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# **Fabrication of PMMA/ZnO nanocomposite: Effect of high nanoparticles loading on the optical and thermal properties**

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## **Abstract**

ZnO nanoparticles (NPs) with a particle size less than 50 nm, were prepared using the solvothermal method from zinc acetate ( $\text{Zn}(\text{CH}_3\text{COO})_2 \cdot 2\text{H}_2\text{O}$ ) and hydrogen peroxide ( $\text{H}_2\text{O}_2$ ) in presence of organic solvent (methanol or ethanol). PMMA/ZnO nanocomposite films were prepared by blending ZnO NPs with polymethylmethacrylate (PMMA) using solution mixing at concentration range 0-40wt%. The results show that the methanol was more efficient in the synthesis of ZnO NPs than ethanol. The optical bandgap of PMMA/ZnO films has been decreased from 4.4 to 3.7eV by increasing the ZnO NPs content from 0 wt% (neat PMMA) to 40 wt%. The thermal stability glass temperature ( $T_g$ ), UV-shielding efficiency (reduced

transmission) were enhanced by about 60 K, 20 K, 40% with increasing ZnO NPs loading from 0 to 40wt%, respectively. The combination of the thermal and optical enhanced properties gives potential applications in different domains *i.e.* optics, photonics, and electronics.

**Keywords:** Zinc oxide, solvothermal method, polymethylmethacrylate nanocomposite, high nanoparticles loading, optical and thermal properties.

## 1. Introduction

Polymer nanocomposites are an attractive class of materials providing novel performance, due to their remarkable properties by using a low to high nanofiller loading [1-2]. Adding inorganic nanoparticle (NPs) to polymers can introduce specific functionalities (luminescence, catalytic activity, magnetism, etc.) and/or improve the electrical conductivity, mechanical properties, thermal stability, and flame retardancy [3]. Keeping the inorganic NPs homogeneously dispersed (isolated) is crucial for creating smart composite materials that retain the flexibility and fluidity of the polymer matrix (host) while incorporating a high fraction of inorganic NPs with their associated properties. However, aggregation of NPs deteriorates the polymer properties (*i.e.* optical, thermal and mechanical properties) [4-5] except for a few cases where a limited NPs aggregation exhibited some improvements [6].

Generally, there are four general approaches for dispersing NPs in polymers [7-9]: (i) High shear mixing of the polymer and NPs by extrusion for polymer melts; (ii) High shear mixing or ultrasonic irradiation for dispersion of NPs in the polymer solution and then evaporation of the solvent. However, dispersion of NPs in polymer matrix cannot be achieved only by mechanical mixing without any chemical bonding and optimized interaction between the matrix and NPs; (iii) in-situ polymerization in the presence of NPs [10]; (iv) modification of the surface by chemical treatment and the grafting of functional polymeric molecules to the active sites (e.g. -OH) existing on the particle surface; (v) both the formation of the

nanoparticles and the polymerization in-situ [11,12]. It has been proved from the literature that homogenous distribution of inorganic NPs in the polymer matrix at a high content of >10wt% NPs is not possible because of the immiscibility and high surface energy of NPs.

Transparent polymer nanocomposites which exhibit various optical properties such as tailored emission/absorption properties, low/high refractive index, or strong nonlinearities optical properties attracts great interest because of the potential optoelectronic applications [13]. These polymer nanocomposites are typically obtained by the incorporation of metal oxide NPs into a transparent polymer matrix (e.g. PMMA). While the polymeric component contributes to good processability, transparency, and flexibility, the NPs provide the desired thermal, electrical and optical properties [14,15]. The technological and economic advantages of polymers have been amply demonstrated in the fields of automotive, chemical sensors and electronics [16,17]. Scientists are particularly interested in polymer-based nanocomposites to combine the properties of NPs and polymers [18-20].

PMMA as a transparent polymeric material offers excellent transparency in the visible and near-infrared range, as well as dielectric constants generally lower than those of the inorganic materials [21], which allows fast response times, shorter than in the case of crystals inorganic compounds. It is often used as an alternative to glass. PMMA has a transmission of more light and much lighter than silica glass (1.19 versus 2.20 g cm<sup>-3</sup> in density, respectively). However, neat PMMA exhibits some disadvantages, for example, poor thermal stability, which restricts it from high-temperature applications. PMMA transmits UV light thus manufacturers usually apply UV coatings on PMMA to overcome this deficiency [22]. A variety of inorganic nanofillers were mixed with an organic matrix, including PMMA/nanoclay [23-24], PMMA/SiO<sub>2</sub> [25,26] and PMMA/MWCNTs [27]. These types of systems were showed improved optical properties allow applications in optoelectronics and optical domains [28,29]. PMMA/ZnO nanocomposites were extensively studied in the literature [22,30-34]. It may be

due mainly to the wide number of advantages presented in different domains (optics, electronics and the simple methods of preparation). ZnO NPs is a well-known multifunctional inorganic filler that has outstanding properties such high refractive index, large exciton binding energy (nearly 60 meV) [35-36], high thermal conductivity and antibacterial and UV-protection properties [22]. ZnO NPs usually exhibit a near-band-edge (NBE) UV line accompanied by a strong visible luminescence. However, the most crucial aspects of high luminescence efficiencies from ZnO NPs are the surface texture and an efficient surface passivation. Recent studies show that capping of ZnO NPs with polymers can effectively passivate the surface and reduce the surface-related visible emission [37-38].

Many studies were focused on the thermal degradation behavior of the PMMA/ZnO nanocomposites, because of its properties and applications in chromatography, medicine, and optics [34]. The blending of ZnO NPs with PMMA is believed to improve the optical properties of the polymer, however, refractive index mismatch may cause strong optical scattering which diminishes the transparency of the composites. The mismatch is tolerable when the particle size ( $2r$ ) is much smaller than the wavelength ( $\lambda$ ) of the incident light ( $2r < \lambda/10$ ) [9]. ZnO presents other properties as n-type semiconductor with a wide bandgap (3.36-3.39eV) [39], and presents appealing features such as biocompatibility, nontoxicity and inexpensiveness [40,41]. Thus it has been extensively used in the fabrication of electronic and optical devices [30]. ZnO NPs are environmentally friendly [42] and nontoxic to the cells comparing with many other inorganic NPs such as CdTe, CdS and CdSe [43].

