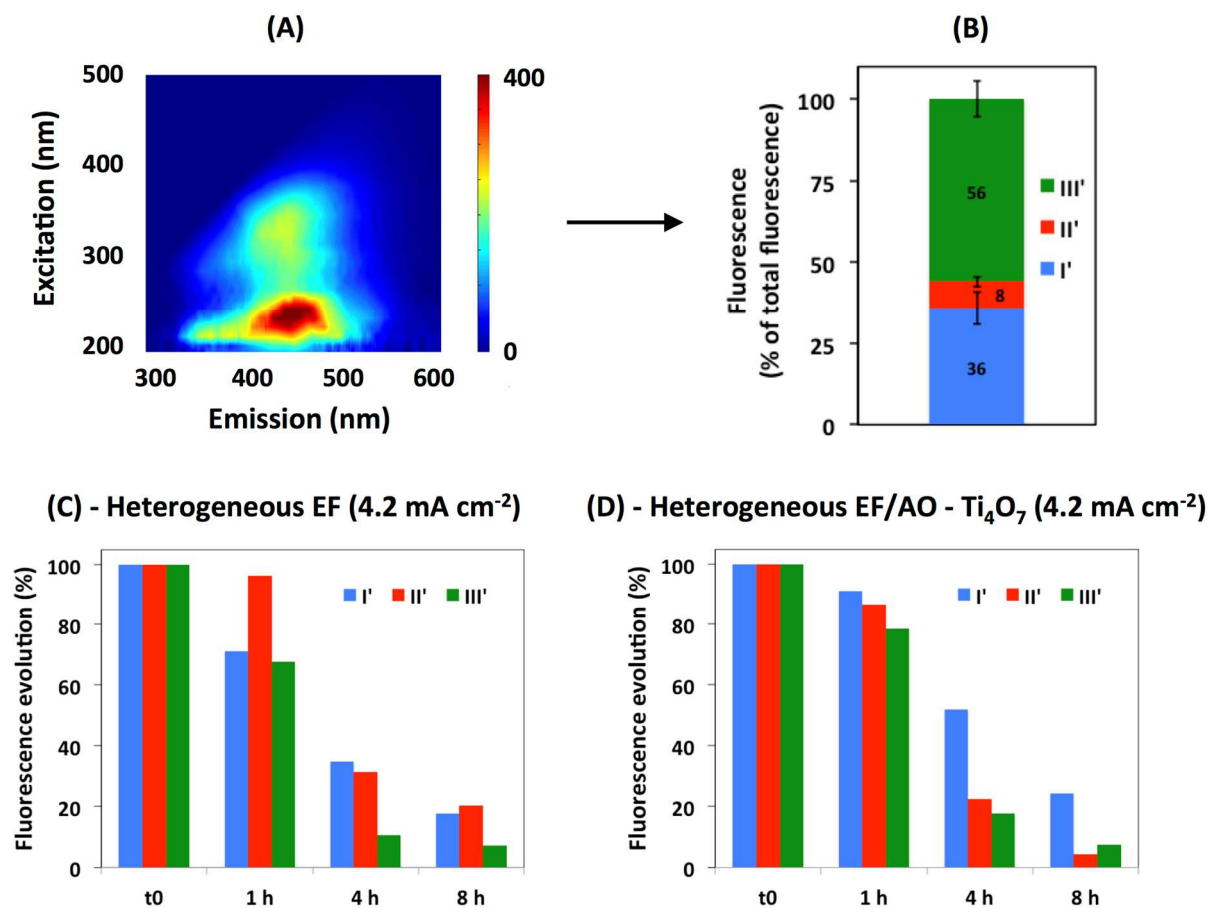
358

3.5 Characterization of the organic matter by 3DEEM

3DEEM has been reported to be a useful tool for characterization of colloidal and dissolved organic matter. Particularly, Jacquin et al. (2017) recently emphasized a correlation between the volume of fluorescence of zone III' (HA+FA-like fluorophores) and the concentration in a full-scale MBR of humic substances (MW \approx 1000 Da) and building blocks (degradation by-products from humic substances, with MW \approx 300-500 Da) measured by size exclusion liquid chromatography coupled with organic carbon and organic nitrogen detector. Similarly, a correlation was also obtained between the volume of fluorescence of zone II' (SMP-like fluorophores) and the concentration of proteins from biopolymers (MW \approx 20,000 - 7.5×10^{11} Da). Besides, no correlation was obtained for the volume of fluorescence of zone I' because fluorophores of this zone would be mostly associated with colloidal proteins that could not be analyzed by size exclusion liquid chromatography (Jacquin et al., 2017). In this study, we proposed to use these results with the view to obtain indications on the evolution of the nature of the organic matter during EAOPs.