The present study focuses on the effect of blending ZnO NPs with PMMA. Firstly, the ZnO NPs was synthesized using the solvothermal method. The obtained ZnO NPs were blended with PMMA by solution mixing to prepare PMMA/ZnO nanocomposite films. The dispersion of ZnO NPs and morphology, thermal stability and UV-shielding capability of the nanocomposite films were investigated by scanning electron microscopy (SEM), thermal

gravimetric analysis (TGA), differential scanning calorimetry (DSC), and UV-vis spectroscopy.

## **2. Experimental Section**

### **2.1. Synthesis of ZnO nanoparticles**

Approximately 8.4 g of zinc acetate dihydrate ( $\text{Zn}(\text{CH}_3\text{COO})_2 \cdot 2\text{H}_2\text{O}$ , Panreac, 99%) was added to 266 ml of doubly distilled water with continuous stirring for 25 min at 50°C. Subsequently, 400 ml of methanol ( $\text{CH}_3\text{OH}$ , Sigma-Aldrich, 99.7%) or ethanol ( $\text{C}_2\text{H}_5\text{OH}$ , Sigma-Aldrich, 96%) was added gradually with continuous stirring for 3 h. Then 4 ml of hydrogen peroxide ( $\text{H}_2\text{O}_2$ , Panreac, 47%) was added dropwise and with stirring until a clear solution is obtained. This solution was incubated for 24 h at ambient temperature and pressure and then dried at 80°C for 6 h to obtain a white powder. The white powder was washed three times with doubly distilled water. After washing, the ZnO powder was dried at 80 °C in a drying oven and then calcined at 600 °C for 1 h.

### **2.2. Preparation of nanocomposite films**

The granulate PMMA (Vedril Spa-Resina Metallica) manufactured (Motedison as a supplier) was used, with  $T_g = 105^\circ\text{C}$ . In order to prepare the nanocomposite film, the PMMA granulate (1 g) was dissolved in 50 ml of THF solvent when the mixture becomes white and viscous. Different amounts of the ZnO powder (ZnO NPs) were added. The solution was stirred at 60°C for 90 min. When the solution becomes white, the gel was spread on the glass plate to obtain a nanocomposite film. The thin film was placed in an oven heated to 70 °C for 11 min to remove the trace solvent. The thickness of PMMA and PMMA/ZnO films was measured by a micrometer (Vernier Digital Palmer) at different positions (4 to 6 points), the mean film thickness kept at about ( $20 \pm 1 \mu\text{m}$ ). The composition of the nanocomposites film was summarized in Table 1.

**Table 1:** Nanocomposites prepared with different contents of ZnO in THF solvent.

Samples	PMMA (g)	ZnO (g)
Neat PMMA	1.00	0.00
PMMA/ZnO 1%	0.99	0.01
PMMA/ZnO 2%	0.98	0.02
PMMA/ZnO 5%	0.95	0.05
PMMA/ZnO 10%	0.90	0.10
PMMA/ZnO 40%	0.60	0.40

### 2.3. Characterization of ZnO NPs and PMMA/ZnO nanocomposite

Phase identification of the prepared ZnO NPs was performed using an X-ray diffractometer (PANalyticalXpert-PRO diffractometer equipped with an accelerator detector using Ni-filtered Cu-radiation) with Cu K $\alpha$  radiation ( $\lambda=0.15406$  nm) at 40 kV and 20 mA. The scanning step size was 0.0167113°. The crystalline size of ZnO particles was determined from the XRD data using the Debye-Scherrer equation [44].

$$d = \frac{K\lambda}{\beta \cos \theta}$$

Where ( $d$ ) is the crystallite size, ( $K$ ) is a dimensionless shape factor, with a value close to unity. The shape factor has a typical value of about 0.9, but varies with the actual shape of crystallite, ( $\beta$ ) is the full-width at half maximum (FWHM) of the most intense diffraction peak, ( $\lambda$ ) is the wavelength of Cu target, and ( $\theta$ ) is the Bragg angle.

The morphology and the chemical composition of the PMMA/ZnO films were examined by field-emission scanning electron microscopy (FESEM) using Quanta 650 model with an acceleration voltage of 10 kV and low vacuum.

A Varian-Cary system 500 UV-VIS-NIR spectrophotometer was used to determine the optical properties (UV-vis absorption) of the nanocomposites films at ambient temperature. The optical bandgap energy of different nanocomposites PMMA/ZnO was calculated using the following formula [45].

$$(\alpha h\nu) = A(h\nu - E_g)^{1/2}$$

Where  $\alpha$ ,  $h$ ,  $\nu$ ,  $E_g$  and  $A$  are absorption coefficient, Planck constant, light frequency, band gap energy and  $A$  constant, respectively. The bandgap energy ( $E_g$ ) values of different nanocomposites films were calculated from the intercept of  $(\alpha h\nu)^2$  versus  $(h\nu)$  curves by extra-plotting tangent to them.

Thermogravimetric analysis (TA Instruments, TGA Q500, USA) was used to detect the thermal degradation of neat PMMA and PMMA/ZnO nanocomposites with different amounts of ZnO. The samples were heated to 60°C and held at this temperature for 30 min and heated at 10 K.min<sup>-1</sup> from 60 to 700 °C under air atmosphere. The residue (at 600 °C) on the TGA curves presents the estimated amount of ZnO in the prepared nanocomposites. Thermal characterization was performed using a TA Instruments Q20 DSC to evaluate the glass transition of the nanocomposites. Samples were heated from 20 to 200°C at a heating/cooling rate of 10 K.min<sup>-1</sup> under N<sub>2</sub> atmosphere. The glass transition temperature was determined using universal analysis software provided by TA.