The initial effluent was mainly constituted of HA+FA-like fluorophores from zone III' as shown in Fig. 5A and 5B. This result is consistent with the pre-treatment of the effluent in the MBR because microfiltration membranes have lower retention capacity for these low MW compounds. A strong decrease of fluorophores of zone III' was then observed during both heterogeneous EF (Fig. 5C) and EF/AO (Fig. 5D) due to (i) precipitation of humic acids at acidic pH and (ii) fast degradation and mineralization of these low MW compounds with an aromatic structure that reacts quickly with $\cdot\text{OH}$ (Trellu et al., 2016b). By comparison, fluorescence from colloidal proteins of zone I' decreased much more slowly. This phenomenon might be ascribed to the lower availability of colloids for reaction with $\cdot\text{OH}$ in the aqueous phase. In the context of the global combined process, this result means that improving the retention of colloidal proteins in the MBR would then have a positive effect on

384 the efficiency of the EAOP. Finally, it was also observed that the use of Ti_4O_7 anode for the
 385 heterogeneous AO/EF process promoted significantly the decrease of the fluorescence of
 386 proteins from biopolymers (zone II', Fig. 5D). This might be ascribed to the higher electro-
 387 catalytic activity of Ti_4O_7 for anodic oxidation of organic compounds, compared to Pt anode
 388 (Ganiyu et al., 2016).



389
 390 **Figure 5** – Fluorescence evolution of the NF concentrate effluent treated by EAOPs, as a
 391 function of time treatment. Zone I': colloidal proteins ; zone II': Soluble Microbial by-
 392 Product (SMP)-like fluorophores ; zone III': Humic and Fulvic acids (HA + FA)-like
 393 fluorophores.

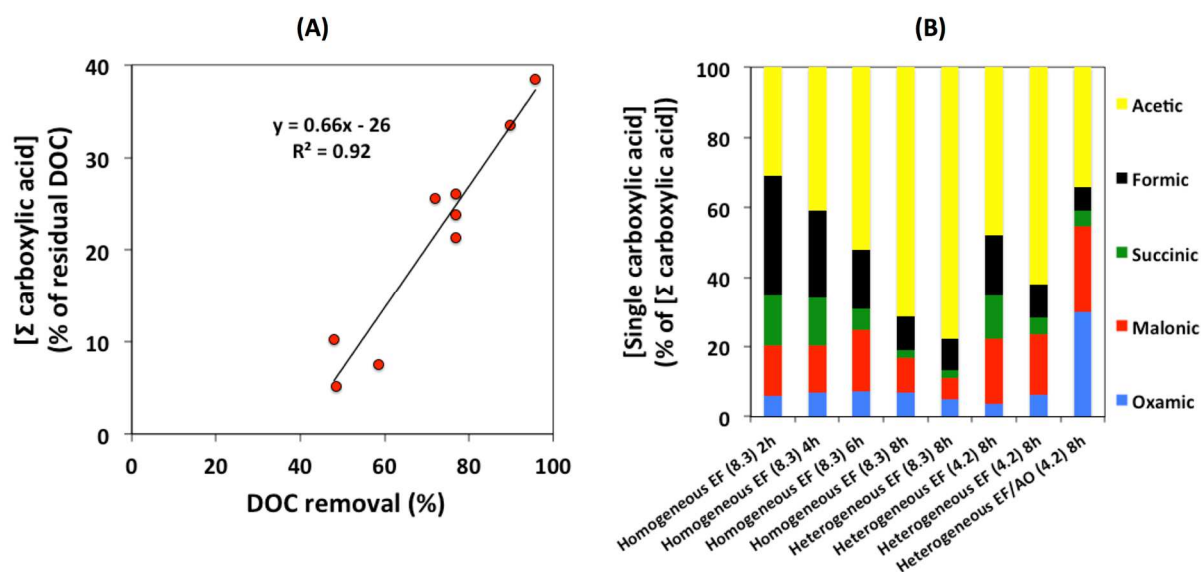
394

395

3.6 Identification and quantification of short-chain carboxylic acids

Short-chain carboxylic acids are common degradation by-products generated during the degradation of organic compounds by EAOPs (Oturán et al., 2008). In this study, they were identified by ion-exclusion HPLC and chromatograms revealed five well defined peaks corresponding to acetic, formic, succinic, malonic and oxamic acids. Results were firstly analyzed by following the evolution of the concentration of the sum of carboxylic acids analyzed. The evolution of the concentration according to treatment time was determined for homogeneous EF at 8.3 mA cm^{-2} . An initial increase of the concentration was observed until $t = 6 \text{ h}$ ($[\Sigma \text{ carboxylic acids}] = 55 \text{ mgC L}^{-1}$) because of the degradation of aromatic pollutants. Then, the concentration decreased ($[\Sigma \text{ carboxylic acids}] = 34 \text{ mgC L}^{-1}$ at $t = 8 \text{ h}$). In fact, the lower concentration of organic compounds after 6 hours of treatment resulted in lower formation rate of carboxylic acids compared to the degradation rate. Interestingly, Fig. 6A shows a linear correlation ($R^2 = 0.92$) between DOC removal (%) and proportion of carboxylic acids among the residual DOC. This correlation took into consideration all analyzes performed for the different configurations tested. The higher the DOC removal achieved, the higher was the proportion of carboxylic acids among the residual DOC. This result is consistent with the lower reaction rate constant of short-chain carboxylic acids with $\bullet\text{OH}$ ($10^7 - 10^8 \text{ M}^{-1} \text{ s}^{-1}$) compared to aromatic compounds ($10^9 - 10^{10} \text{ M}^{-1} \text{ s}^{-1}$) from which they are formed (Oturán et al., 2008). The evolution of the proportion of each single carboxylic acid was also studied (Fig. 6B). It was noticed that the proportion of acetic acid among total carboxylic acid concentration was continuously increased over time during homogeneous EF at 8.3 mA cm^{-2} . This phenomenon is also consistent with the lower reaction rate constant of acetic acid with $\bullet\text{OH}$ ($1.6 \times 10^7 \text{ M}^{-1} \text{ s}^{-1}$) compared to other carboxylic acids (Oturán et al., 2008). Similar trend was observed after 8 h of electrolysis using heterogeneous EF at 8.3 mA cm^{-2} or 4.2 mA cm^{-2} and homogeneous EF at 4.2 mA cm^{-2} , i.e. the higher the DOC removal

421 rate (Fig. 1), the higher the proportion of acetic acid. Heterogeneous EF/AO using Ti_4O_7 was
 422 the only experiment exhibiting a different trend, thus indicating that Ti_4O_7 might modify
 423 mineralization mechanisms, compared to Pt anode. Higher proportion of N-containing oxamic
 424 acid and lower proportion of acetic acid were observed with Ti_4O_7 anode, which might be
 425 attributed to the better electro-catalytic ability of Ti_4O_7 anode for the degradation of organic
 426 nitrogen and acetic acid, compared to Pt anode. As regards to oxamic acid, these results were
 427 consistent with 3DEEM. In fact, Ti_4O_7 anode further degraded proteins from biopolymers
 428 (zone II'), which usually contain high concentration of N (ratio C/N around 3) (Jacquin et al.,
 429 2017). Overall, the increase of the proportion of carboxylic acids over time confirms the
 430 suitability of the combination of EAOPs with a biological treatment (Fig. 3), because of the
 431 well-known high biodegradability of such compounds.



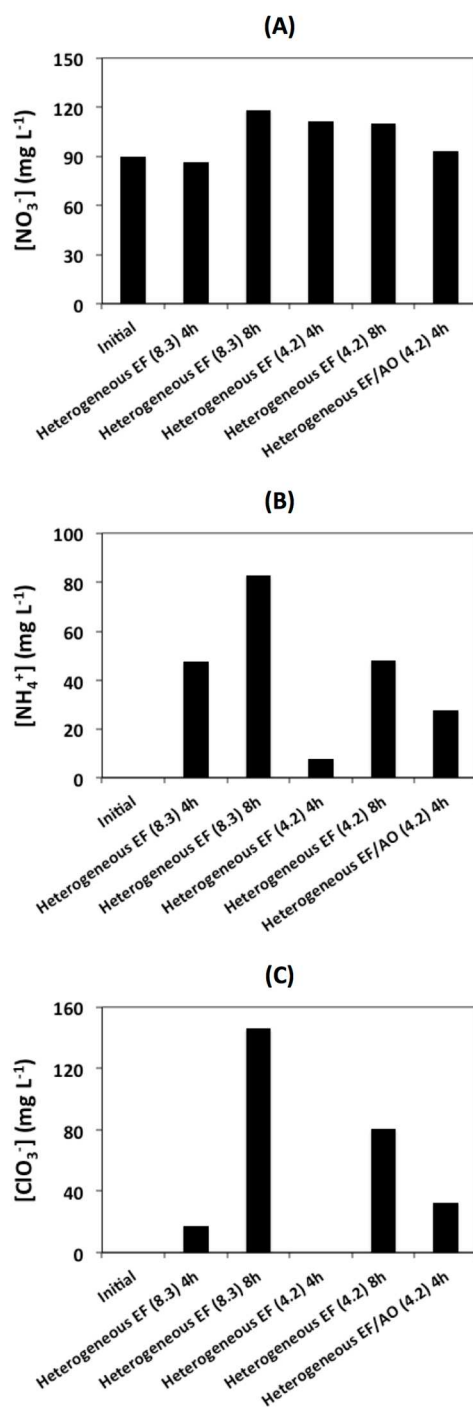
432
 433 **Figure 6** – Generation of short-chain carboxylic acids during EAOPs: (A) correlation
 434 between DOC removal (%) and proportion of carboxylic acids among the residual DOC;
 435 (B) evolution of the proportion of each single carboxylic acid among the total carboxylic acid
 436 concentration as a function of configurations and operating conditions (current density in
 437 bracket, $mA\ cm^{-2}$).