### **3. Results and discussion**

#### **3.1. Influence of solvent nature on ZnO NPs preparation**

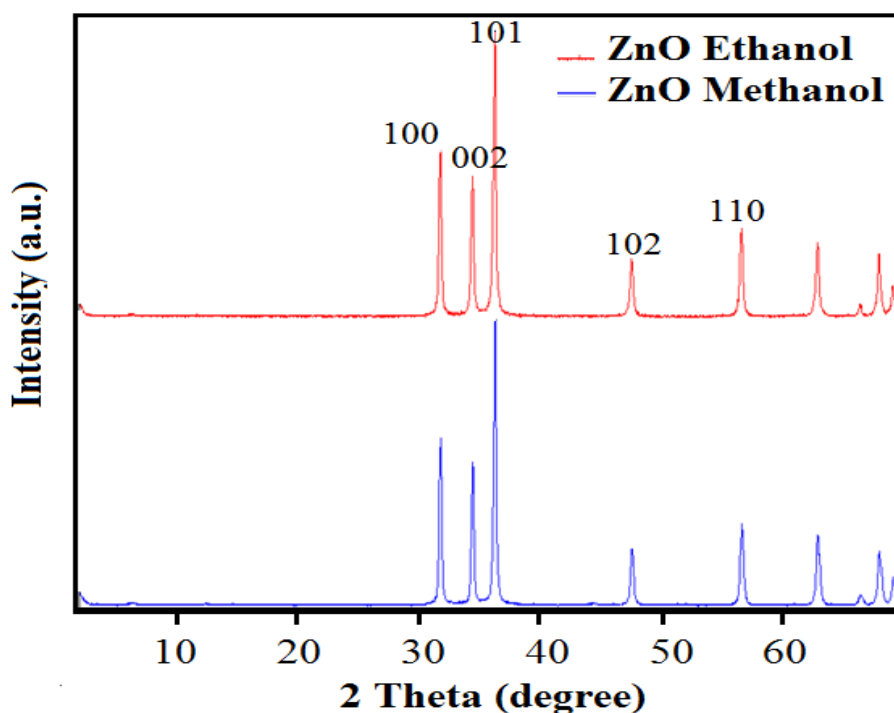
The solvothermal method is one of the most efficient and easiest methods to produce ZnONPs. In a solvothermal synthesis of ZnO NPs, a solvent (methanol or ethanol) acts as a reaction medium. Using different types of organic solvents with different structures, single crystalline ZnO NPs of different morphologies can be obtained [46]. In this study, two types of



solvent were used in the synthesis of the ZnO NPs in order to evaluate the effect of different solvents. ZnO NPs with different yields were obtained in methanol and ethanol solvent. The mass yield of ZnO NPs obtained in presence of methanol is much higher than that obtained in ethanol. This means that the conversion of the hydroxides into ZnO was promoted in presence of methanol (CH<sub>3</sub>-OH) more than ethanol (CH<sub>3</sub>-CH<sub>2</sub>-OH). There are several factors may contribute the hydrolysis of Zn<sup>2+</sup> and formation of ZnO NPs: length of the alkyl chain, steric hindrance, and polarity of the solvent molecules. Based on this result the methanol solvent was suggested in all preparations.

### **3.2. Polymorph and crystal structure of ZnO NPs**

The X-ray diffraction pattern of the obtained ZnO nanoparticles is shown in Fig. 1. All sharp diffraction peaks at  $2\theta = 31.7$  (100),  $34.4$  (002) and  $36.2$  (101) in the two graphs prepared in ethanol and methanol solvents. The synthesis of ZnO generally gives wurtzite [47] or zincite hexagonal crystal structure [13,48]. The prepared ZnO were indexed to the zincite hexagonal crystal structures. The peak positions were in an agreement with those observed by Yadav *et al.* [48]. The X-ray diffraction results and the average crystallite size value calculated by the Scherer equation confirm that the ZnO crystallites elaborated either in methanol or in ethanol gave the crystalline size of about 44 nm