438 3.7 Evolution of inorganic species

439 Mineralization of Cl^- and N-containing organic compounds was accompanied by the
440 formation of inorganic ions. The main inorganic species of interest (NO_3^- , NH_4^+ , ClO_3^- and
441 ClO_4^-) were analyzed by ion chromatography. Results are presented in Fig. 7. In all
442 experiments, a stronger increase of the concentration of NH_4^+ was observed compared to NO_3^-
443 (Fig. 7A and 7B). These results can be explained by (i) the direct release of NH_4^+ from
444 mineralization of organic nitrogen and (ii) the release of NO_3^- followed by reduction of NO_3^-
445 into NH_4^+ at the cathode. Actually, the latter reaction is often observed during EF and AO,
446 while oxidation of NH_4^+ at the anode is hindered by the positive charge of this ion (Martin de
447 Vidales et al., 2016; Mousset et al., 2018). The amount of NH_4^+ in the solution increased with
448 treatment time and current density because of the higher mineralization rate of organic
449 nitrogen (Fig. 7B). Higher concentration of NH_4^+ was also obtained using Ti_4O_7 anode
450 (heterogeneous EF/AO), compared to Pt anode (heterogeneous EF). This phenomenon is
451 consistent with the results obtained from 3DEEM and with the higher concentration of
452 oxamic acid previously reported, and supports the higher electro-catalytic activity of Ti_4O_7 for
453 the degradation of organic nitrogen. In the context of a treatment strategy including a
454 recirculation of the effluent back to the MBR, the release of NH_4^+ during EAOPs would
455 require further biological nitrification in the MBR.

456 Oxidation of Cl^- resulted in a strong increase of ClO_3^- concentration (Fig. 7C). The reaction
457 mechanisms usually go through Cl^- oxidation into Cl_2 , followed by hydrolysis of Cl_2 into
458 hypochlorous acid HOCl . Further oxidation of HOCl lead to the formation of ClO_2^- (which
459 was not detected because of the fast oxidation kinetic) and subsequently ClO_3^- . No formation
460 of ClO_4^- was detected in any experiment because higher current density ($> 30 \text{ mA cm}^{-2}$) is
461 required to form this compound (Mousset et al., 2018). Besides, it was also observed that the
462 formation of ClO_3^- mainly occurred between 4 and 8 h of treatment, most probably because of

463 the preferential reaction of HOCl with organic species and NH_4^+ (break-point chlorination)
464 between 4 and 8 h of treatment (Martin de Vidales et al., 2016; Mousset et al., 2018). These
465 results confirm the suitability to stop the treatment after 4 h of treatment in order to avoid the
466 accumulation of ClO_3^- and NH_4^+ . While ClO_3^- can be toxic for the biomass in the MBR, the
467 formation of NH_3 at basic pH is also highly toxic for the autotrophic biomass (Jacquin et al.,
468 2018).



469 **Figure 7** – Concentration evolution of main inorganic species of interest during EAOPs as a
 470 function of configurations and operating conditions (current density in bracket, mA cm⁻²): (A)
 471 ClO₃⁻, (B) NO₃⁻ and (C) NH₄⁺. ClO₄⁻ was not detected.

472

473

474 **4. Conclusion**

475 NF concentrate of landfill leachate pre-treated in a MBR contains high concentration of
476 biorefractory organic pollutants that makes very difficult to treat or detoxify by conventional
477 techniques. The investigation of different configurations of EAOPs applied to this complex
478 effluent showed that heterogeneous EF/AO using Ti_4O_7 anode and $\text{Fe}^{\text{II}}\text{Fe}^{\text{III}}$ -LDH modified CF
479 cathode is the most efficient process for mineralization of organic pollutants. Without initial
480 pH adjustment and external addition of Fe^{2+} in the bulk, 45% removal of DOC was achieved
481 at 4.2 mA cm^{-2} with limited energy consumption (0.11 kWh g^{-1} of DOC removed, i.e. 49.5
482 kWh m^{-3}). Up to 96% removal of DOC was also obtained by using higher current density (8.3
483 mA cm^{-2}) during the heterogeneous EF process. The use of Ti_4O_7 anode appeared to be a key
484 parameter for improving the degradation and mineralization of organic nitrogen. From
485 3DEEM analysis, colloidal proteins were observed to be the most refractory organic
486 compounds. Therefore, improving the retention of such compounds in the MBR could
487 improve the efficiency of the EAOP.

488 The acute toxicity of the effluent was not removed but strong biodegradability enhancement
489 was observed after 4 h of treatment by heterogeneous EF/AO, thus making possible the
490 recirculation of the residual DOC towards the MBR in order to achieve total COD removal
491 without longer electrochemical treatment time. This result was consistent with the
492 identification and quantification of more biodegradable and less toxic by-products such as
493 carboxylic acids. In fact, a linear correlation was observed between DOC removal rate and
494 proportion of carboxylic acids in the residual DOC. As regards to the fate of inorganic
495 species, the formation of ClO_3^- could be limited by stopping electro-oxidation at 4 h, but
496 enhanced nitrification would be required in the MBR because of the release of NH_4^+ from
497 mineralization of organic nitrogen. Overall, these results emphasized the interdependence

498 between the MBR process and the EAOP in a combined treatment strategy and demonstrated
499 that the use of EAOPs using suitable electrode materials can be useful for the management of
500 such complex effluent.

501 **Acknowledgements**

502 We gratefully acknowledge the National French Agency of Research ‘ANR’ for funding the
503 project ECOTS/CELECTRON. Authors are also grateful to Saint Gobain Research Provence
504 for supplying Ti_4O_7 electrodes.

505

506

507

508

509

510

511

512

513

514

515 **References**

- 516 Amaral, M.C., Moravia, W.G., Lange, L.C., Zico, M.R., Magalhães, N.C., Ricci, B.C., Reis, B.G.,
517 2016. Pilot aerobic membrane bioreactor and nanofiltration for municipal landfill leachate
518 treatment. *Journal of Environmental Science and Health, Part A* 51, 640–649.
- 519 Ammar, S., Oturan, M.A., Labiadh, L., Guersalli, A., Abdelhedi, R., Oturan, N., Brillas, E., 2015.
520 Degradation of tyrosol by a novel electro-Fenton process using pyrite as heterogeneous
521 source of iron catalyst. *Water Research* 74, 77–87.
522 <https://doi.org/10.1016/j.watres.2015.02.006>
- 523 Brillas, E., Sirés, I., Oturan, M.A., 2009. Electro-Fenton Process and Related Electrochemical
524 Technologies Based on Fenton's Reaction Chemistry. *Chemical Reviews* 109, 6570–6631.
525 <https://doi.org/10.1021/cr900136g>
- 526 Campagna, M., Çakmakçı, M., Büşra Yaman, F., Özkaya, B., 2013. Molecular weight
527 distribution of a full-scale landfill leachate treatment by membrane bioreactor and
528 nanofiltration membrane. *Waste Management* 33, 866–870.
529 <https://doi.org/10.1016/j.wasman.2012.12.010>
- 530 Chen, W., Westerhoff, P., Leenheer, J.A., Booksh, K., 2003. Fluorescence Excitation–Emission
531 Matrix Regional Integration to Quantify Spectra for Dissolved Organic Matter. *Environmental*
532 *Science & Technology* 37, 5701–5710. <https://doi.org/10.1021/es034354c>
- 533 Comninellis, C., Kapalka, A., Malato, S., Parsons, S.A., Poulios, I., Mantzavinos, D., 2008.
534 Advanced oxidation processes for water treatment: Advances and trends for R&D. *J. Chem.*
535 *Technol. Biotechnol.* 83, 769–776. <https://doi.org/10.1002/jctb.1873>
- 536 Cui, Y.-H., Xue, W.-J., Yang, S.-Q., Tu, J.-L., Guo, X.-L., Liu, Z.-Q., 2018.
537 Electrochemical/peroxydisulfate/Fe³⁺ treatment of landfill leachate nanofiltration
538 concentrate after ultrafiltration. *Chemical Engineering Journal* 353, 208–217.
539 <https://doi.org/10.1016/j.cej.2018.07.101>
- 540 Ganiyu, S.O., Huong Le, T.X., Bechelany, M., Oturan, N., Papirio, S., Esposito, G., van
541 Hullebusch, E., Cretin, M., Oturan, M.A., 2018. Electrochemical mineralization of
542 sulfamethoxazole over wide pH range using FeII/FeIII LDH modified carbon felt cathode:
543 Degradation pathway, toxicity and reusability of the modified cathode. *Chemical Engineering*
544 *Journal* 350, 844–855. <https://doi.org/10.1016/j.cej.2018.04.141>
- 545 Ganiyu, S.O., Le, T.X.H., Bechelany, M., Esposito, G., Hullebusch, E.D. van, Oturan, M.A.,
546 Cretin, M., 2017a. A hierarchical CoFe-layered double hydroxide modified carbon-felt
547 cathode for heterogeneous electro-Fenton process. *J. Mater. Chem. A* 5, 3655–3666.
548 <https://doi.org/10.1039/C6TA09100H>
- 549 Ganiyu, S.O., Oturan, N., Raffy, S., Cretin, M., Esmilaire, R., van Hullebusch, E., Esposito, G.,
550 Oturan, M.A., 2016. Sub-stoichiometric titanium oxide (Ti₄O₇) as a suitable ceramic anode
551 for electrooxidation of organic pollutants: A case study of kinetics, mineralization and

552 toxicity assessment of amoxicillin. *Water Research* 106, 171–182.
553 <https://doi.org/10.1016/j.watres.2016.09.056>

554 Ganiyu, S.O., Oturan, N., Raffy, S., Esposito, G., van Hullebusch, E.D., Cretin, M., Oturan,
555 M.A., 2017b. Use of Sub-stoichiometric Titanium Oxide as a Ceramic Electrode in Anodic
556 Oxidation and Electro-Fenton Degradation of the Beta-blocker Propranolol: Degradation
557 Kinetics and Mineralization Pathway. *Electrochimica Acta* 242, 344–354.
558 <https://doi.org/10.1016/j.electacta.2017.05.047>