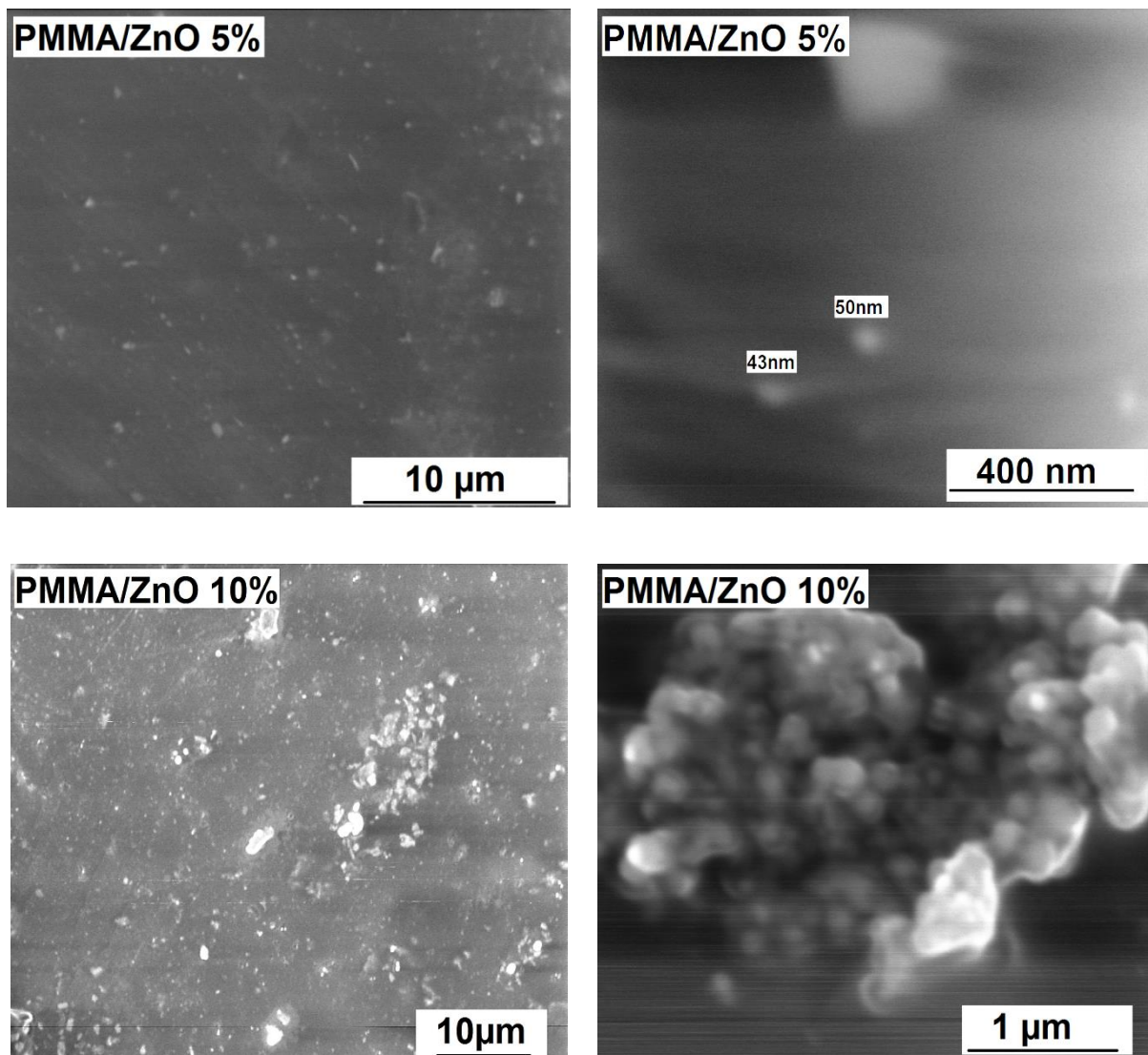


**Fig. 1** XRD pattern of synthesis of ZnO NPs prepared in presence of ethanol and methanol.

### 3-3. Morphological development of PMMA/ZnO nanocomposites

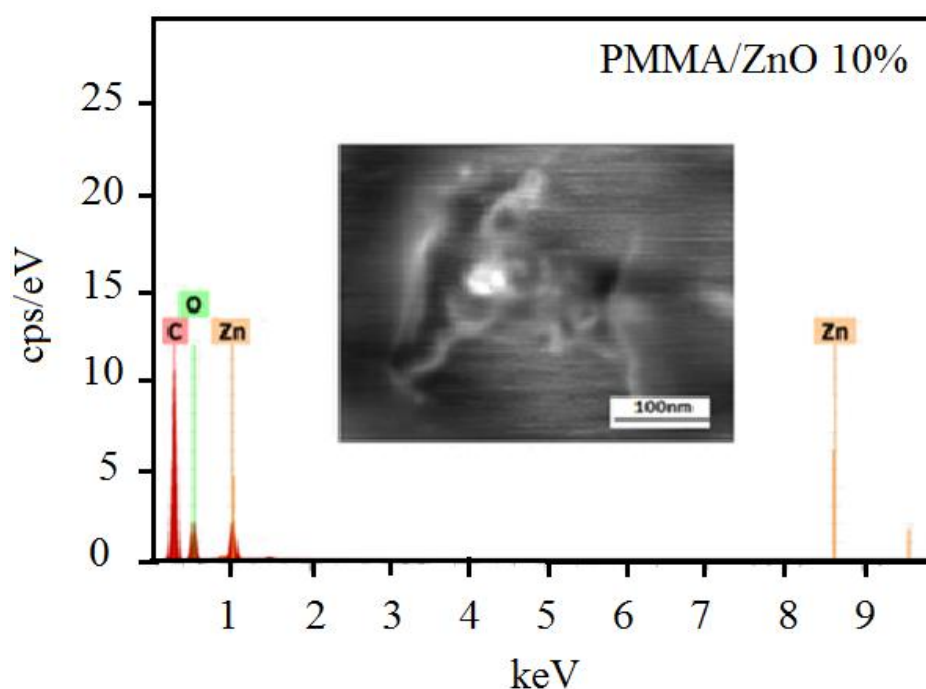
The evolution of ZnO NPs dispersion is important since the well-dispersed NPs can lead to a significant improvement in PMMA properties. Homogeneous dispersion of ZnO NPs is desired to avoid creation of crack-initiatives in the PMMA film. In order to visualize the effect of high ZnO NPs loaded (0-40wt%) on the distribution and the dispersion of the nanofiller in the polymer PMMA matrix, SEM was used. Fig. 2 shows SEM images of the PMMA/ZnO films prepared with 5 and 10 wt% ZnO NPs. The 5wt% PMMA/ZnO film shows a good dispersion of the ZnO NPs without agglomeration. The average diameter of the dispersed phase (Fig. 2) was about 50 nm. The bright phase presents the ZnO NPs with nonuniform distribution when the amount of ZnO NPs increases up to 10wt%. The agglomeration observed at 10wt% PMMA/ZnO film presents a poor dispersion of ZnO NPs in the polymer matrix (PMMA). The average size of the agglomeration was estimated equal to 3  $\mu\text{m}$ , the average size of the agglomerate increases from 0.3  $\mu\text{m}$  to 3  $\mu\text{m}$  with increasing the ZnO NPs loading from 5wt% to 10wt%.

This phenomenon was the most challenging in the dispersion of the nanoparticles in the polymeric matrix, because the NPs tend to agglomerate due to high surface energy [45,49]. The size of the aggregation has a considerable influence on thermal and optical properties of the nanocomposites [13]. A very narrow size distribution is strictly required because scattering intensity is proportional to  $(r^3)$  ( $r$ : particle radius). The presence of even a small percentage of particles or aggregates, therefore, larger than 100 nm leads to strong scattering in the visible and therefore causes haze or even turbidity [13].



**Fig. 2:** Morphological development of PMMA/ZnO nanocomposites with 5 and 10wt% of ZnO in PMMA.

Fig. 3 presents the energy dispersive X-ray (EDX) analysis of 10wt% PMMA/ZnO nanocomposites. The EDX analysis shows the existence of three characteristic elements of the polymer and nanoparticle (PMMA/ZnO) which are represented by the C, Zn and O peaks. The different elements exist in the materials that we are used during the preparation of the nanocomposites, indicating the successful formation of nanocomposite with high purity. The C peak is attributed to the PMMA while those for Zn are attributed to ZnO NPs and the O peak is a contribution between PMMA and ZnO.