559 Ganiyu, S. O., Zhou, M., Martínez-Huitle, C. A., 2018. Heterogeneous electro-Fenton and
560 photoelectro-Fenton processes: a critical review of fundamental principles and application
561 for water/wastewater treatment. *Applied Catalysis B: Environmental* 235, 103–129.
562 <https://doi.org/10.1016/j.apcatb.2018.04.044>

563 Ganzenko, O., Huguenot, D., van Hullebusch, E.D., Esposito, G., Oturan, M.A., 2014.
564 Electrochemical advanced oxidation and biological processes for wastewater treatment: a
565 review of the combined approaches. *Environmental Science and Pollution Research* 21,
566 8493–8524. <https://doi.org/10.1007/s11356-014-2770-6>

567 Ganzenko, O., Trelu, C., Papirio, S., Oturan, N., Huguenot, D., van Hullebusch, E.D., Esposito,
568 G., Oturan, M.A., 2018. Bioelectro-Fenton: evaluation of a combined biological—advanced
569 oxidation treatment for pharmaceutical wastewater. *Environmental Science and Pollution*
570 *Research* 25, 20283–20292. <https://doi.org/10.1007/s11356-017-8450-6>

571 García-Rodríguez, O., Bañuelos, J.A., El-Ghenymy, A., Godínez, L.A., Brillas, E., Rodríguez-
572 Valadez, F.J., 2016. Use of a carbon felt–iron oxide air-diffusion cathode for the
573 mineralization of Malachite Green dye by heterogeneous electro-Fenton and UVA
574 photoelectro-Fenton processes. *Journal of Electroanalytical Chemistry* 767, 40–48.
575 <https://doi.org/10.1016/j.jelechem.2016.01.035>

576 Guo, L., Jing, Y., Chaplin, B.P., 2016. Development and Characterization of Ultrafiltration TiO₂
577 Magnéli Phase Reactive Electrochemical Membranes. *Environ. Sci. Technol.* 50, 1428–1436.
578 <https://doi.org/10.1021/acs.est.5b04366>

579 Hu, Y., Lu, Y., Liu, G., Luo, H., Zhang, R., Cai, X., 2018. Effect of the structure of stacked
580 electro-Fenton reactor on treating nanofiltration concentrate of landfill leachate.
581 *Chemosphere* 202, 191–197. <https://doi.org/10.1016/j.chemosphere.2018.03.103>

582 Iglesias, O., Fernández de Dios, M.A., Pazos, M., Sanromán, M.A., 2013. Using iron-loaded
583 sepiolite obtained by adsorption as a catalyst in the electro-Fenton oxidation of Reactive
584 Black 5. *Environmental Science and Pollution Research* 20, 5983–5993.
585 <https://doi.org/10.1007/s11356-013-1610-4>

586 Jacquin, C., Lesage, G., Traber, J., Pronk, W., Heran, M., 2017. Three-dimensional excitation
587 and emission matrix fluorescence (3DEEM) for quick and pseudo-quantitative determination
588 of protein- and humic-like substances in full-scale membrane bioreactor (MBR). *Water*
589 *Research* 118, 82–92. <https://doi.org/10.1016/j.watres.2017.04.009>

590 Jacquin, C., Monnot, M., Hamza, R., Kouadio, Y., Zaviska, F., Merle, T., Lesage, G., Héran, M.,

591 2018. Link between dissolved organic matter transformation and process performance in a
592 membrane bioreactor for urinary nitrogen stabilization. *Environmental Science: Water*
593 *Research & Technology* 4, 806–819. <https://doi.org/10.1039/C8EW00029H>

594 Kjeldsen, P., Barlaz, M.A., Rooker, A.P., Baun, A., Ledin, A., Christensen, T.H., 2002. Present
595 and Long-Term Composition of MSW Landfill Leachate: A Review. *Critical Reviews in*
596 *Environmental Science and Technology* 32, 297–336.
597 <https://doi.org/10.1080/10643380290813462>

598 Li, X., Zhu, W., Wu, Y., Wang, C., Zheng, J., Xu, K., Li, J., 2015. Recovery of potassium from
599 landfill leachate concentrates using a combination of cation-exchange membrane
600 electrolysis and magnesium potassium phosphate crystallization. *Separation and Purification*
601 *Technology* 144, 1–7. <https://doi.org/10.1016/j.seppur.2015.01.035>

602 Ma, L., Zhou, M., Ren, G., Yang, W., Liang, L., 2016. A highly energy-efficient flow-through
603 electro-Fenton process for organic pollutants degradation. *Electrochimica Acta* 200, 222–
604 230. <https://doi.org/10.1016/j.electacta.2016.03.181>

605 Mahindrakar, K.V., Rathod, V.K., 2018. Utilization of banana peels for removal of strontium
606 (II) from water. *Environmental Technology & Innovation* 11, 371–383.
607 <https://doi.org/10.1016/j.eti.2018.06.015>