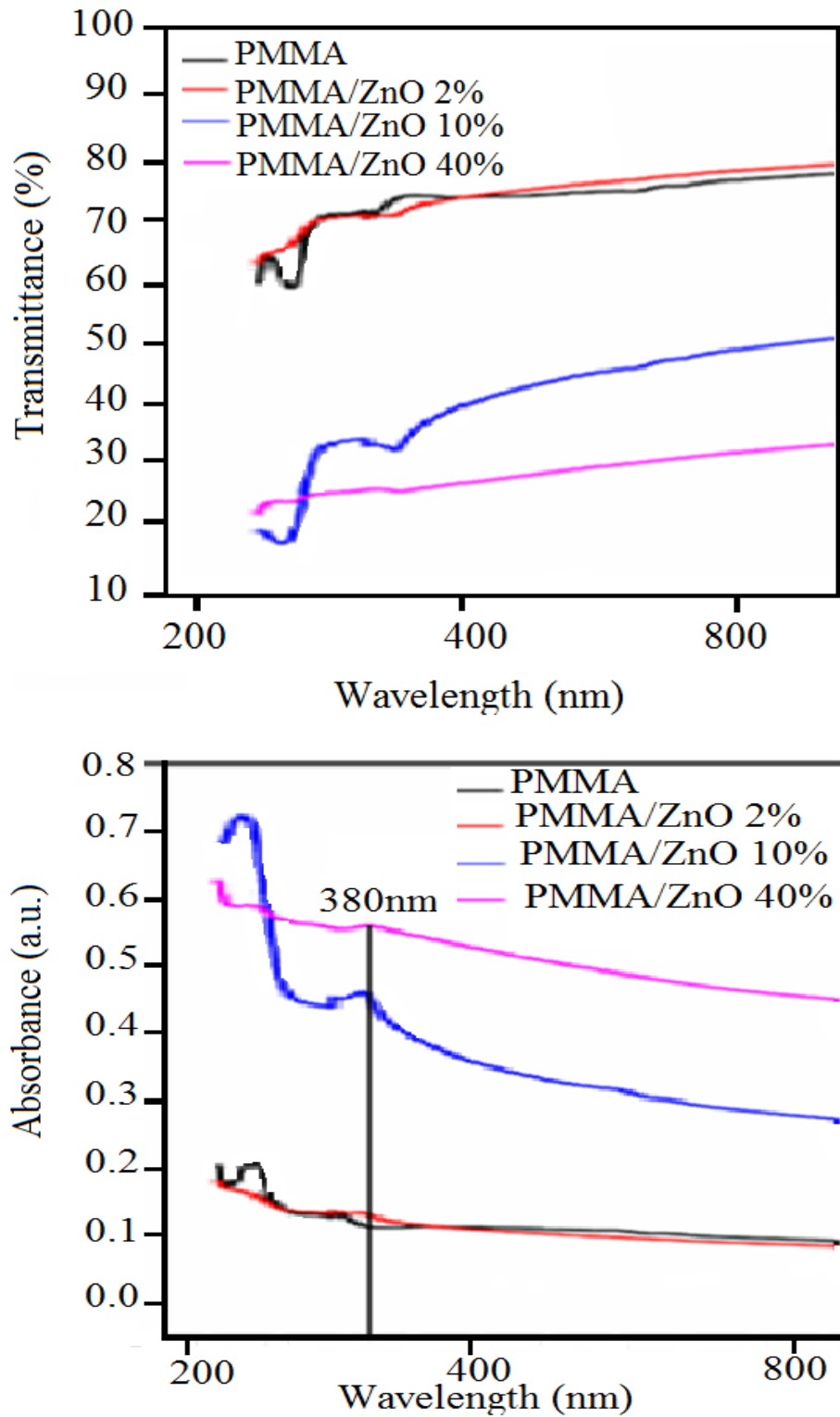


**Fig. 3 SEM-EDX analysis of the nanocomposite film PMMA/ZnO 10 wt%.**

### **3.5. Tuning the optical properties of PMMA/ZnO film**

Fig. 4 shows the UV-vis transmission and absorption spectra of PMMA and PMMA/ZnO nanocomposites as a function of wavelength. The PMMA/ZnO films loaded with 2wt% ZnO NPs has a high transparency in the visible range compared to the films loaded with higher ZnO contents 10-40 wt%. The increase of ZnO NPs content in the PMMA matrix from 2 to 10 wt% reduces the transmission of light in the visible range up to 33-37%. This drop-in transparency is mainly attributed to the scattering of the light with relatively short wavelengths

by the larger aggregates formed in the PMMA/ZnO films as ZnO NPs concentration increases from 2-10 wt%. The transmission of PMMA/ZnO films decreases further with increasing ZnO NPs concentration from 10-40wt%. Aggregation of ZnO NPs at high loading from 10-40 wt% results in a high UV-reflectivity but also leads to a lower factor of transmission for the visible spectrum. Demir et al. [13] reported that blending of ZnO NPs with PMMA is believed to improve the optical properties of the PMMA/ZnO film. However, the difference between refractive index of ZnO and that of PMMA (refractive index mismatch) results in a strong optical scattering which diminishes the transparency of the nanocomposites [13]. Fig. 4b shows that the absorption maximum of the PMMA/ZnO films is shifted to higher wavelengths compared to the neat PMMA. The neat PMMA and 2wt% PMMA/ZnO films do not absorb UV light down to 300 nm in wavelength, while the PMMA/ZnO films containing ZnO NPs (10-40wt%) absorb UV-light at around 380 nm ( $\lambda_{onset}$ ), down to 250 nm. The peaks of absorption of PMMA/ZnO films shown at ~380nm (3.29eV) corresponding to the exciton state in the pure ZnO NPs [50]. The 10wt% PMMA/ZnO films containing show good UV-shielding efficiency, and the more ZnO NPs in PMMA, the more UV light is absorbed.



**Fig. 4** UV-vis transmission and absorbance spectra of the nanocomposite films with different ZnO NPs concentrations (0, 2, 10 and 40wt% of ZnO NPs).

The measurement of optical bandgap energy was performed using Tau's plot, plotting  $((\alpha h\nu)^2)$  as a function of  $(h\nu)$ ,  $\alpha$  is absorption coefficient and  $\nu$  is the frequency of light used. The optical bandgap was determined from the extrapolation of the linear portion of the graph to the photon energy axis (Fig. 5a). The optical bandgap of PMMA/ZnO films has been decreased from  $4.4\pm 0.2$  to  $3.7\pm 0.2$  eV by increasing the ZnO NPs content from 0 wt% (neat PMMA) to 40 wt%.

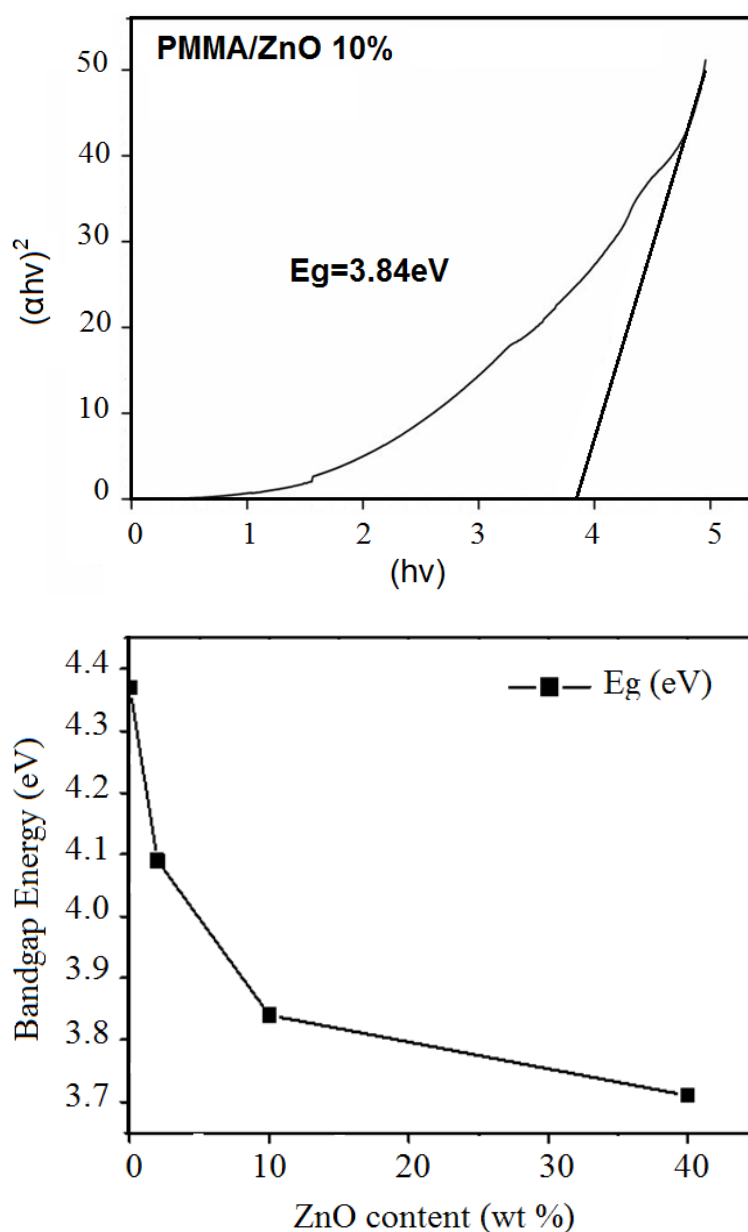


Fig. 5. (a) Variation of  $(\alpha h\nu)^2$  as a function of  $(h\nu)$  of PMMA/ZnO 10wt% nanocomposites, (b) Variation of the optical bandgap of the PMMA/ZnO nanocomposite films as a function of ZnO NPs content.

Figure 5b shows the variation of the optical bandgap energy of nanocomposites films versus the ZnO NPs loading. The optical bandgap is inversely proportional to the ZnO NPs loading in PMMA/ZnO nanocomposites. The bandgap values decrease from 4.37eV for pure PMMA to 3.71eV when the amount of ZnO NPs reached 40 wt%. Therefore, this decrease in the bandgap energy is mainly due to the charges are given by the ZnO NPs and the increase of the ZnO size particle due to the agglomeration phenomenon. The experimental results of the bandgap energy are in good agreement with predicted theoretical results [51,52]. These results also are constituents with by Lee *et al.* [52] who reported that the increase of the particle size (aggregate size) of the NPs increases significantly the bandgap energy of the nanocomposite. Similar results were shown by Thakur et al. [53] in PVC/PE/TiO<sub>2</sub>, the optical bandgap decreases with the increase of the TiO<sub>2</sub> NPs content in the blend polymer.

### **3-6. Glass transition temperature**

The thermograms of neat PMMA and PMMA/ZnO nanocomposites are shown in Fig. 6. All the PMMA/ZnO nanocomposite films show glass transition temperature (T<sub>g</sub>) higher than that of the neat PMMA. T<sub>g</sub> of the neat PMMA was about 90.5°C (Table 2). This value of T<sub>g</sub> is much lower of about 10 K in comparison to the value gives by the supplier (102 °C). This may be due to the dissolution of the PMMA in the solvent that diminishes the molecular weight distribution (MWD) since the glass transition temperature (T<sub>g</sub>) is directly related to the polymer chain mobility [54].



Table 2 shows that all nanocomposites present a higher T<sub>g</sub> compared to that of neat PMMA. Enhancement of T<sub>g</sub> of the PMMA matrix with the addition of inorganic NPs was discussed in the literature. For example, Nikolaidis *et al.* [55] show that the increase of T<sub>g</sub> of nanocomposites (PMMA/Cloisite) was due to the production of materials with higher MWD compared to the neat PMMA. This can be attributed to the increase in the length of the macroradicals before they find one another and terminate. The dispersion of ZnO NPs in the PMMA matrix will hinder the movement of macromolecules and increase the T<sub>g</sub> of the sample. At a low loading of ZnO in the nanocomposites, the T<sub>g</sub> increase of 5 to 10 K (Table 2). At the high amount of ZnO (40 wt%), a spectacular enhancement of T<sub>g</sub> of about 22K is observed. Khan *et al.* [56] show that the addition of 14 nm ZnO NPs enhances the T<sub>g</sub> of the nanocomposites by about 26 K at 1 wt% of ZnO. The difference between our results and Khan *et al.* [56] is probably due to the aggregate size of ZnO NPs in the PMMA matrix.