608 Martín de Vidales, M.J., Millán, M., Sáez, C., Cañizares, P., Rodrigo, M.A., 2016. What
609 happens to inorganic nitrogen species during conductive diamond electrochemical oxidation
610 of real wastewater? *Electrochemistry Communications* 67, 65–68.
611 <https://doi.org/10.1016/j.elecom.2016.03.014>

612 Martínez-Huitle, C.A., Rodrigo, M.A., Sirés, I., Scialdone, O., 2015. Single and Coupled
613 Electrochemical Processes and Reactors for the Abatement of Organic Water Pollutants: A
614 Critical Review. *Chem. Rev.* 115, 13362–13407.
615 <https://doi.org/10.1021/acs.chemrev.5b00361>

616 Mousset, E., Pontvianne, S., Pons, M.-N., 2018. Fate of inorganic nitrogen species under
617 homogeneous Fenton combined with electro-oxidation/reduction treatments in synthetic
618 solutions and reclaimed municipal wastewater. *Chemosphere* 201, 6–12.
619 <https://doi.org/10.1016/j.chemosphere.2018.02.142>

620 Oller, I., Malato, S., Sánchez-Pérez, J.A., 2011. Combination of Advanced Oxidation Processes
621 and biological treatments for wastewater decontamination—A review. *Science of The Total*
622 *Environment* 409, 4141–4166. <https://doi.org/10.1016/j.scitotenv.2010.08.061>

623 Olvera-Vargas, H., Cocerva, T., Oturan, N., Buisson, D., Oturan, M.A., 2015. Bioelectro-
624 Fenton: A sustainable integrated process for removal of organic pollutants from water:
625 Application to mineralization of metoprolol. *J. Hazard. Mater.*
626 <https://doi.org/10.1016/j.jhazmat.2015.12.010>

627 Oturan, M.A., Pimentel, M., Oturan, N., Sirés, I., 2008. Reaction sequence for the
628 mineralization of the short-chain carboxylic acids usually formed upon cleavage of aromatics
629 during electrochemical Fenton treatment. *Electrochimica Acta* 54, 173–182.

630 <https://doi.org/10.1016/j.electacta.2008.08.012>

631 Oturan, N., van Hullebusch, E.D., Zhang, H., Mazeas, L., Budzinski, H., Le Menach, K., Oturan,
632 M.A., 2015. Occurrence and Removal of Organic Micropollutants in Landfill Leachates
633 Treated by Electrochemical Advanced Oxidation Processes. *Environmental Science &*
634 *Technology* 49, 12187–12196. <https://doi.org/10.1021/acs.est.5b02809>

635 Özcan, A., Sahin, Y., Koparal, A.S., Oturan, M.A., 2008. Prophan mineralization in aqueous
636 medium by anodic oxidation using boron-doped diamond anode: Influence of experimental
637 parameters on degradation kinetics and mineralization efficiency. *Water Res.* 42, 2889–
638 2898. <https://doi.org/10.1016/j.watres.2008.02.027>

639 Panizza, M., Cerisola, G., 2009. Direct and mediated anodic oxidation of organic pollutants.
640 *Chem. Rev.* 109, 6541–6569. <https://doi.org/10.1021/cr9001319>

641 Ponthieu, M., Pinel-Raffaitin, P., Le Hecho, I., Mazeas, L., Amouroux, D., Donard, O.F.X.,
642 Potin-Gautier, M., 2007. Speciation analysis of arsenic in landfill leachate. *Water Research*
643 41, 3177–3185. <https://doi.org/10.1016/j.watres.2007.04.026>

644 Poza-Nogueiras, V., Rosales, E., Pazos, M., Sanromán, M. Á. (2018). Current advances and
645 trends in electro-Fenton process using heterogeneous catalysts—a review. *Chemosphere* 201,
646 399-416. <https://doi.org/10.1016/j.chemosphere.2018.03.002>

647 Radjenovic, J., Sedlak, D.L., 2015. Challenges and Opportunities for Electrochemical
648 Processes as Next-Generation Technologies for the Treatment of Contaminated Water.
649 *Environ. Sci. Technol.* 49, 11292–11302. <https://doi.org/10.1021/acs.est.5b02414>

650 Reuschenbach, P., Pagga, U., Strotmann, U., 2003. A critical comparison of respirometric
651 biodegradation tests based on OECD 301 and related test methods. *Water Research* 37,
652 1571–1582. [https://doi.org/10.1016/S0043-1354\(02\)00528-6](https://doi.org/10.1016/S0043-1354(02)00528-6)

653 Sirés, I., Brillas, E., Oturan, M.A., Rodrigo, M.A., Panizza, M., 2014. Electrochemical advanced
654 oxidation processes: today and tomorrow. A review. *Environ Sci Pollut Res* 21, 8336–8367.
655 <https://doi.org/10.1007/s11356-014-2783-1>