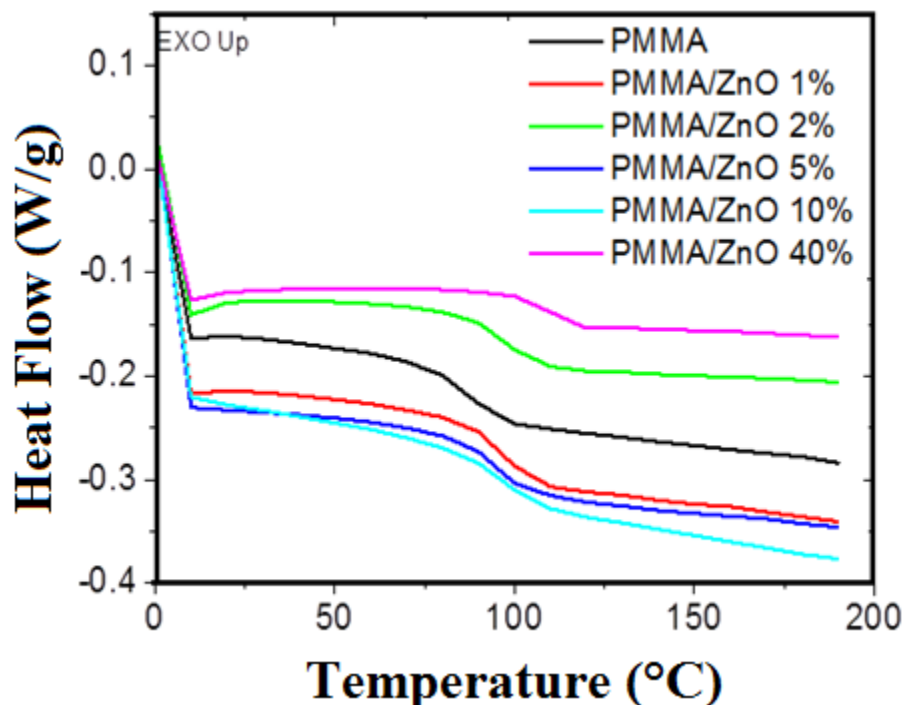


Fig. 6: DSC curves of the PMMA/ZnO nanocomposite films at different ZnO NPs content.

### 3.7. Thermal stability

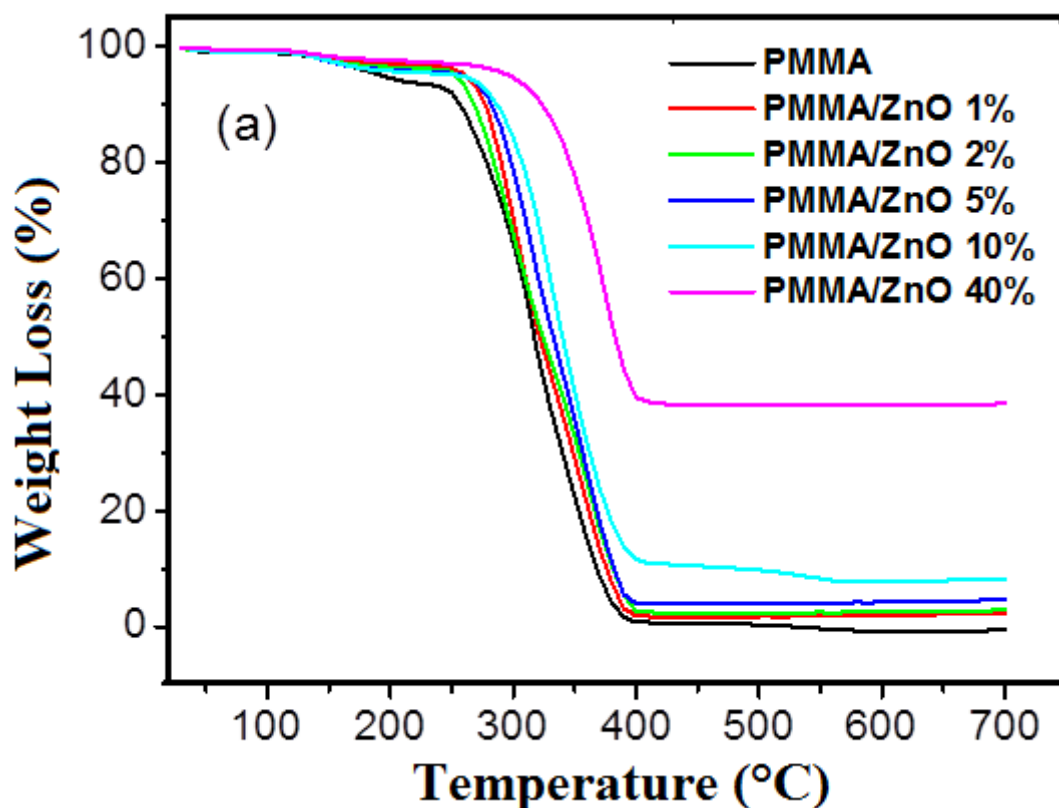
Figure 7 shows TG and DTG curves of neat PMMA and PMMA/ZnO nanocomposites with different contents of ZnO NPs. The neat PMMA (Fig. 8b) was presented by three main steps of degradation. The last step was completely separated from the other. The first step was between  $T=102^{\circ}\text{C}$  and  $230^{\circ}\text{C}$ , mass loss of about 7 wt% was attributed to the evaporation of residual solvent [56-58] and probably a depolymerization step initiated at weak head-to-head linkages arising from chain termination by combination and vinylidene chain ends that result from disproportionation [59,60]. The interval of the second step was situated between  $230^{\circ}\text{C}$  and  $404^{\circ}\text{C}$ . This DTG peak refers to degradation initiated by random scission of the polymer backbone of neat PMMA, about 92 wt% of the total weight of the PMMA sample. The peak presents an overlapping due probably to the complexity of the decomposition of PMMA. This degradation was attributed to the radical transfer to unsaturated chain ends and random scission [61].