656 Trellu, C., Chaplin, B.P., Coetsier, C., Esmilaire, R., Cerneaux, S., Causserand, C., Cretin, M.,
657 2018a. Electro-oxidation of organic pollutants by reactive electrochemical membranes.
658 *Chemosphere* 208, 159–175. <https://doi.org/10.1016/j.chemosphere.2018.05.026>

659 Trellu, C., Coetsier, C., Rouch, J.-C., Esmilaire, R., Rivallin, M., Cretin, M., Causserand, C.,
660 2018b. Mineralization of organic pollutants by anodic oxidation using reactive
661 electrochemical membrane synthesized from carbothermal reduction of TiO₂. *Water*
662 *Research* 131, 310–319. <https://doi.org/10.1016/j.watres.2017.12.070>

663 Trellu, C., Ganzenko, O., Papirio, S., Pechaud, Y., Oturan, N., Huguenot, D., van Hullebusch,
664 E.D., Esposito, G., Oturan, M.A., 2016a. Combination of anodic oxidation and biological
665 treatment for the removal of phenanthrene and Tween 80 from soil washing solution.
666 *Chemical Engineering Journal* 306, 588–596. <https://doi.org/10.1016/j.cej.2016.07.108>

667 Trellu, C., Oturan, N., Pechaud, Y., van Hullebusch, E.D., Esposito, G., Oturan, M.A., 2017.

668 Anodic oxidation of surfactants and organic compounds entrapped in micelles – Selective
669 degradation mechanisms and soil washing solution reuse. *Water Research* 118, 1–11.
670 <https://doi.org/10.1016/j.watres.2017.04.013>

671 Trellu, C., Péchaud, Y., Oturan, N., Mousset, E., Huguenot, D., van Hullebusch, E.D., Esposito,
672 G., Oturan, M.A., 2016b. Comparative study on the removal of humic acids from drinking
673 water by anodic oxidation and electro-Fenton processes: Mineralization efficiency and
674 modelling. *Applied Catalysis B: Environmental* 194, 32–41.
675 <https://doi.org/10.1016/j.apcatb.2016.04.039>

676 Van der Bruggen, B., Koninckx, A., Vandecasteele, C., 2004. Separation of monovalent and
677 divalent ions from aqueous solution by electrodialysis and nanofiltration. *Water Research*
678 38, 1347–1353. <https://doi.org/10.1016/j.watres.2003.11.008>

679 Van der Bruggen, B., Lejon, L., Vandecasteele, C., 2003. Reuse, Treatment, and Discharge of
680 the Concentrate of Pressure-Driven Membrane Processes. *Environmental Science &*
681 *Technology* 37, 3733–3738. <https://doi.org/10.1021/es0201754>

682 Wang, Yujing, Zhao, G., Chai, S., Zhao, H., Wang, Yanbin, 2013. Three-Dimensional
683 Homogeneous Ferrite-Carbon Aerogel: One Pot Fabrication and Enhanced Electro-Fenton
684 Reactivity. *ACS Applied Materials & Interfaces* 5, 842–852.
685 <https://doi.org/10.1021/am302437a>

686 Westerhoff, P., Prapaipong, P., Shock, E., Hillaireau, A., 2008. Antimony leaching from
687 polyethylene terephthalate (PET) plastic used for bottled drinking water. *Water Research* 42,
688 551–556. <https://doi.org/10.1016/j.watres.2007.07.048>

689 Xu, Y., Chen, C., Li, X., Lin, J., Liao, Y., Jin, Z., 2017. Recovery of humic substances from
690 leachate nanofiltration concentrate by a two-stage process of tight ultrafiltration membrane.
691 *Journal of Cleaner Production* 161, 84–94. <https://doi.org/10.1016/j.jclepro.2017.05.095>

692 Zhang, G., Wang, S., Yang, F., 2012. Efficient Adsorption and Combined
693 Heterogeneous/Homogeneous Fenton Oxidation of Amaranth Using Supported Nano-FeOOH
694 As Cathodic Catalysts. *The Journal of Physical Chemistry C* 116, 3623–3634.
695 <https://doi.org/10.1021/jp210167b>

696 Zhang, H., Fei, C., Zhang, D., Tang, F., 2007. Degradation of 4-nitrophenol in aqueous
697 medium by electro-Fenton method. *Journal of Hazardous Materials* 145, 227–232.
698 <https://doi.org/10.1016/j.jhazmat.2006.11.016>

699 Zhang, L., Li, A., Lu, Y., Yan, L., Zhong, S., Deng, C., 2009. Characterization and removal of
700 dissolved organic matter (DOM) from landfill leachate rejected by nanofiltration. *Waste*
701 *Management* 29, 1035–1040. <https://doi.org/10.1016/j.wasman.2008.08.020>

702 Zhang, Q.-Q., Tian, B.-H., Zhang, X., Ghulam, A., Fang, C.-R., He, R., 2013. Investigation on
703 characteristics of leachate and concentrated leachate in three landfill leachate treatment
704 plants. *Waste Management* 33, 2277–2286. <https://doi.org/10.1016/j.wasman.2013.07.021>