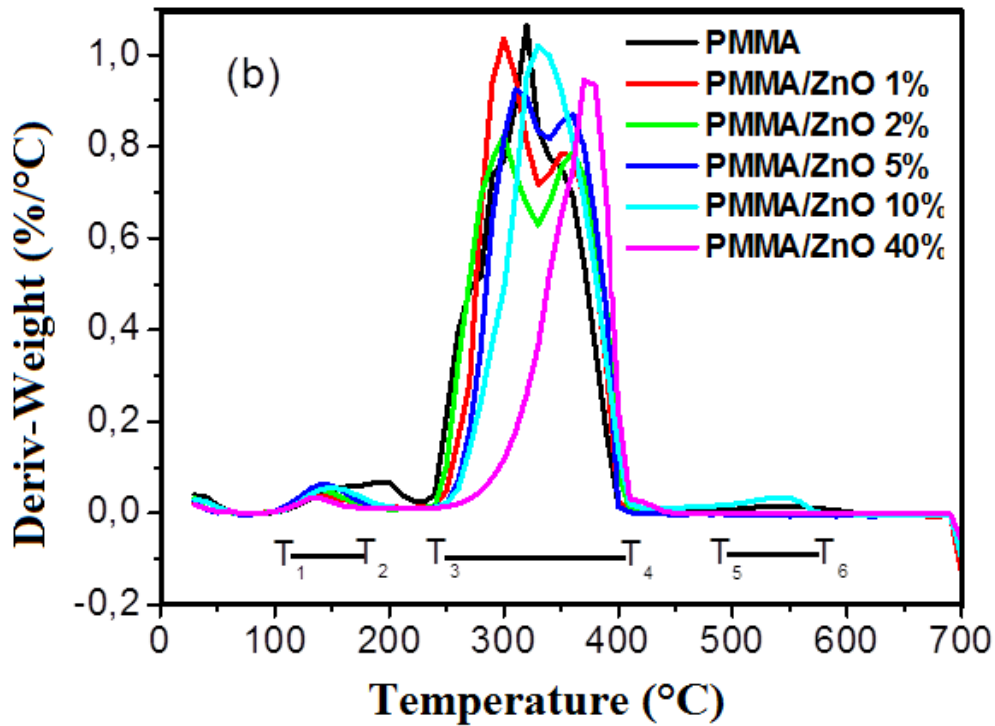
It can be seen that TG and DTG curves of PMMA/ZnO nanocomposites were different from neat PMMA. All PMMA/ZnO nanocomposites show that the first peak ( $T_{\text{peak}}$ ) was delayed of about 50 K compared to the neat PMMA with increasing the ZnO NPs content upto 40 wt% (Figure 7a). This may be due to the facile liberation of the residual solvent due to the enhancement of the thermal conductivity in the samples with the presence of the ZnO NPs.

The main peak of DTG of neat PMMA (second degradation) was disturbed by different little peaks. The addition of 1-5wt % of ZnO to PMMA eliminated this anomaly and the main degradation was presented by two succeeds peaks. The same shape of degradation was preserved in PMMA/ZnO with 2 and 5wt% nanocomposites. The DTG curve of the 10wt% and 40wt% PMMA/ZnO nanocomposites show only one sharp peak at  $415\pm 5$  and  $443\pm 5$  °C. The disappearance of the third step (see Table 2) occurs gradually with increasing ZnO NPs but it is very significant even at 10-40% ZnO NPs. This result was in good agreement with Demir *et*

al. [13] results, who prepared PMMA/ZnO composite through in-situ polymerization of (MMA) monomers in the presence of modified ZnO NPs. The thermal stability properties were increased with increasing the ZnO NPs loading.

The addition of ZnO NPs in PMMA enhances the stability of the nanocomposites (Fig. 7b). A spectacular enhancement was observed at higher ZnO content when the main degradation peak increases of about 60K compared to the main peak of PMMA. The enhancement of the thermal stability of PMMA prepared with increasing ZnO NPs can be attributed to the formation of char which hinders the out-diffusion of the volatile decomposition products. The ZnO NPs remaining at the surface hinder the diffusion of oxygen into the bulk PMMA film thus slowing down the last step of the degradation process [3]. The analysis of the residue results by TG curves (Table 2) shows that the amounts of ZnO NPs revealed by TG curves was superior to that estimated by our preparation in a low amount of ZnO NPs (Table 2).





**Fig. 8: Thermograms of the neat PMMA and nanocomposites PMMA/ZnO with different ZnO NPs content: (a) TGA curves; and (b) derivative DTG curves**

**Table 2: TGA and DSC data of the neat PMMA and PMMA/ZnO nanocomposite films**

Sample	T <sub>g</sub>	Thermal Stability									Residue
		Step 1			Step 2			Step 3			
	°C	T <sub>1</sub> (°C)	T <sub>2</sub> (°C)	Wt loss (%)	T <sub>3</sub> (°C)	T <sub>4</sub> (°C)	Wt loss (%)	T <sub>5</sub> (°C)	T <sub>6</sub> (°C)	Wt loss (%)	ZnO wt%
Neat PMMA	90.5	132	230	6.5	230	403	92.5	499	637	0.8	0.0
PMMA/ZnO(1%)	97.8	115	212	5.0	230	331	52	331	402	50.3	1.7
PMMA/ZnO(2%)	100.2	104	220	4.7	220	326	47.5	326	409	44.2	2.3
PMMA/ZnO(5%)	95.6	117	222	4.3	241	337	49.2	337	405	45.2	4.3
PMMA/ZnO(10%)	96.0	118	230	4.5	239	415	89.6	470	577	2.5	7.9

PMMA/ZnO(40%)	112.6	107	214	2.5	241	443	59.2	-	-	-	38.3
T1, T3, and T5 represent the start points of weight loss steps 1, 2, 3 where T2, T4, and T6 represent the endpoints of the weight loss steps 1, 2, 3, respectively.											

#### 4. Conclusion

Homogeneous dispersion of inorganic NPs, especially at high NPs content >10wt%, in a polymeric matrix is highly challenging. In this study, the effect ZnO NPs loading at PMMA films (0-40 wt%) on the morphological, thermal and optical properties of PMMA was investigated. The optical properties of the PMMA/ZnO nanocomposites show that all films are transparent at low ZnO loading. The increase of ZnO NPs content from 0 to 40 wt% in PMMA matrix decreases the optical bandgap from 4.4 to 3.7eV, respectively. The increase of ZnO NP content in the nanocomposites films increases the glass transition temperature (+20 K) and enhances the thermal stability (+60 K) of the PMMA/ZnO films along with UV-shielding capability. However, the pure PMMA and PMMA films loaded with 1-2 wt% ZnO NPs has a high transparency in the visible range compared to the films loaded with higher ZnO contents 10-40 wt%. The combination of the enhanced properties of PMMA/ZnO films opens a large way to these nanocomposites in different applications (solar energy and optical domains).

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**No conflicts of interest exist**

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